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Defective ZnCo$_2$O$_4$ with Zn vacancies: synthesis, property and electrochemical application

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ABSTRACT: Through the liquid-phase co-precipitation and alkaline-tailored method, the defective ZnCo$_2$O$_4$ with Zn vacancies (Zn$_{0.95}$Co$_2$O$_4$) has been synthesized, which is similar to the crystal phase, morphology, and particle size of the pure ZnCo$_2$O$_4$ before etched, except the enlarged BET specific surface. For the first time, the Zn$_{0.95}$Co$_2$O$_4$ has been evaluated as an anode material for lithium-ion batteries. The Zn vacancies in defective ZnCo$_2$O$_4$ may decrease the probability of the reversible by-reaction between Zn and Li-Zn alloy by the cyclic voltammogram measurement. Compared to the traditional ZnCo$_2$O$_4$, the Zn vacancies in defective ZnCo$_2$O$_4$ can provide larger interface, activate more reaction sites and expand faster transport paths for both of Li-ions and electronics insertion/extraction, so the electrochemical performance of defective ZnCo$_2$O$_4$ has been enhanced highly. The discharge capacity

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retains at 652.2 mAh·g^{-1} under 0.4 A·g^{-1} after 200 cycles. When the rate returns to 0.4 Ah·g^{-1}, the average discharge capacity could be recovered to 748.9 mAh·g^{-1} under the multiple-step high rates after many cycles.

KEYWORDS: $ZnCo_2O_4$, vacancy, defect, nanosheet, anode, energy storage materials

1. Introduction

Li-ion batteries (LIBs) with high energy density are considered to apply in energy storage systems (ESSs), hybrid electric vehicles (HEVs) and electric vehicles (EVs) [1,2]. But the commercial anode (graphite) of LIBs is the shortest board due to the low theoretical capacity (372 mAh·g^{-1}) and the limited hosts of Li-ion intercalation [3]. So kinds of potential anode materials with higher theoretical capacity have been widely investigated including $Li_4Ti_5O_{12}$, Si, graphene, TiO_2, SnO_2, Co_3O_4, MoS_2 and so on [4,5]. Among them, the transition metal oxides (TMOs) (Co_3O_4, NiO, CoO, MnO_2, *et al.*) have been focused on account of hosting more Li-ions on the conversion reaction [6-11]. However, the practical applications are still hampered by the poor electrical conductivity, high voltage plateaus, and large volume expansion of TMOs [1,3,4]. As a result, $ZnCo_2O$, one of binary TMOs, is considered as an attractive candidate, since Zn owns good electrochemical activity for Li-ion insertion/extraction [12-17]. Thus far, kinds of $ZnCo_2O_4$ materials and composites have been prepared via kinds of methods as follows. Through a solvothermal self-assembling process, flower-like porous $ZnCo_2O_4$ microspheres (1.0-1.2 µm) delivers a high reversible specific capacity of 940 mAh·g^{-1} after 100 cycles at 0.1 A·g^{-1} [15]. $ZnCo_2O_4$/graphene

has been prepared via a urea-assisted auto-combustion method with an enhanced rate capability of 378.1 mAh·g^{-1} at 4.5 C [16]. Self-supported ZnO/ZnCo$_2$O$_4$ submicron rod arrays on Ni foil by an ammonia-evaporation-induced method have improved the Li-ion transfer and electrolyte permeability [17]. ZnCo$_2$O$_4$@C with 3D porous wrinkled-paper-like structure via a hydrothermal method has shown a fair cycling property [18]. Yolk-shell ZnCo$_2$O$_4$ through an ultrasonic spray pyrolysis has possessed a reasonable structure and heat stability at 60 °C under 3 A·g^{-1} after 200 cycles [19]. Nevertheless, the electrochemical performance of ZnCo$_2$O$_4$, especially the rate capacity and the cycling stability, should be further developed.

Remarkably, defects and particular vacancies at metal oxides (MOs) play an important role in a variety of applications [20,21]. These vacancies may be considered to be reactive agents, adsorption sites, electron donor sites, or small metal clusters, which directly influence electronic structure, charge transport, surface steady, and optoelectronic reactivity [21]. By annealing in an appropriate atmosphere, single-crystalline ZnO samples with Zn or O vacancies could affect the near band edge (NBE) signals directly [22]. And defective TiO$_2$ with O vacancies via air plasma treatment is intimately related to the optical absorption region/red shift (from 3.22 to 3.00 eV) [23,24]. In addition, the defect adjustment can regulate effectively the electronic properties of different catalysts. For example, Co$_3$O$_4$ with O vacancies on the surface is an enhanced catalyst for ethylene hydrogenation [25]; and MnO$_2$ via adjusting Mn Vacancies has a dramatic catalytic reaction active for HCHO oxidation [26]. So it is necessary to realize the distinct relationship between the Zn vacancies in

$ZnCo_2O_4$ and the electrochemical performance. Hopefully, the research would stimulate the interest in the development of TMOs with defects.

Herein, novel defective $ZnCo_2O_4$ with Zn vacancies has been designed and prepared successfully by alkaline-tailored method at low temperature. To the best of our knowledge, this is the first report on the electrochemical performance of defective $ZnCo_2O_4$ with Zn vacancies. Evaluated as an anode material for LIBs and compared with the similar normal $ZnCo_2O_4$, the exceptional electrochemical performance including higher rate capability and better capacity retention is identified. And the finding may reveal an innovative method to enhance the electrochemical activity of TMOs via vacancies. The details are further discussed below.

2. **Experimental Section**

Materials. All reagents including $CoSO_4 \cdot 7H_2O$ (molecular weight 281.10), $ZnSO_4 \cdot 7H_2O$ (molecular weight 287.56), NaOH (molecular weight 40.00), Tween-40 (Polyoxyethylene 40 sorbitan monopalmitate, molecular weight 1283.62) are the analytical grade.

Materials Preparation. The method of $ZnCo_2O_4$ precursor (mixture of $Co(OH)_2$ and $Zn(OH)_2$) by liquid-phase co-precipitation was similar to our previous work [27-29]. As a typical synthesis, $CoSO_4 \cdot 7H_2O$ (2.8110 g) and $ZnSO_4 \cdot 7H_2O$ (2.876 g) were dissolved in deionized water to form a 100 mL homogeneous solution (Solution A); while NaOH (0.800 g) as a precipitation agent and Tween-40 (0.010 g) as a dispersing agent were dissolved in deionized water to form the other 100 mL homogeneous solution (Solution B). In order to adjust the pH value to 10.0~11.0, an

appropriate amount of NaOH was added to 50 mL deionized water (Solution C), which was transferred into a 500 mL flask as mother liquor firstly. Both Solution A and Solution B were added dropwise into this flask under stirring at 25 °C; while the pH value of reaction system was controlled at 10.0~11.0 through adding a little saturated NaOH solution. The liquid-phase reaction was kept on 30 min, and the pink precipitate as a $ZnCo_2O_4$ precursor was obtained. After washed by deionized water and dried in a vacuum oven at 40 °C, the precursor was thermally treated at 450 °C in air for 6 h with the ramp of 5 °C·min^{-1}. Subsequently, the black $ZnCo_2O_4$ sample was prepared. At last, NaOH (16.000 g) was dissolved in deionized water to form a 200 mL homogeneous solution (Solution D), which was transferred into a 400 mL beaker as the alkaline etchant. And all of the above $ZnCo_2O_4$ powders were added into the beaker under stirring at 25 °C, and the etching reaction was kept on 1 h. After washed by deionized water and dried in a vacuum oven at 40 °C for 24 h, the defective $ZnCo_2O_4$ sample was synthesized.

Materials Characterization. The elemental analysis was characterized by inductively coupled plasma atomic emission spectroscopy (ICP-AES) (Baird, PS-6). The thermal behavior was analyzed by thermogravimetric analysis (TGA) (Netzsch, 409PC) in air atmosphere with temperature range from 30 °C to 800 °C at ramp rate of 5 °C·min^{-1}. The XRD patterns were recorded by X-ray powder diffraction (XRD) (Rigaku, D/Max-2000). The nitrogen adsorption data were calculated by Nitrogen adsorption data (Quantachrome, Autosorb-iQ). The X-ray photoelectron spectroscopy (XPS) characterization was carried out by an ESCALAB 250 instrument. The micro

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morphological and structural characterization was investigated by scanning electron microscope (SEM) (TESCAN, MIRA3 and JEOL), atomic force microscope (AFM) (Bruker, Dimension Icon) and transmission electron microscope (TEM) (FEI, Tecnai G2 F20).

Electrochemical Measurements. All of electrochemical measurements were similar to our previous work including the type of coin cell, the composition of working electrode, the weight of active material, the electrolyte and so on [30-32]. The galvanostatic cycling test was conducted at 25 °C under different rates by battery chargers (Land, CT2001A). Using the electrochemical workstation (Metrohm, Autolab), the cyclic voltammogram (CV) test was performed under the scan rate of 0.1 mV·s^{-1} between 0.01 V and 3.00 V ($vs.$ Li$^+$/Li), and the electrochemical impedance spectroscopy (EIS) was carried out at AC voltage of 5 mV amplitude under the open-circuit condition over a frequency range of 1.00 MHz to 0.01 Hz.

3. Results and Discussion

The as-prepared $ZnCo_2O_4$ precursor by simple liquid-phase co-precipitation is studied by XRD to confirm the crystalline phase [Eq. (1)]. As shown in Fig. S1 (*Supporting Information*), the main diffraction peaks correspond to the mixture of $Co(OH)_2$ (JCPDS no.30-0443) and $Zn(OH)_2$ (JCPDS no.20-1437). And the typical morphology of precursor is irregular nano-particles shown in the Fig. S2 (*Supporting Information*). In order to gain the best heating treatment temperature of $ZnCo_2O_4$, the thermal behavior of precursor is characterized by TGA at air atmosphere shown in Fig. 1. The TGA curve exhibits two distinct weight loss steps. The first weight loss

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between 30 °C and 146 °C is ~ 2.0%, which may be assigned to the evaporation of

moisture. And the second weight loss between 146 °C and 398 °C is ~ 12.3%, which

may be attributed to the transformation of $M(OH)_2$ (M = Co and Zn) into $ZnCo_2O_4$

[Eq. (2)]. After 398 °C, almost no weight loss appears. Hence, in the corresponding

$ZnCo_2O_4$ synthesis, the thermal treatment temperature is controlled at 450 °C.
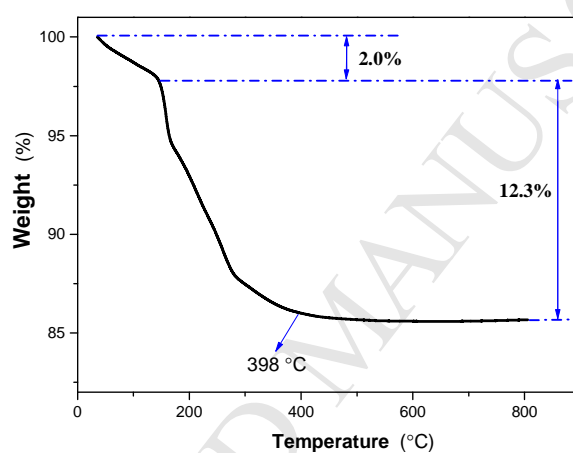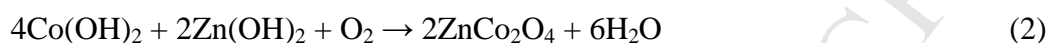
$$M^{2+} + 2OH^- \rightarrow M(OH)_2 \downarrow \qquad (M = Co \text{ and } Zn) \qquad (1)$$

$$4Co(OH)_2 + 2Zn(OH)_2 + O_2 \rightarrow 2ZnCo_2O_4 + 6H_2O \qquad (2)$$

Fig. 1 TGA curve of as-prepared $ZnCo_2O_4$ precursor under air atmosphere (5 °C·min^{-1})

Generally, in strong alkaline solution, Zn^{2+} in the solid-phase could react with OH^-

to form soluble $Zn(OH)_4^{2-}$ [33,34]. Base on the mechanism, a handful of Zn-ions in

$ZnCo_2O_4$ particles are etched through this alkaline-tailored method at low temperature,

so the pure $ZnCo_2O_4$ is transformed into the defective $ZnCo_2O_4$ with Zn vacancies.

Measured by ICP-AES, the Mole ratio of Zn and Co is 0.95:2.00 in the defective

$ZnCo_2O_4$, so its molecular formula can be recorded as $Zn_{0.95}Co_2O_4$. In addition, Na,

Fe, Ca and Mg contents of the $Zn_{0.95}Co_2O_4$ sample are 58 ppm, 13 ppm, 22 ppm and

18 ppm, respectively. And Fig. 2 shows the phase purity of $ZnCo_2O_4$ powders before

and after etched. Both of the diffraction peaks match well with the spinel $ZnCo_2O_4$

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(JCPDS no.23-1390), which locate at 2 theta of 18.94°, 31.25°, 36.82°, 38.60°, 44.81°, 55.50°, 59.28°, 65.23° and 77.30° corresponding to (111), (220), (311), (222), (400), (422), (511), (440) and (533), respectively. Hence, it indicates that the etch treatment does not affect the crystal structure of $ZnCo_2O_4$. In addition, The BET specific surface area of $ZnCo_2O_4$ is measured to 105.07 $m^2 \cdot g^{-1}$; after etched, the corresponding specific surface area of defective $ZnCo_2O_4$ increases to 170.12 $m^2 \cdot g^{-1}$, which is probably caused by the additional interface due to Zn vacancies in nano-particles (Fig. 3).

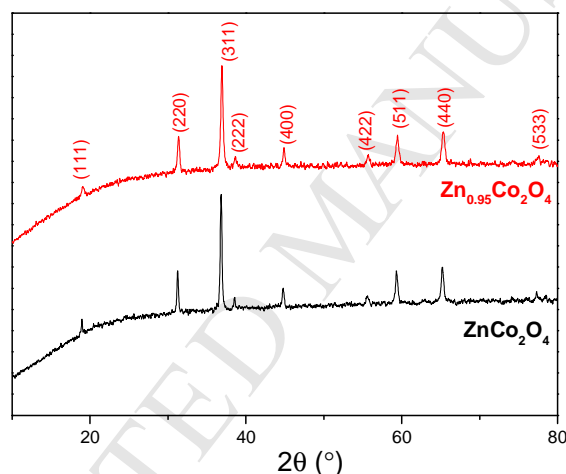**Fig. 2** XRD patterns of $ZnCo_2O_4$ and $Zn_{0.95}Co_2O_4$

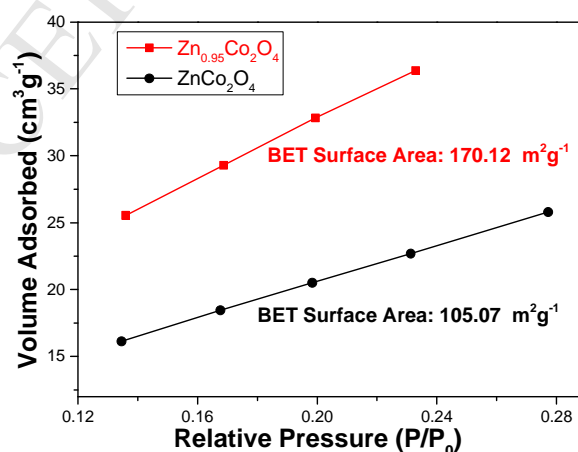**Fig. 3** N_2 adsorption isotherms of $ZnCo_2O_4$ and $Zn_{0.95}Co_2O_4$

Morphology and micro-structure of as-synthesized samples are characterized by

the SEM technique in Fig. 4. Figs. 4a and 4b clearly display the typical irregular lamellar structure of $ZnCo_2O_4$ with good dispersion, and the particle size of $ZnCo_2O_4$ nanosheets is 50-200 nm. Figs. 4c shows that the morphology and particle size of $Zn_{0.95}Co_2O_4$ keep similar to the above $ZnCo_2O_4$ powders, which proves that the etch treatment does not damage the micro-structure of nano-particles. The higher-magnification SEM image (Fig.4d) elucidates that the $Zn_{0.95}Co_2O_4$ nanosheets pile up obviously, and the phenomenon of agglomeration is more serious than the $ZnCo_2O_4$ powders', which may be attributed to the additional specific surface energy by the etch treatment. Further information about the $Zn_{0.95}Co_2O_4$ nanosheets is obtained form TEM and high-resolution TEM (HRTEM) images. Nanosheets with diameter of 50-200 nm touched each other is clearly observed again in Fig. 5a. In Figs. 5b-5d, all of the lattice fringes are uniform, and the corresponding interplanar spacings of 0.24 nm, 0.29 nm and 0.20 nm are indexed to the (311), (220) and (400) planes of spinel $ZnCo_2O_4$ phase, respectively. Remarkably, the thin layered structure of $Zn_{0.95}Co_2O_4$ particles are further confirmed by the AFM image (Fig. 6); and the thickness of $Zn_{0.95}Co_2O_4$ nanosheets is measured to 10-12 nm. What is more, in order to understand the surface electronic states of $Zn_{0.95}Co_2O_4$ particles, the XPS analysis is also carried out in Fig. 7. The full wide-scan spectrum shows the presence of Zn, Co, O and C in Fig. 7a (C 1s line is used as the reference for calibration). And the two peaks at the binding energies of 1020.4 eV and 1043.5 eV are associated with Zn $2p_{3/2}$ and Zn $2p_{1/2}$ of Zn^{2+} in Fig. 7b; while the major peaks at 779.6 eV and 794.5 eV with a spin-orbit splitting of 15.1 eV are attributed to Co $2p_{3/2}$ and Co $2p_{1/2}$ in Fig. 7c,

confirming the presence of Co^{2+} and Co^{3+} [35-37]. The XPS results demonstrate that the surface Zn and Co cation valences of $Zn_{0.95}Co_2O_4$ after etched are the same as the traditional $ZnCo_2O_4$ [16-19].

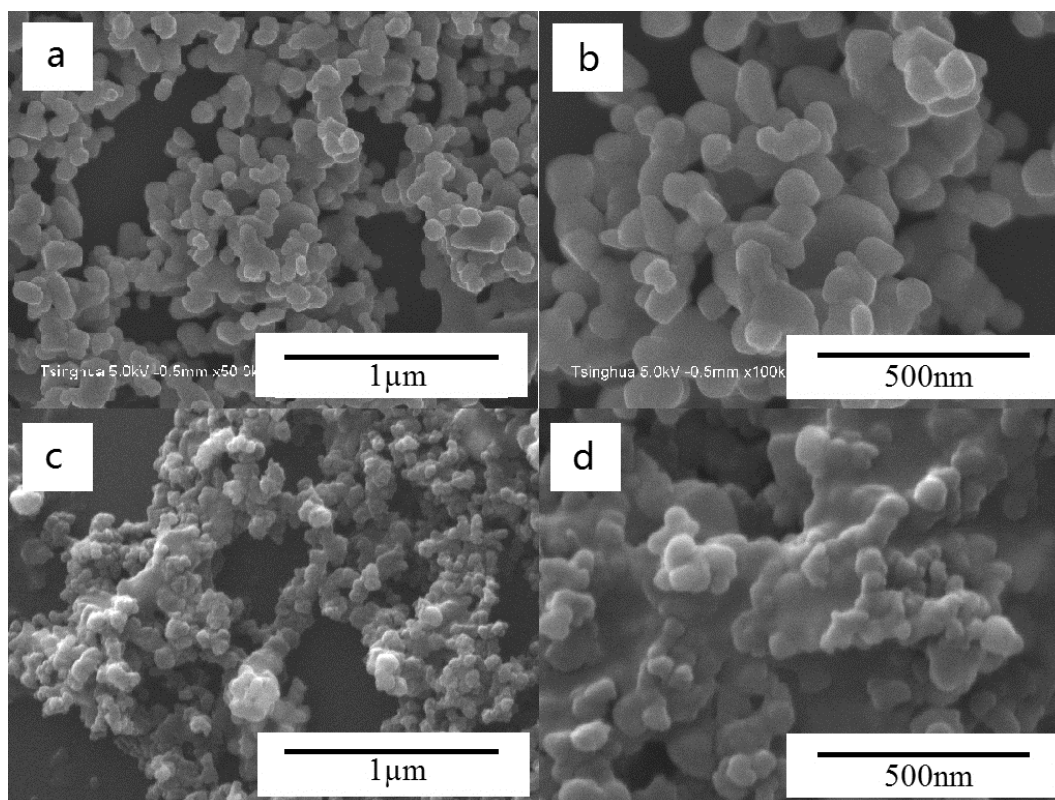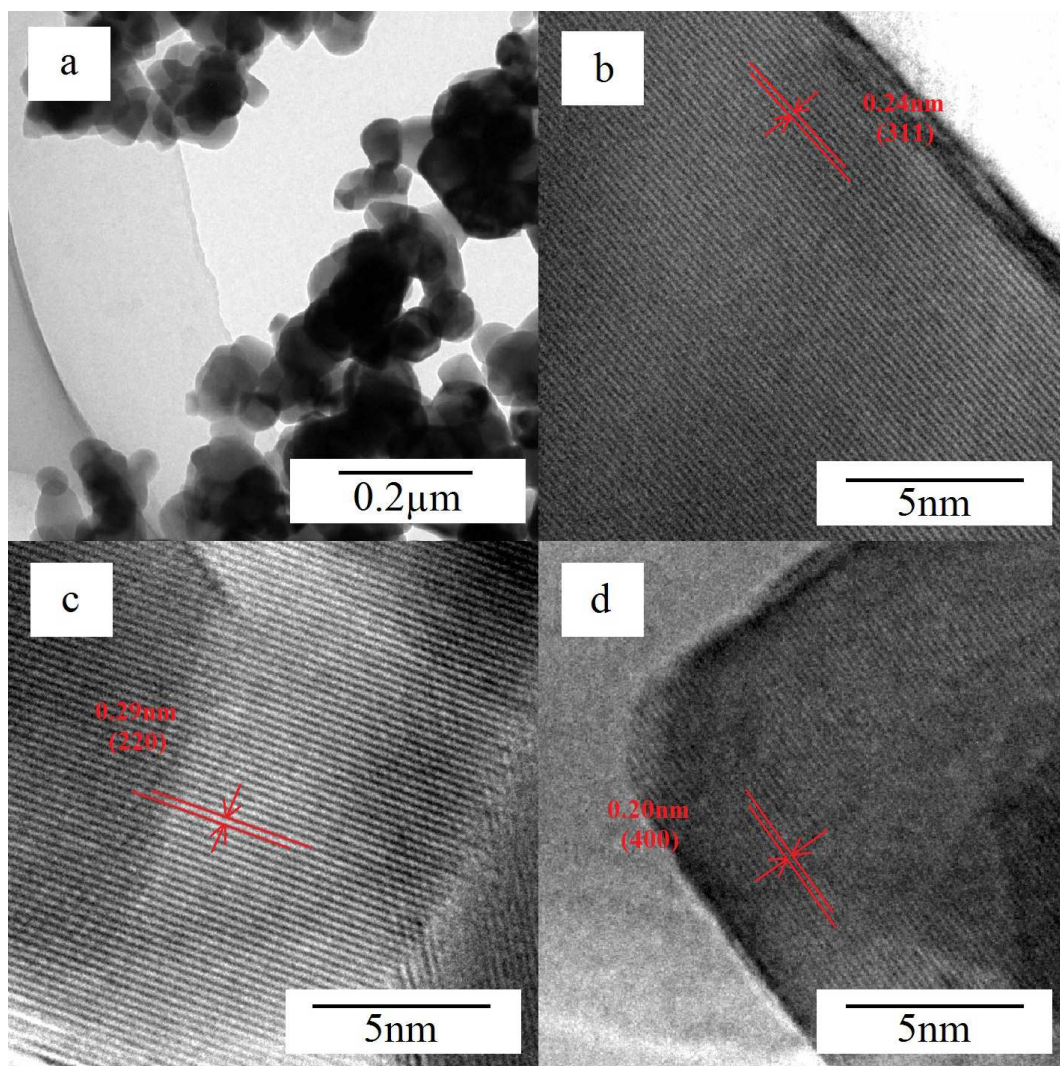**Fig. 4** (a), (b) SEM images of $ZnCo_2O_4$ powders and (c), (d) SEM images of $Zn_{0.95}Co_2O_4$ powders

Fig. 5 (a) TEM and (b)-(d) HRTEM images of $Zn_{0.95}Co_2O_4$ powders
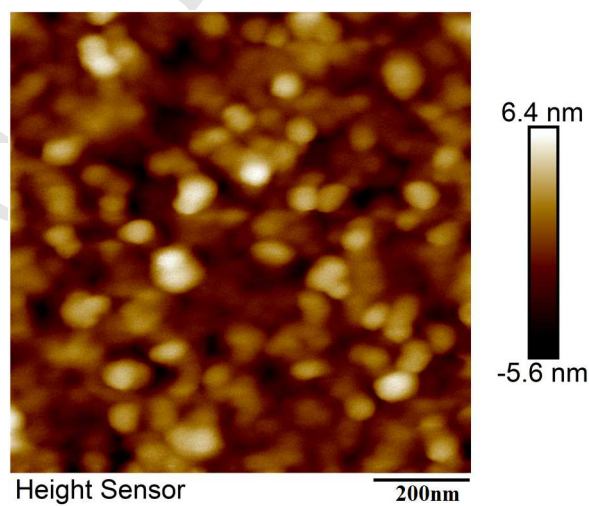
Fig. 6 AFM image of $Zn_{0.95}Co_2O_4$ powders
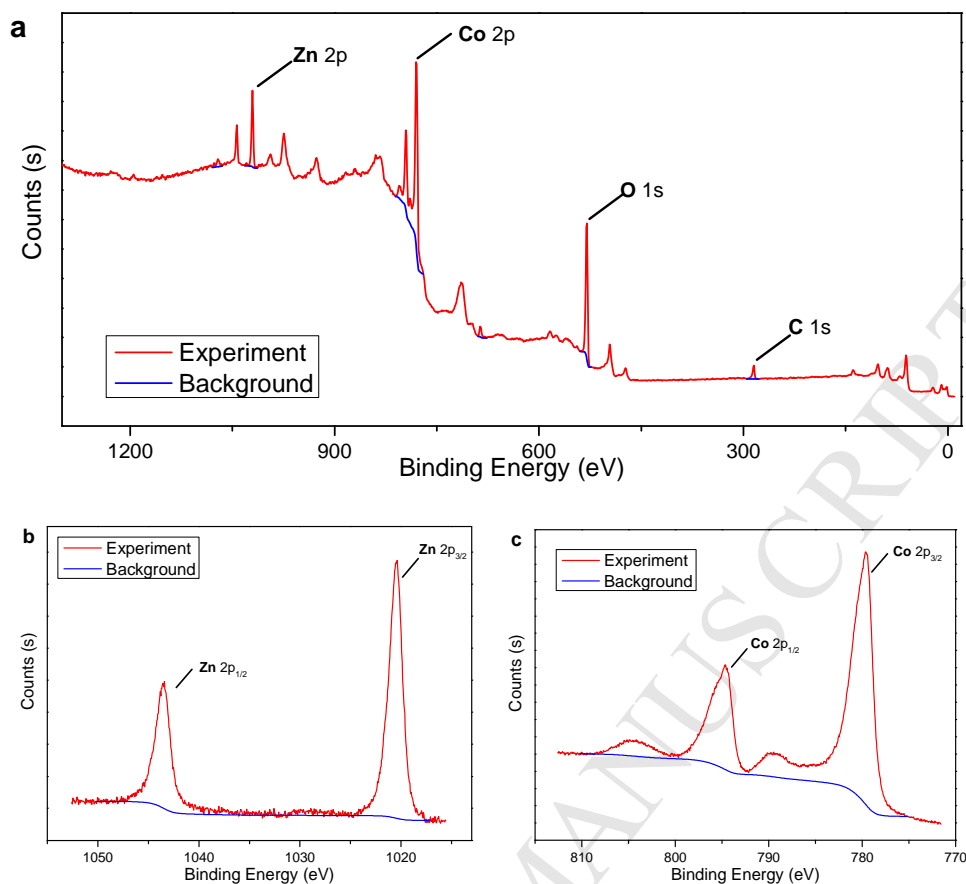
Fig. 7 XPS spectra of $Zn_{0.95}Co_2O_4$

To further understand the reaction mechanism of $Zn_{0.95}Co_2O_4$ electrode, the CV curves between 0.01 V and 3.00 V (*vs*. Li^+/Li) at 0.1 $mV \cdot s^{-1}$ is conducted for the initial five cycles in Fig. 8. In the 1st cathodic sweep, an well-defined reduction peak is recorded at ~ 0.78 V, which may be attributed to the intercalation of Li-ions into the spinel structure and the reduction of $Zn_{0.95}Co_2O_4$ to Zn and Co (Mole ratio of Zn:Co = 0.95:2) [Eq. (3)]; and a broad anodic peak is observed at ~ 2.15 V due to the oxidations of Zn/defective ZnO and Co/CoO [Eq. (4)] [38-41]. In the subsequent cathodic sweep, the reduction peak is moved to ~ 1.00 V and becomes broader than the 1st cycle, which belongs to the other reductions of defective ZnO/Zn and CoO/Co [Eq. (5)]. From the 2nd cycle onwards, the CV curves almost remain overlapped, which implies the high reversibility of electrochemical reactions and excellent kinetic

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of $Zn_{0.95}Co_2O_4$ for Li-ions insertion/extraction. In addition, compared to the traditional $ZnCo_2O_4$ [42,43], the peaks indexed to the reversible reaction of Zn and Li-Zn alloy is not observed in this research, so the Zn vacancies in defective $ZnCo_2O_4$ may decrease the probability of the by-reaction.

$$Zn_{0.95}Co_2O_4 + 8Li^+ + 8e^- \rightarrow 0.95Zn + 2Co + 4Li_2O \qquad (3)$$

$$0.95Zn + 2Co + 3Li_2O \rightarrow Zn_{0.95}O + 2CoO + 6Li^+ + 6e^- \qquad (4)$$

$$Zn_{0.95}O + 2CoO + 6Li^+ + 6e^- \rightarrow 0.95Zn + 2Co + 3Li_2O \qquad (5)$$
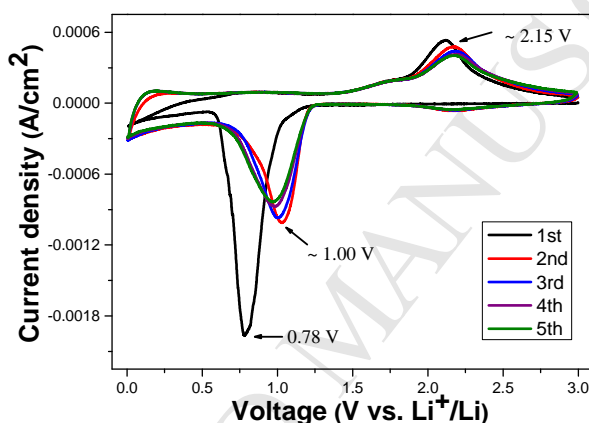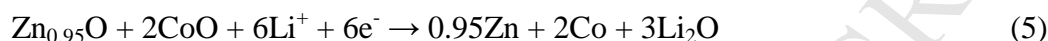
Fig. 8 First five cyclic voltammogram curves of $Zn_{0.95}Co_2O_4$ electrode at a scan rate of 0.1 $mV \cdot s^{-1}$ in the range of 0.01-3.00 V

The plots of specific capacity *vs.* voltage at 0.4 $A \cdot g^{-1}$ with different cycles are shown in Fig. 9. An evident long plateaus in every discharge curve exhibits at ~ 1.0 V, and an unobvious short plateaus in every charge curve presents between 1.9~2.2 V, which match well with the above CV results. The initial discharge and charge process are up to 1057.0 $mAh \cdot g^{-1}$ and 776.3 $mAh \cdot g^{-1}$, respectively, corresponding to a Coulombic efficiency (CE) of 73.4%. The irreversible capacity loss may come from the consumption of active materials in the formation of the SEI film and the irreversible reaction [Eq. (3)], which is reflected in the above CV profiles, too [44,45]. Subsequently, the 50th and 100th CEs are high to 99.5% and 98.6%, respectively,

demonstrating the super stability of defective $ZnCo_2O_4$ and the fast transport paths of

Li-ions insertion/extraction due to the Zn vacancies again.

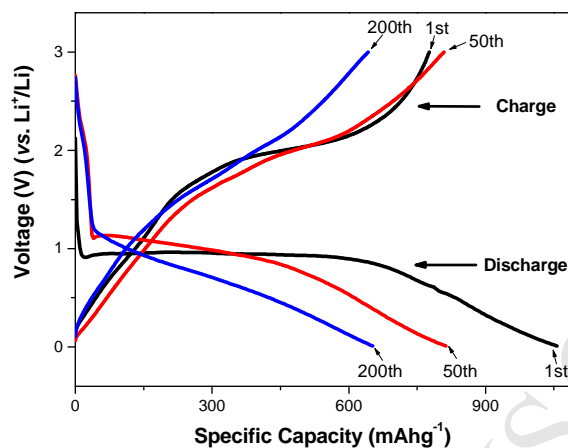**Fig. 9** Different discharge and charge profiles of $Zn_{0.95}Co_2O_4$ electrode in the range of 0.01-3.00 V at $0.4\,A\cdot g^{-1}$

The cycling performance of the $Zn_{0.95}Co_2O_4$ after etched and the pure $ZnCo_2O_4$

before etched as the reference is displayed in Fig. 10a. At the high rate of $0.4\,A\cdot g^{-1}$,

the discharge/charge capacity of $Zn_{0.95}Co_2O_4$ stays very stable and remains at a high

value between 805.0/800.0 $mAh\cdot g^{-1}$ and 795.0/785.0 $mAh\cdot g^{-1}$ from the 2nd cycle to

the 100th cycle; then, it begins to fade slowly to 652.2/642.5 $mAh\cdot g^{-1}$ at the 200th

cycle. In comparison, the discharge/charge capacities of pure $ZnCo_2O_4$ are lower than

the above $Zn_{0.95}Co_2O_4$'s at the same cycle obviously, which suggests that the Zn

vacancies may activate more reaction sites efficiently [46]. And the discharge/charge

capacity rises up gradually from 707.1/688.0 $mAh\cdot g^{-1}$ to 760.4/750.4 $mAh\cdot g^{-1}$ at the

2nd~27th cycle, corresponding to the traditional activation process of Li-ion diffusion

[47]. Further prolonged cycling to the 200th cycle, it decreases to 483.2/476.3

$mAh\cdot g^{-1}$ quickly. It is noted that the cycling stability of defective $ZnCo_2O_4$ with Zn

vacancies is substantially enhanced.

The rate capabilities are evaluated by a multiple-step galvanostatic strategy at various current densities of 0.6~6.0 $Ah \cdot g^{-1}$ in Fig. 10b. At 0.6 $Ah \cdot g^{-1}$ and 1.0 $Ah \cdot g^{-1}$, the property of $Zn_{0.95}Co_2O_4$ and $ZnCo_2O_4$ are similar. The former average discharge/charge capacities except the 1st cycle are 618.0/608.3 $mAh \cdot g^{-1}$ and 561.8/557.5 $mAh \cdot g^{-1}$; while the latter's capacities are 578.3/562.6 $mAh \cdot g^{-1}$ and 534.7/523.5 $mAh \cdot g^{-1}$. When the rate increases to 1.6 $Ah \cdot g^{-1}$, 2.0 $Ah \cdot g^{-1}$ and 4.0 $Ah \cdot g^{-1}$, the rate behavior of $Zn_{0.95}Co_2O_4$ illustrates much better than the $ZnCo_2O_4$'s. And the corresponding average discharge/charge capacities of $Zn_{0.95}Co_2O_4$ are 483.5/480.2 $mAh \cdot g^{-1}$ and 340.0/333.2 $mAh \cdot g^{-1}$, 158.2/152.6 $mAh \cdot g^{-1}$, respectively; in contrast, the latter's capacities are 377.6/362.4 $mAh \cdot g^{-1}$, 179.8/175.0 $mAh \cdot g^{-1}$ and 44.7/40.5 $mAh \cdot g^{-1}$, respectively. Even at the highest current density of 6.0 $Ah \cdot g^{-1}$, both of them may lose the cycling ability. Remarkably, when the rate finally returns to 1.0 $Ah \cdot g^{-1}$, 0.6 $Ah \cdot g^{-1}$ and 0.4 $Ah \cdot g^{-1}$ after many cycles at high rates, the average discharge capacities of $Zn_{0.95}Co_2O_4$ are recovered to 594.9/592.8 $mAh \cdot g^{-1}$, 654.2/645.5 $mAh \cdot g^{-1}$ and 748.9/742.1 $mAh \cdot g^{-1}$, respectively; however, the corresponding latter's capacities are just restored to 388.4/380.2 $mAh \cdot g^{-1}$, 475.1/462.2 $mAh \cdot g^{-1}$ and 543.0/532.2 $mAh \cdot g^{-1}$, respectively. Hence, compared to pure $ZnCo_2O_4$, the $Zn_{0.95}Co_2O_4$ with larger interface area and more Li^+ insertion/extraction sites may facilitate faster Li-ion transfer between electrolyte and electrode.
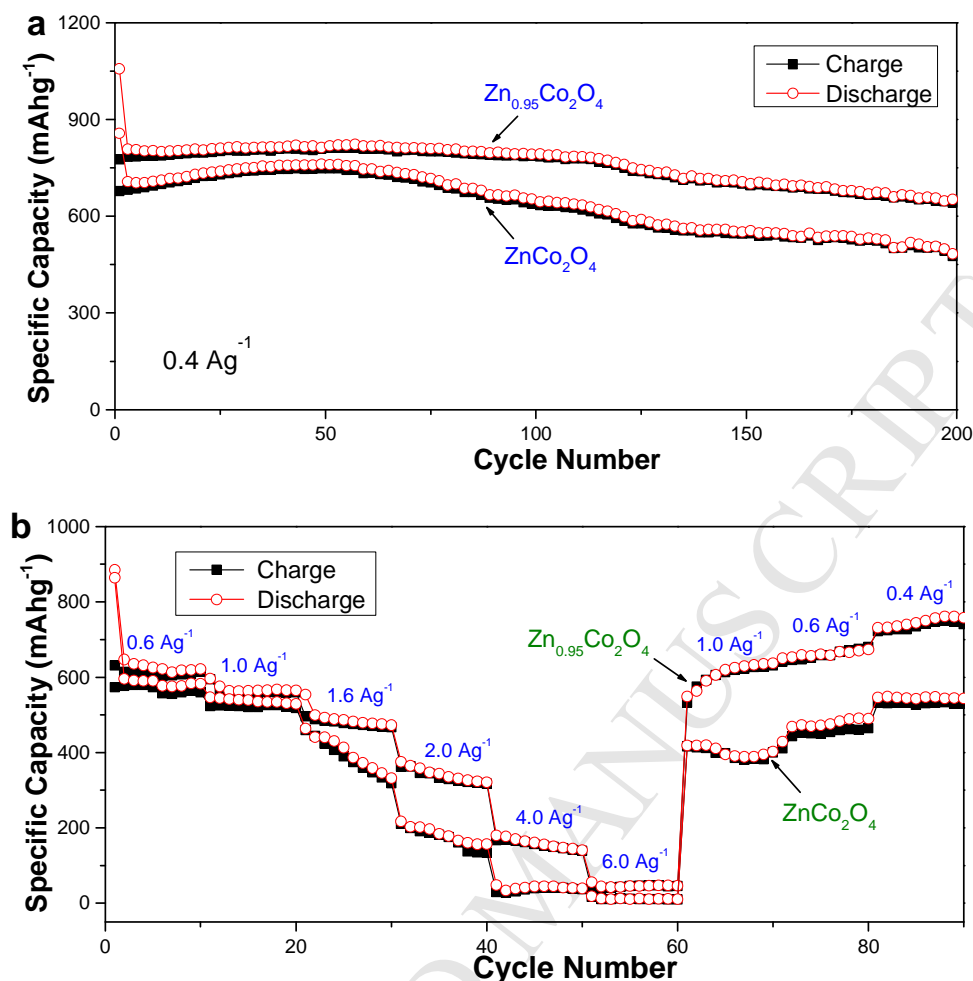
Fig. 10 (a) Cycling performance of $Zn_{0.95}Co_2O_4$ and $ZnCo_2O_4$ electrodes at 0.4 A·g^{-1} and (b) rate performance of $Zn_{0.95}Co_2O_4$ electrode at different current densities

The EIS curves are tested to confirm the different kinetic process of $Zn_{0.95}Co_2O_4$ and $ZnCo_2O_4$ after 200 cycles at 0.4 A·g^{-1} (Fig. 11). Both of the Nyquist plots are similar in shape except the diameter of approximate semi-circle in the high-frequency region. The diameter of $Zn_{0.95}Co_2O_4$'s is much smaller than the $ZnCo_2O_4$'s, which suggests that the electron conductivity of the former is still better than the latter's after 200 cycles [48]. In addition, the $Zn_{0.95}Co_2O_4$ electrode (the mixture of $Zn_{0.95}Co_2O_4$/ATB/PVDF) under 0.4 A·g^{-1} after 200 cycles is also shown in Fig. 12. The SEM images clearly illustrate that the micro-structure of nanosheets is well maintained without collapse and shedding, so the excellent structural stability agrees
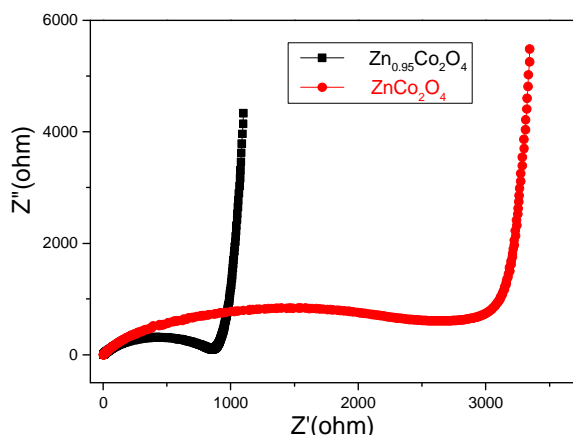
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well with the above advanced capacity retention.

Fig. 11 Nyquist plots of $Zn_{0.95}Co_2O_4$ and $ZnCo_2O_4$ electrodes under 0.4 Ag^{-1} after 200 cycles measured with an amplitude of 10 mV over the frequency range of 1 MHz and 0.01 Hz
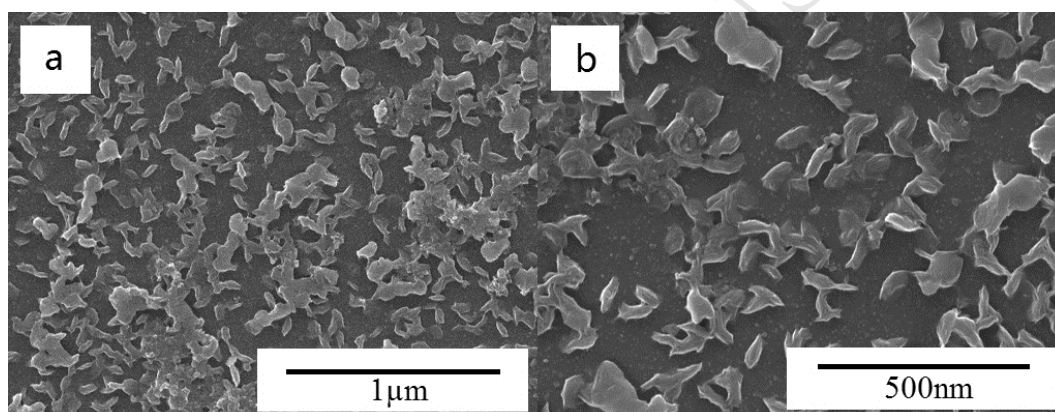
Fig. 12 SEM images of the $Zn_{0.95}Co_2O_4$ electrode (the mixture of $Zn_{0.95}Co_2O_4$/ATB/PVDF) under 0.4 Ag^{-1} after 200 cycles

Conclusions

In summary, the defective $ZnCo_2O_4$ with Zn vacancies ($Zn_{0.95}Co_2O_4$) via the liquid-phase co-precipitation and alkaline-tailored method has been synthesized successfully. The $Zn_{0.95}Co_2O_4$ powders are similar to the crystal phase, morphology, and particle size of the pure $ZnCo_2O_4$ before etched except the enlarged BET specific surface. The Zn vacancies in defective $ZnCo_2O_4$ may decrease the probability of the reversible by-reaction between Zn and Li-Zn alloy through the CV measurement. Compared to the traditional $ZnCo_2O_4$, the Zn vacancies in the defective $ZnCo_2O_4$ can

provide larger interface, activate more reaction sites and expand faster transport paths for both of Li-ions and electronics insertion/extraction, so the electrochemical performance of defective $ZnCo_2O_4$ has been enhanced highly. The discharge capacity retains at 652.2 mAh·g^{-1} under 0.4 A·g^{-1} after 200 cycles. When the rate returns to 0.4 Ah·g^{-1}, the average discharge capacity is recovered to 748.9 mAh·g^{-1} under the multiple-step high rates after many cycles. Furthermore, this facile strategy may be extended to design and construct other electrode materials with atomic vacancies for conversion devices.

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Notes

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Highlights

1. The defective $ZnCo_2O_4$ with Zn vacancies has been synthesized via the alkaline-tailored method.

2. The Zn vacancies may decrease the probability of the reversible by-reaction between Zn and Li-Zn alloy.

3. The discharge capacity retains at 652.2 mAh·g^{-1} under 0.4 A·g^{-1} after 200 cycles.

4. When the rate returns to 0.4 Ah·g^{-1}, the average discharge capacity is recovered to 748.9 mAh·g^{-1} under multiple-step high rates.