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Published in:
Sustainable Energy and Fuels

Link to article, DOI:
[10.1039/c9se00799g](https://doi.org/10.1039/c9se00799g)

Publication date:
2020

Document Version
Peer reviewed version

[Link back to DTU Orbit](#)

Citation (APA):
McEnaney, J. M., Rohr, B. A., Nielander, A. C., Singh, A. R., King, L. A., Nørskov, J. K., & Jaramillo, T. F. (2020). A cyclic electrochemical strategy to produce acetylene from CO₂, CH₄, or alternative carbon sources. *Sustainable Energy and Fuels*, 4(6), 2752-2759. <https://doi.org/10.1039/c9se00799g>

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Title: A Cyclic Electrochemical Strategy to Produce Acetylene from CO₂, CH₄, or Alternative Carbon Sources

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Abstract

Electrochemical transformation of potent greenhouse gases such as CO₂ and CH₄ to produce useful carbon-based products is a highly desirable sustainability goal. However, selectivity challenges remain in aqueous electrochemical processes as selective CO₂ reduction to desired products is difficult and electrochemical CH₄ oxidation often proceeds at very low rates. The formation of C-C coupled products in these fields is particularly desirable as this provides a path for the production of high-value fuels and chemicals. We have developed a cyclic electrochemical strategy which can produce acetylene, a C-C coupled product, from such carbon sources and water, with favorable current density and selectivity. This strategy is exemplified with a lithium-mediated cycle: an active Li⁰ surface is electrochemically generated from LiOH, the newly formed Li⁰ reacts with a carbon source to form Li₂C₂, and Li₂C₂ is hydrolyzed to form acetylene and regenerate LiOH. We demonstrate this process primarily using CO₂ gas, achieving a current efficiency of 15% to acetylene (which represents 82% of the maximum based on stoichiometric production of oxygenated byproducts, e.g. LiCO₃ and/or Li₂O), as verified by gas chromatography and Fourier transform infrared radiation studies. We also explore CH₄, CO, and C as alternative precursors in the acetylene synthesis. Notably, the use of graphitic carbon at higher temperatures resulted in over 55% current efficiency to acetylene, with opportunity for further optimization. Importantly, this cycling method avoids the formation of common side products observed during aqueous electrochemical CO₂ and CH₄ redox reactions, such as H₂, CO, HCO₂⁻, or CO₂. Theoretical considerations elucidate the feasibility and general applicability of this cycle and the process steps have been characterized with specific electrochemical and materials chemistry techniques. The continued development of this strategy may lead to a viable route for the sustainable production of C-C coupled carbon fuels and chemicals.

Introduction

Carbon dioxide and methane are two major emitted chemicals with global impact as greenhouse gases. The limited chemical versatility and limited demand vs a massive supply for these gases also makes them inexpensive materials. Thus, these gasses are targeted as precursors for chemical processing to produce higher value, more useful products. Aqueous electrochemical CO₂ reduction or CH₄ oxidation are attractive routes toward sustainably

converting these chemicals into higher value chemicals and fuels. However, under such aqueous conditions, CO₂ reduction is often dominated by the hydrogen evolution reaction (HER), motivating research into non-aqueous routes.¹⁻³ On the other hand, CH₄ is difficult to selectively oxidize on surfaces and often fully oxidizes to CO₂, if it reacts at all.⁴ In order to form higher value chemical products such as C₂+ alcohols and long chain hydrocarbons, however, C-C coupling capability is necessary, thus this is a key challenge in electrochemical production processes.⁵⁻⁶ For CO₂ electrochemical reduction, some progress has been made on copper based catalysts, one of very few systems capable of demonstrating C-C coupling and promoting selectivity to a range of higher value products.⁵⁻⁹ With very few electrochemical routes to form C-C coupled products from these sources, new directions to the formation of many higher value fuels and chemicals are highly desirable.

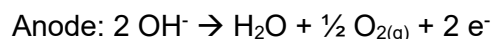
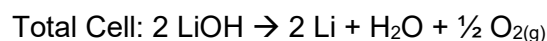
Acetylene, as a triple bonded, C-C coupled chemical and fuel, is an important product with exceptionally diverse chemistry that can serve as a chemical precursor to polymers, fuels, coatings, solvents, adhesives, and other carbon-based products.¹⁰ An elegant approach to low-carbon footprint chemical processes would be to produce acetylene as a platform chemical, employing renewable electricity and atmospheric greenhouse gasses (GHGs), e.g. CO₂ and CH₄, as the feedstocks. In fact, long ago, acetylene was used to produce a broad range of carbon-based chemicals before the advent of fossil fuel refinery technologies. The acetylene itself was traditionally produced by thermally combining calcium oxides with carbon but required very high temperatures (~2200°C) to form calcium carbide as an acetylene precursor. Nowadays, acetylene is primarily produced by partial methane combustion.¹⁰⁻¹¹ There is an opportunity to develop pathways to acetylene production in a low carbon-footprint manner.

We have developed a lithium-mediated electrochemical cycling strategy as an alternative pathway for the production of acetylene from H₂O and a variety of low-cost carbon feedstocks including CO₂ and CH₄. Similar to our previous work in which we produced NH₃ from N₂ and H₂O (achieving ~90% current efficiency),¹² the lithium-mediated production process comprises three distinct steps: 1) electrochemical reactive surface preparation, 2) carbon reactant activation, and 3) hydrolysis to synthesize acetylene. These steps can be cycled for continuous acetylene production (Figure 1). Herein, we primarily focus on CO₂ reduction to acetylene via lithium carbide as a baseline for this strategy. This electrochemical strategy circumvents the HER for CO₂ reduction, and for CH₄ oxidation, prevents the complete oxidation to CO₂ while still using water as a proton source. As one of very few electrochemical routes to producing C-C coupled products, this strategy represents an important direction for carbon material electrochemical synthesis with opportunities for improved efficiency.

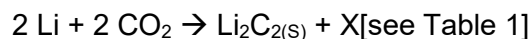
Results

The cyclic electrochemical strategy for acetylene synthesis is conceptually outlined in Figure 1. The three reaction steps shown in the figure are generalized as follows:

Step 1: LiOH electrolysis:



Step 2: Direct reaction of metallic Li with CO₂ to form Li₂C₂:



Step 3: Hydrolytic release of C₂H₂ by reaction with H₂O



The electrochemical nature of the metal electrolysis step, which is likely the major energy input in this reaction with the remaining reactions thermodynamically downhill, allows the cycle to be coupled directly to sustainable electricity sources.

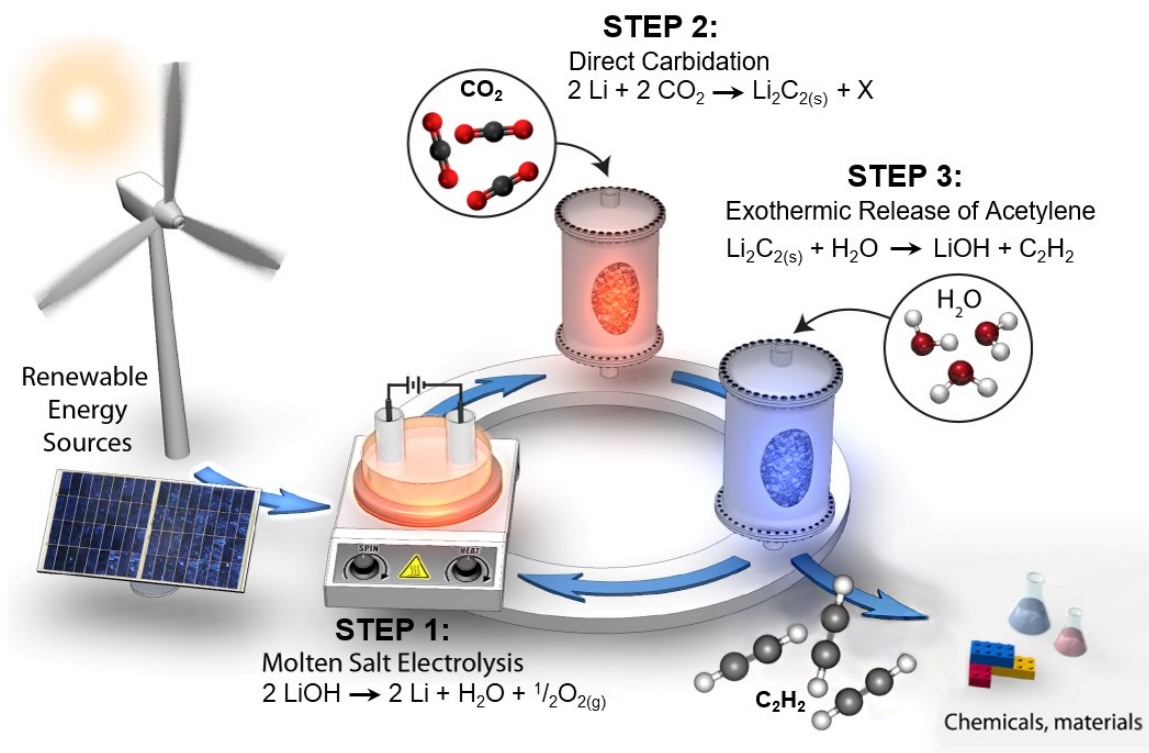


Fig. 1. Electrochemical concept cycle for acetylene production. Parallel concept cycles exist for other carbon sources such as CH₄, CO, and C for this strategy. This figure is adapted from Ref. 11 with permission from the Royal Society of Chemistry.¹²

We have experimentally demonstrated the Li-based electrochemical cycle to produce acetylene using CO₂. The first step of this electrochemical cycle was reported and discussed in our previous work on electrochemical ammonia synthesis.¹² Briefly, an LiCl-KCl/LiOH-LiCl (catholyte/anolyte) molten salt system held at 450°C was used to produce Li metal at the cathode, and O₂ and H₂O at the anode from LiOH. A constant current (~0.5 A cm⁻²) was applied for several hours to generate significant molten Li product. The electrolytically produced Li was allowed to cool, then Li was cut and pressed with steel into a thin (1 mm) pellet or slab and exposed to a flowing CO₂ atmosphere while being held at temperatures between 22 and 450°C (Step 2). We tested the conversion efficiency of Li to C₂H₂ by adding the product of the Li+CO₂ reaction to H₂O, which rapidly hydrolyzed the solid product (Step 3). Li to C₂H₂ yields are shown in Figure 2A for

the various temperature and conditions applied. We observed small yields of acetylene product after Li carbidation at 250°C and acetylene yields increased with increasing carbidation temperature up to 350°C (Figure 2A). Increasing carbidation reaction time from 4h to 8h at 350°C, resulted in an increase in acetylene yields (from ~4% to ~16%). Increasing temperature beyond 400°C resulted in a black product which did not react with water. The bulk composition of the Li+CO₂ products were examined by X-Ray diffraction (XRD) (Figure 2B). The low temperature representative Li+CO₂ product treated at 200°C, appeared nearly unreacted and gave only a signal for Li metal with preferred orientation for the (110) peak. A moderate temperature (350°C) representative sample was dark grey in color and non-lustrous, yielding broad peaks for Li₂O, and surprisingly no peaks for Li₂C₂. We assume the Li-C regions of the product must be amorphous to yield the significant amount of acetylene observed. The higher temperature (450°C) representative sample shows Li₂O peaks with greater intensity as well as the formation of LiCO₃.

To confirm the presence of C≡C, as would be expected in the Li₂C₂ structure, Raman spectroscopy was used on a moderate temperature (350°C, 8 h), CO₂ treated Li sample, showing a characteristic peak at 1873 cm⁻¹ for the bonding feature in this structure.¹³⁻¹⁴ This indicates the presence of the desired product (Li₂C₂) despite the lack of bulk organization to produce a prominent Li₂C₂ XRD signal (Figure 2C). Two regions are shown, highlighting the non-uniformity of the sample. In region 1, a broad peak stretches to lower wavenumbers, which has been observed previously in amorphous Li₂C₂ samples.¹⁴ In region 2, a small but clear peak is observed at 1873 cm⁻¹, in line with the control. The Li₂C₂ control sample was produced based on a reported procedure¹⁵ using a Li metal and graphitic carbon powder mixture heated to 350°C for 4 h then to 700°C for 4 h under Ar.

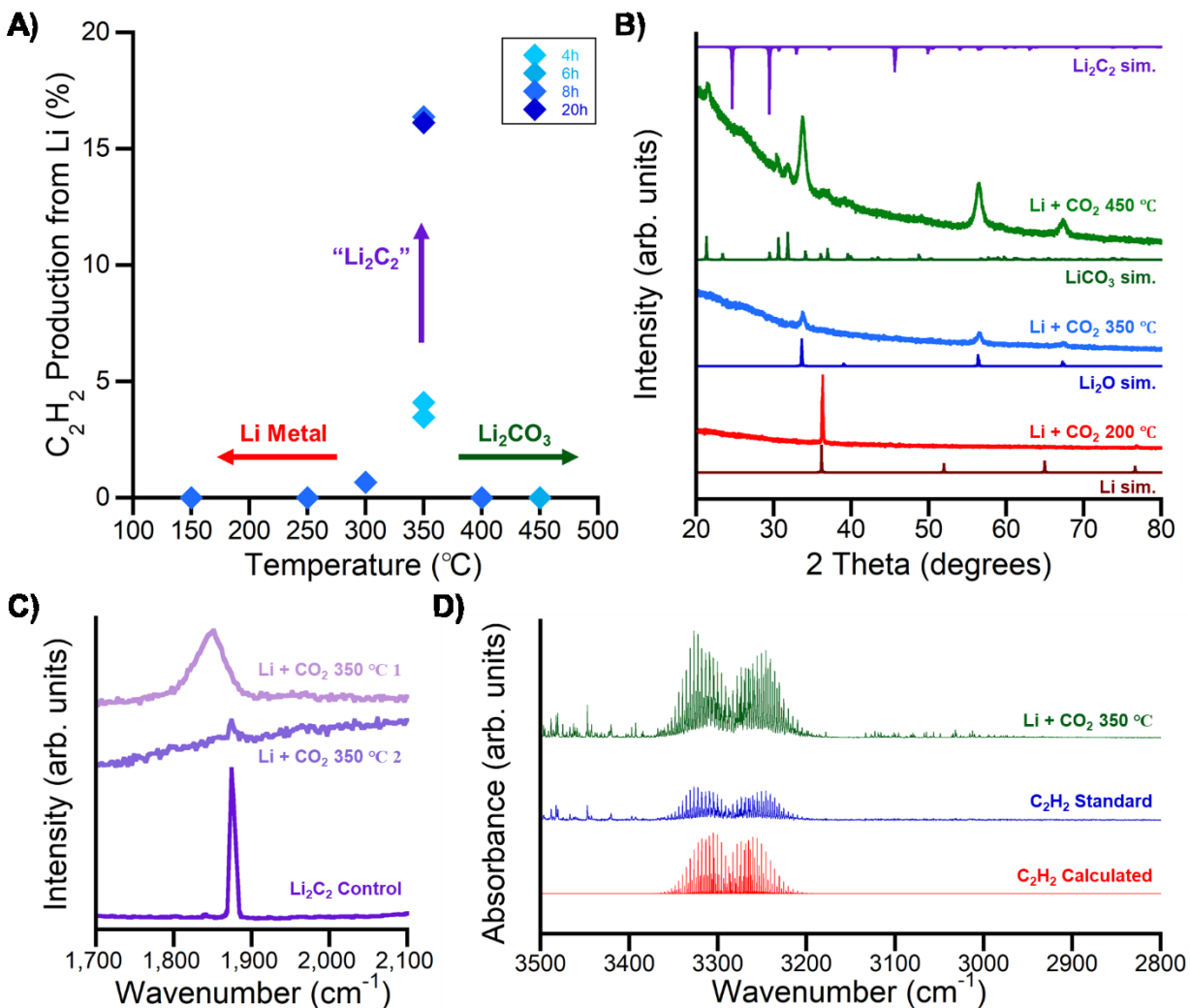


Fig. 2. (A) Acetylene yield percent based on the mols of Li used, obtained from Li carbidation from CO_2 under indicated reaction temperatures and times. (B) Representative XRD data for the experimentally synthesized $Li+CO_2$ products at given temperatures with relevant simulated (sim.) spectra overlaid. (C) Raman spectra, collected with an excitation laser at 432 nm, showing the characteristic peak¹³⁻¹⁴ at 1873 cm^{-1} for Li_2C_2 on two different surface regions of a 350 $^{\circ}C$ treated sample as well as for a Li_2C_2 synthesized control. (D) Representative FTIR signal for $Li-C$ hydrolysis product sample with an acetylene standard and a HITRAN calculated acetylene spectra¹⁶ for reference.

To detect and quantify the acetylene product, two complementary methods were employed. Gas chromatography (GC) was used to quantify our yields against a calibration curve made from serial dilutions of 1000 ppm and 100 ppm C_2H_2 gas standards. To verify acetylene production, with a chemically specific signal gas phase Fourier Transform Infrared Spectroscopy (FTIR) was used. A representative spectrum of the gas product is shown in Figure 2D (green) versus an acetylene standard gas (blue) and a HITRAN database calculated acetylene FTIR spectrum (red).

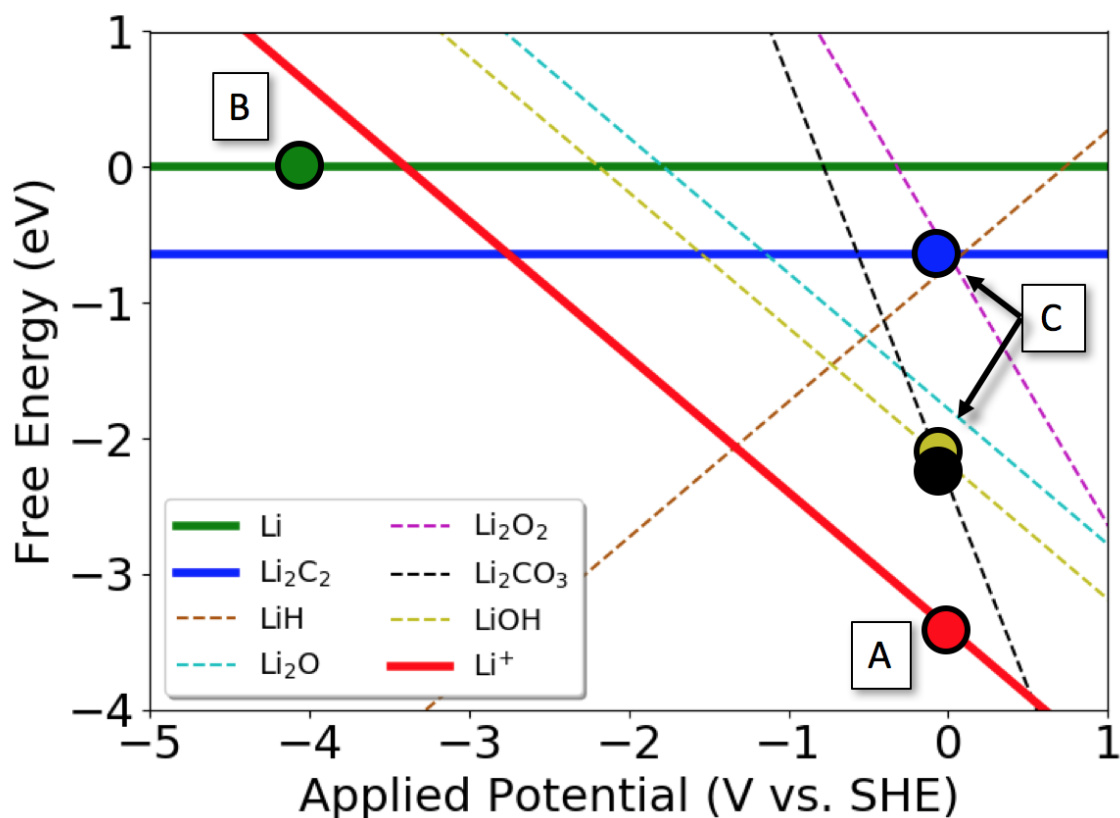


Fig. 3. Phase diagram showing the relative thermodynamic behavior of Li-C-O-H species as well as the cyclic process for electrochemical production of acetylene. In this figure, steps 1, 2, and 3 of our process are $A \rightarrow B$, $B \rightarrow C$, and $C \rightarrow A$, respectively. State A is lithium ions at 0 V vs. SHE. State B is lithium metal, which is produced by changing the applied potential to -4 V vs. SHE, and state C is a mixture of Li-C-O products, obtained by exposing Li to CO_2 .

The Pourbaix diagram and overlaid cycle process steps (Figure 3) shows that lithium carbide, the desired species to for acetylene, is not the most stable species at any applied potential. Under reducing conditions, the hydride is by far the most stable species, and under oxidizing conditions, the carbonate is by far the most stable species. This agrees well with the observed formation of LiCO_3 at high temperatures as a dominant thermodynamic product. We are, however, able to create a reasonable amount of the carbide, by only allowing the system to access certain species on the Pourbaix diagram during different steps of the cycle by controlling which elements are present at each step of the cycle. In the first step, plating lithium ions to form lithium metal at -4 V vs. SHE, the most stable species is lithium hydride, but no hydrogen is available at the cathode, thus lithium metal is formed with high selectivity. In the second step, the resulting hydride is exposed to carbon dioxide gas under no applied potential. Figure S1 shows a Pourbaix diagram containing only the species accessible in this step of the reaction. Lithium carbonate is by far the most stable species in this step, however, the only oxygen present in the system is the oxygen in carbon dioxide gas. Therefore, while a substantial amount of carbonate

is produced, any excess carbon is free to react with any excess lithium, and since this is a favorable reaction, we are able to produce a reasonable amount of lithium carbide (Table 1). Additionally, density functional theory (DFT) calculations show that the barrier for the formation of the carbide in bulk Li is much lower than the barrier for the formation of the carbonate in bulk (Figure S2). This agrees well with the observed data. At low temperatures, Li does not react quickly with the CO₂ at all, indicating that all barriers are too high to be quickly surmounted at low temperature. At high temperature, when the reaction is under thermodynamic control, LiCO₃ is the dominant product. At intermediate temperatures, some Li₂C₂ is formed. In the final step, water is added, providing the system with access to the Li⁺ line on the Pourbaix diagram, which is very stable at 0 V vs. SHE. In this final step of the cycle, water spontaneously reacts with the slab, converting the carbide and oxide lithium species to Li⁺ and OH⁻ ions, and the cycle is complete. Future work may aim to minimize the amount of carbonate formed in the reaction or increase its solubility in the electrolyte of step 1 to promote the cycle despite this thermodynamic sink.

To improve the yields of this process, there are several strategies to look to in future studies. For example, the ratio of carbon to other elements and type of elements is likely to be important. By using CO, rather than CO₂, one may limit the amount of LiCO₃ and Li₂O side products, increasing the Li-C formation and thus C₂H₂ product yield. This is evident by the reaction equations shown below as the relative amount of Li required per acetylene produced changes drastically depending on the side product made and the carbon source used:

Feedstock	Reasonable Mechanisms	Li _{(C₂H₂):Li_(total) Ratio}
CO ₂	6 CO ₂ + 10 Li → Li ₂ C ₂ + 4 Li ₂ CO ₃	1:5
	2 CO ₂ + 10 Li → Li ₂ C ₂ + 4 Li ₂ O	1:5
CH ₄	2 CH ₄ + 6 Li → Li ₂ C ₂ + 4 LiH	1:3
CO	3 CO + 4 Li → Li ₂ C ₂ + Li ₂ CO ₃	1:2
	2 CO + 6 Li → Li ₂ C ₂ + 2 Li ₂ O	1:3
C	2 C + 2 Li → Li ₂ C ₂	1:1

Table 1. List of possible carbon feedstocks, corresponding reasonable reaction mechanisms of these feedstocks with Li metal to produce Li₂C₂ and subsequent acetylene, and the resulting stoichiometric expectations for the ratio of Li atoms that go toward acetylene production.

In the first reaction equations case, using CO₂ and forming Li₂CO₃ or Li₂O as biproducts, only 1 in 5 reduced Li metal atoms is directly used for making acetylene (Table 1). Considering this, we are approaching the hypothetical maximum yield from CO₂ of 20% unless the oxygen atoms can be creatively scavenged (16.4% Li conversion to acetylene represents 82% of the maximum based on the stoichiometry presented in Table 1). By using CH₄ we change the byproduct to LiH, which will hydrolyze to form hydrogen and reform LiOH. Methane as a carbon source increases the ratio to 1 in 3 Li metal atoms toward making acetylene. In the third case, CO as a carbon source allows up to 1 in 2 Li metal atoms toward C₂H₂ (for an LiCO₃ byproduct, 1 in 3 if Li₂O is the byproduct). The ideal case for efficiency, reaction equation 4, is a process to convert Li metal and carbon directly into Li₂C₂, where 1 in 1 Li metal atoms could be used toward acetylene, however, carbon itself is typically a significantly more expensive reagent, resulting in a tradeoff. Each of these cases assumes that one is not losing gas in the form of O₂, CO, or H₂ (from CH₄),

nor forming solid carbon instead of the indicated products, which may occur to some extent on the Li reagent. Probing the reaction dynamics of product formation here will be important for further studies. We performed several preliminary tests using CH₄, CO, and C reagents. Each also successfully yielded acetylene, however, CH₄ can result in carbon coking of the Li surface, preventing further reaction, and limiting yields to ~3 % so far (see Figure S3 in the Supporting Information). This seemingly low yield from Li results in an over 2% current efficiency to acetylene. However, considering that aqueous electrochemical methane oxidation typically operate at low current densities (<1 mA/cm²) for primarily CO₂ and CO production, this alternative method implores further study.¹⁷ Solid oxide fuel cell methane dehydrogenation cells can produce some C₂ products but require significantly higher operating temperatures (~700-1000°C).¹⁷ Next, an initial attempt at using a 10% CO in argon source, at a temperature of 350°C and time of 8 h resulted in an ~11% acetylene yield from Li. Finally, by using a graphitic carbon reagent we were able to observe 66% acetylene production from Li (or 58% current efficiency, using the same 700°C high temperature treatment as in the Raman control sample). It is likely that the CH₄, CO, and C results can be significantly improved with further study and optimization.

One can imagine a number of future opportunities to improve yield. For example, by continuing to hold the Li product under a negative potential while applying the carbon source. The Pourbaix diagram (Figure 3) indicates that at -4 V, Li₂C₂ becomes the thermodynamically favored product over Li₂O or LiCO₃. In the case of CH₄, however, the LiH becomes more thermodynamically favored with more negative potential, which may limit efficiency. Interestingly, intentional formation of LiH may be beneficial as LiH is reported to react directly with C at a low temperature of only 400-500°C (rather than the 700°C reaction of C with Li in this paper) to form Li₂C₂ and C₂H₂.¹⁸ One may also improve yield if one can prevent coking, surface coating, or byproduct formation which may be possible with the addition of specific defects into the Li lattice, or possibly by seeding the Li with Li₂C₂ crystals to promote the formation of Li-C. Finally, lithium is not uniquely capable of these carbidation and hydrolysis reactions, and other metal carbides may be formed with larger windows of temperature and reactant opportunity to be selective for acetylene yield. Calcium, for example, is an obvious candidate for electrochemical cycling as calcium carbide was the basis for acetylene synthesis before fossil fuels became the primary source.¹⁰

Conclusions

In summary, we have introduced and demonstrated a unique approach for electrochemical production of C-C coupled acetylene using CO₂ and H₂O at atmospheric pressure. This was enabled by a lithium mediated cycling process and resulted in ~15% current efficiency to acetylene, based on Li produced and used (which represents 82% of the maximum based on stoichiometric production of oxygenated byproducts, e.g. LiCO₃ and/or Li₂O). We have also outlined the generalization of this method to use CH₄, CO, or other sources of carbon with opportunity for even greater selectivity to acetylene based on reactant stoichiometries and an understanding of the corresponding Pourbaix diagram. Preliminary testing with graphitic carbon as the carbon source increased current efficiency to over 55%. This electrochemical C-C coupling process can circumvent the competing HER reaction of typical water-based electrocatalysis, by physically and temporally separating the electrochemical activation and protonation steps of the acetylene synthesis, though it does have selectivity issues to overcome in the form of undesirable lithium side products. If the selectivity of this production process can be further improved, it may

lead to a potentially sustainable source of acetylene, which is an excellent precursor to many of the chemicals that fossil fuels provide for us today.

Experimental

Chemicals and materials

Lithium hydroxide [$\geq 98\%$, LiOH, powder, Sigma-Aldrich], lithium chloride [$\geq 99\%$, LiCl, anhydrous powder, Sigma-Aldrich], potassium chloride [$\geq 99\%$, KCl, anhydrous powder, Sigma-Aldrich], stainless steel foil [Fe : Cr : Ni; 70 : 19 : 11 wt%, 0.5 mm thick, Alfa Aesar], graphite rod, nickel rod, carbon dioxide [$\geq 99.999\%$, CO₂, gas, Airgas], methane [$\geq 99.999\%$, CH₄, gas, Praxair], DI Millipore water, porous alumina diffusion barrier tube [80 mm height \times 27 mm outer diameter (OD) \times 3 mm thickness, P-3-C material, 1.7 mm average pore diameter, single closed end tube, Coorstek], alumina round dish [35 mL, 25 mm height \times 50 mm OD \times 3 mm thick, AdValue Technology], cylindrical alumina crucible [5 mL, 26 mm height \times 20 mm OD \times 1 mm thick, 2 mm diameter side wall hole, AdValue Technology].

Electrolysis of LiOH

The first step of the electrochemical process was executed as described previously in our work on electrochemical ammonia synthesis.¹² (See Figure S4 in the Supporting Information for additional electrochemical characterization.) NOTE: lithium reacts violently with water and hot molten salts can be hazardous, thus only appropriately trained personnel with proper safety precautions and PPE should attempt these reactions. An aluminum oxide cylindrical crucible was used as the major cell container. This crucible was equipped with a steel foil cathode and graphite rod anode for electrolysis reactions. A small porous alumina container surrounded the anode to prevent undesirable side reactions by isolating the anodic and cathodic reactants and products. Li was produced by LiOH electrolysis in this cell by applying a constant current (-0.5 A) over a 1-4 h time period. Li product was collected in a solid alumina cylindrical crucible surrounding the cathode, with a small 2 mm hole bored out to maintain electrical contact. Further electrochemical characterization, product characterization, and experimental notes can be found in our previous work.

Li-C synthesis from Li and CO₂, CH₄, or C

Li was produced in excess from LiOH electrolysis and collected from the small crucible containing the steel working electrode and the molten salt mixture. The melt was transferred to a steel crucible and allowed to cool. Li metal was brought above its melting temperature (180°C) to isolate it from solid salts of the electrolyte, and the liquid product was decanted into a second steel crucible. Solid, cooled Li was cut and pressed between steel into ~ 1 mm thick pellets. The mass of these Li pellets was measured (approx. 0.01 g each) and the pellets were transferred from the Ar glovebox in a septum capped vial to a CO₂ or CH₄ purged tube furnace. Li was very briefly (15 s) exposed to air where the surface would slightly tarnish during transfer. Heat was applied under the carbonated gas flow for the desired duration (0-500°C, 0-20 hours). Resulting gray to black pellets were collected in scintillation vials and stored in the Ar glovebox briefly until hydrolysis testing. For use of C instead of a gas reactant, graphitic carbon powder was added to solid Li in a steel crucible in the Ar-filled glovebox, using a 2:1 C to Li ratio. The Li + graphitic

carbon was pressed into a pellet in the glovebox as above. In this case, Ar was used as the carrier gas in the tube furnace for step 2, to keep an inert atmosphere during this reaction, which was held at 350°C for 4 hours then at 700°C for 4 hours.

Acetylene synthesis from Li-C and H₂O

For the acetylene synthesis step, we hydrolyzed our Li-C samples using the following procedure. The Li-C pellet was removed from the glovebox in an 8 mL, capped glass vial. An inverted graduated cylinder was equipped with a septum capped top such that gas samples could be collected from the septum. The graduated cylinder and large beaker were filled with <Millipore pure de-ionized water> and the sample vial containing the Li-C pellet was uncapped directly into the submerged graduated cylinder. Thus, all gases evolved from the pellet sample, including the gas from the sample vial, were collected in the graduated cylinder for detection and quantification of gas species. Caution: This hydrolysis reaction is exothermic in general, and possible unreacted Li (and LiH in the CH₄ case) can rapidly evolve hydrogen gas, thus extra caution, preparation, and appropriate PPE should always be used. To demonstrate cyclability, we have dried the oxidized lithium products (hydroxide +/- carbonate species) and re-dissolved them in the molten anolyte with successful continued production of Li. We note, however that the solubility of carbonate in the electrolyte is limited and could become problematic in cases where carbonate is produced.

Acetylene detection and quantification

The acetylene yield was determined by the concentration signal from a gas chromatograph (GC) with flame ionization detection (FID) and the total volume of evolved product gas, considering the average volume of non-product gas from the sample vial. The sample gas (100 µL) was injected into the GC. The sample peak was quantified against a calibration curve using 1000 ppm acetylene gas and a series of dilutions using the same septum capped graduated cylinder reaction apparatus described above. While the elution time for gas samples matched that of the acetylene standard, exact chemical signal verification was obtained using a gas phase FTIR cell. For this, the gas cell was first evacuated with the equipped house vacuum. The remaining sample gas (using up to 100 mL) was added via syringe directly to the gas phase FTIR cell. Resulting spectra gave clear signals for acetylene when present and were used to double check relative expected signal strength for quantity verification.

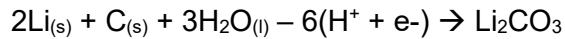
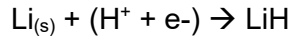
Current efficiency calculation

$$\begin{aligned} CE(\%) &= \frac{\text{Coulombic equivalent of } C_2H_2 \text{ produced}}{\text{Coulombic equivalent of the maximum possible } C_2H_2 \text{ that could have been produced}} \\ &= \frac{\text{Charge (C) required to form Li (measured by mass of Li)}}{\text{Total charge (C) passed during chronopotentiometry}} \\ &\quad \times \frac{\text{Measured production of } C_2H_2 \text{ (g)}}{\text{Expected production of } C_2H_2 \text{ (g) if all Li reacted to form } Li_2C_2} \times 100 \end{aligned}$$

We calculated current efficiency using the baseline average of 88.5% current efficiency to electrochemically produce lithium, and then multiplied that by the ratio of the measured production of acetylene gas over the maximum possible gas production if all Li had reacted to form Li_2C_2 upon exposure to the carbon source. We refer to that yield multiplied by 100 as the “% yield” or “acetylene production from Li.” Finally, due to the stoichiometric limitations addressed in Table 1 and the discussion we can also consider what percent of the maximum stoichiometric yield was obtained, assuming a given reaction mechanism from the table.

Porbaix diagram calculations

Calculations for the Pourbaix diagram are analogous with our recent work on this Li cycling strategy for ammonia synthesis.¹² For Figure 3, the carbon reference is bulk graphite, the lithium reference is bulk lithium in its most stable crystal structure (BCC), the oxygen reference is $\text{H}_2\text{O}_{(g)} - \text{H}_2$, and the hydrogen reference is the standard hydrogen electrode, in which H_2 gas is in equilibrium with protons and electrons. The slope of each line on the Pourbaix diagram is equal to the coefficient on the proton-electron pairs in the chemical reaction describing the formation of the corresponding species from the reference states normalized to the number of lithium atoms in the formula unit. For example, the slope of the lithium hydride line is equal to 1, and the slope of the line corresponding to lithium carbonate is equal to -3, as can be easily determined from the chemical equations describing their formation from the reference states:



Explicitly, for each species, the slope is given by

$$\text{slope} = N_H - N_{\text{Li}^+} - 2N_O$$

The intercept for each line is the free energy of formation of the corresponding species under ambient conditions (300K, 1 atm). Then enthalpy of formation of each species was taken from the NIST Chemistry Web Book.¹⁸ The ΔS_f for each species was assumed to be dominated by the gas phase species in the reaction describing the formation of each species from the reference states. The entropy of the gas phase species were also found in the NIST Chemistry Web Book¹⁸. Explicitly, the intercept is given by:

$$\text{intercept} = \Delta H_f^0 - T\Delta S_f^0 - (N_{\text{Li}^+})k_b T \ln([\text{Li}^+]) + (N_H)k_b T \ln(pH)$$

C-C Coupling Barrier Calculations

The barrier for the C-C coupling reaction was calculated using density functional theory (DFT) geometric relaxations and the nudged elastic band (NEB) method.¹⁹ For all DFT calculations, ASE-espresso was used to run the calculations,²⁰ the plane wave cutoff was 500 eV, and the density wave cutoff was 5000 eV. The RPBE exchange correlation functional was used,²¹ and a 4, 4, 4 Monkhorst-Pack k-point grid was used.²² The convergence requirement was that the maximum force on any atom be less than 0.05 eV/Å. The unit cell was a cube with side length 10.44Å, and there were 54 atoms of Li in the unit cell. To identify the final state, a geometric relaxation was performed with two carbons occupying the same octahedral site. To identify the

initial state, a relaxation was performed with individual carbon atoms in neighboring octahedral sites, and a separate relaxation was performed with individual carbon atoms in non-neighboring octahedral sites. The latter configuration was found to be more stable, so it was used as the initial state for the NEB calculation.

Physical characterization

X-ray diffraction (XRD) was performed using a Philips PANalytical X'Pert Pro in parallel beam mode with Cu K α radiation and 0.04 rad Soller slits. Prior to XRD scanning, samples were pressed flat and sealed with Kapton (polyimide) tape against a glass slide backing in an Ar filled glove box. Gas chromatography was performed using a Thermofisher TRACE 1310 gas chromatograph equipped with a TGBOND-Q Plot column and a thermal conductivity detector in series with a flame ionization detector for product detection. Fourier transform infrared radiation (FTIR) spectroscopy was performed using a Nicolet IS-50 FTIR Advanced Spectrometer equipped with a Nicolet 2 m gas cell (ZnSe windows) connected to heated vapor and vacuum gas lines. Raman spectroscopy was conducted using a Horiba XploRA Confocal Raman. The Raman measurements were collected with an excitation laser at 432 nm, and a grating size of 1200 gr/mm in ambient air conditions.

Conflicts of Interest

There are no conflicts of interest to declare

Acknowledgements

This work was supported by a research grant (9455) from VILLUM FONDEN.

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