



CO Chemisorption at Metal Surfaces and Overlayers

Hammer, Bjørk; Morikawa, Y.; Nørskov, Jens Kehlet

Published in:
Physical Review Letters

Link to article, DOI:
[10.1103/PhysRevLett.76.2141](https://doi.org/10.1103/PhysRevLett.76.2141)

Publication date:
1996

Document Version
Publisher's PDF, also known as Version of record

[Link back to DTU Orbit](#)

Citation (APA):
Hammer, B., Morikawa, Y., & Nørskov, J. K. (1996). CO Chemisorption at Metal Surfaces and Overlayers. *Physical Review Letters*, 76(12), 2141-2144. <https://doi.org/10.1103/PhysRevLett.76.2141>

General rights

Copyright and moral rights for the publications made accessible in the public portal are retained by the authors and/or other copyright owners and it is a condition of accessing publications that users recognise and abide by the legal requirements associated with these rights.

- Users may download and print one copy of any publication from the public portal for the purpose of private study or research.
- You may not further distribute the material or use it for any profit-making activity or commercial gain
- You may freely distribute the URL identifying the publication in the public portal

If you believe that this document breaches copyright please contact us providing details, and we will remove access to the work immediately and investigate your claim.

CO Chemisorption at Metal Surfaces and Overlayers

B. Hammer,^{1,2} Y. Morikawa,² and J. K. Nørskov¹

¹Center for Atomic-scale Materials Physics and Physics Department, Technical University of Denmark, DK-2800 Lyngby, Denmark

²Joint Research Center for Atom Technology (JRCAT), 1-1-4 Higashi, Tsukuba, Ibaraki 305, Japan

(Received 31 August 1995)

A database of *ab initio* calculations of the chemisorption energy of CO over Ni(111), Cu(111), Ru(0001), Pd(111), Ag(111), Pt(111), Au(111), Cu₃Pt(111), and some metallic overlayer structures is presented. The trends can be reproduced with a simple model describing the interaction between the metal *d* states and the CO $2\pi^*$ and 5σ states, renormalized by the metal *sp* continuum. Our model rationalizes the results by Rodriguez and Goodman [Science **257**, 897 (1992)] showing a strong correlation between the CO chemisorption energy and the surface core level shift.

PACS numbers: 73.20.At, 71.15.Mb, 73.61.At, 82.65.My

Over the past three decades the field of surface science has produced a series of accurate spectroscopical techniques that can provide detailed information about the electronic structure at surfaces [1]. It would be extremely useful if spectroscopic data could also be used directly to give information about the chemical activity of the surface. This would open up new possibilities in the future search for, e.g., more efficient catalysts. Recently, Rodriguez and Goodman [2] have established a spectacular correlation between spectroscopical data (surface core level shifts) and the chemisorption energy of CO on a series of metal surfaces and overlayers. If such an approach can be generalized, we would have a means of predicting the chemical activity of surfaces based on the surface electronic properties alone.

In the present Letter, we discuss the physics of CO adsorption over metal surfaces and overlayers by presenting an extensive *ab initio* database of CO chemisorption energies calculated within density functional theory (DFT) using the generalized gradient approximation (GGA). We demonstrate that the trends in the database can be understood using a simple two-level model describing the coupling of the CO 5σ and $2\pi^*$ states to the metal *d* valence states. One key surface parameter determining the strength of the bonding turns out to be the energy of the center of the metal *d* band. This surface property can be obtained from spectroscopical methods either directly with photoemission (UPS) or indirectly through the surface core level shifts [3,4]. Using this, we demonstrate explicitly how our model of the CO chemisorption energy can account for all of the experimental data of Rodriguez and Goodman.

Before presenting the DFT-GGA database we first discuss our simple model of the trends in CO chemisorption energies. When an atom or a molecule is adsorbed on simple metal surfaces like Na, Mg, or Al without *d* states, the electronic states of the adsorbate are broadened into resonances and shifted down in energy through the interaction with the broad continuum of metal *sp* states [5,6]. This is illustrated in Fig. 1. While the adsorbate states may con-

sist of closed shells the metallic *sp* states have formed open bands, which enable energy gain through hybridization of adsorbate and metal electronic states. These metal surfaces, thereby, build up bonds to many adsorbates, including CO [7]. The surfaces of transition metals and noble metals also have open *sp* bands and therefore also form bonds to adsorbates. However, the presence of the *d* states in these metals enables a further bonding interaction between the metal *d* states and the adsorbate related states (that are already renormalized through the interaction with the metal *sp* states) [8–13] as illustrated in Fig. 1. It can easily be shown in a tight-binding framework that the total energy change caused by this interaction takes the form of a hybridization energy gain and an orthogonalization energy cost. In the limit of a small overlap *S* between the adsorbate states and the metal *d* states and of a small coupling matrix element *V* compared to the energy separation,

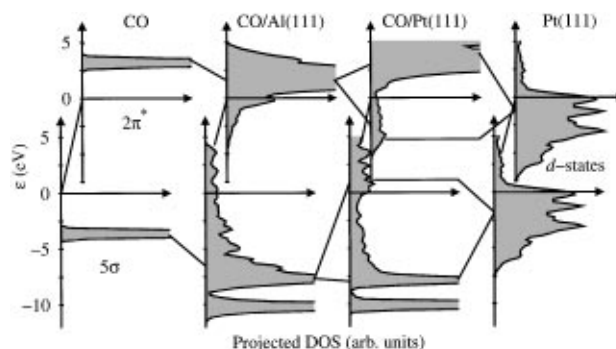


FIG. 1. The self-consistent electronic density of states (DOS) projected onto the 5σ and $2\pi^*$ orbitals of CO: in vacuum and over Al(111) and Pt(111) surfaces. Also shown is the DOS from the *d* bands in the Pt(111) surface. The sharp states of CO in vacuum are seen to broaden into resonances and shift down in energy over the simple metal surface (mixing with the 4σ state causes additional structure in the 5σ resonance). Over the transition metal surfaces the CO resonances further hybridize with the metal *d* states. This leads to shifts in the 5σ and $2\pi^*$ levels and to antibonding 5σ -*d* states at the top of the *d* bands and bonding $2\pi^*$ -*d* states at the bottom. These states have low weight in the 5σ and $2\pi^*$ projections shown.

$\Delta\epsilon$ between the two states, the hybridization energy gain becomes proportional to $V^2/|\Delta\epsilon|$ while the orthogonalization energy cost scales with SV (i.e., roughly as V^2).

For CO, adsorption experiments [14] as well as many theoretical studies [10,11,15,16] suggest that the filled 5σ and the doubly degenerate, empty $2\pi^*$ electronic states are the ones mainly responsible for the bonding to metal surfaces. We therefore write the following simple model expression for the d contribution to the CO chemisorption energy over transition metal surfaces:

$$E_{d\text{-hyb}} \approx -4 \left[f \frac{V_\pi^2}{\epsilon_{2\pi} - \epsilon_d} + f S_\pi V_\pi \right] - 2 \left[(1-f) \frac{V_\sigma^2}{\epsilon_d - \epsilon_{5\sigma}} + (1+f) S_\sigma V_\sigma \right], \quad (1)$$

where 2 is for spin, f is the fractional filling of the d bands, $\epsilon_{2\pi}$ and $\epsilon_{5\sigma}$ are the positions in energy of the (renormalized) adsorbate states, and ϵ_d is the center of the metal d bands. V and S are labeled according to the symmetry of the orbitals they describe.

The fractional filling factors, the coupling matrix elements, and the overlaps in Eq. (1) we take to be dependent only on the atomic number of the metal atom to which the CO bonds, i.e., independent on the environment of this metal atom. The environment will manifest itself through the position of the center of the d states on the metal atom in the surface before the CO adsorption. Values for the center, determined by DFT calculations, are included in Table I. We approximate f with the idealized fractional filling factor $(\nu - 1)/10$ where ν is the valence of the metal atom. As we will be concerned with the variation of the $E_{d\text{-hyb}}$ from one metal to the next it suffices to estimate V_π and V_σ in an LMTO (linear muffin tin orbital) framework [17], where they factorize as products of terms dependent only on the adsorbate and the substrate properties, respectively. This means that the present coupling matrix elements must scale precisely as the LMTO based V_{sd} used in Ref. [13] for a different adsorbate H_2 interacting with the transition metal surfaces. Introducing α and β as adjustable parameters common to all the metals, we write $V_\pi^2 \approx \beta V_{sd}^2$ and $S_\pi \approx -\alpha V_\pi$. From the DFT orbitals of CO and the various metals we find that $S_\sigma/S_\pi \approx 1.3$ is a good approximation and we therefore write $V_\sigma^2 \approx (1.3)^2 \beta V_{sd}^2$ and $S_\sigma \approx -\alpha V_\sigma$.

As the transition metal surfaces considered have very similar half filled s bands, the renormalization of the CO states by the delocalized metal states will be very alike. Guided by the results of DFT calculations for CO adsorption on Al(111) (Fig. 1) and on the transition metal surfaces with small coupling matrix elements, we use +2.5 and -7 eV (with respect to the Fermi level) for the renormalized $\epsilon_{2\pi}$ and $\epsilon_{5\sigma}$ positions, respectively.

In Fig. 2 we present the scaling of the chemisorption energy for CO within the model as compared to the set of *ab initio* DFT calculations given in Table I. Adjusting only α and β in the model, we obtain the excellent correlation in Fig. 2. α and β values of 0.063 eV^{-1}

TABLE I. Parameters and results for CO chemisorption atop a metal atom (first column) in metal surfaces and overlayers [second column: M_1/M_2 means a monolayer of M_1 on a M_2 substrate. “Ni@Cu(111)” refers to a Cu(111) substrate with every fourth surface Cu substituted by a Ni]. The center ϵ_d of the local d bands at the metal atom measured relative to the Fermi level and the fractional filling f of these bands. The coupling matrix element V_{sd}^2 (normalized to 1.0 for Cu) and the chemisorption energy E_{chem} from the DFT-GGA calculations. The last column gives the experimental chemisorption energies for CO on Ni(111) [27], Cu(111) [28], Ru(0001) [29], Pd(111) [30], Ag(111) [31], Pt(111) [32], and Al(111) [7]. All energies are in eV.

Atop	Surface	ϵ_d	f	V_{sd}^2	E_{chem}	E_{exp}
Ni	Ni(111)	-1.48	0.9	1.16	-1.36	-1.26
Ni	Ni/Ru(1000)	-1.27	0.9	1.16	-1.51	
Ni	Ni@Cu(111)	-1.18	0.9	1.16	-1.56	
Cu	Cu(111)	-2.67	1.0	1.00	-0.62	-0.52
Cu	Cu/Pt(111)	-1.88	1.0	1.00	-0.94	
Cu	Ni@Cu(111)	-2.56	1.0	1.00	-0.61	
Cu	Cu ₃ Pt(111)	-2.35	1.0	1.00	-0.53	
Ru	Ru(0001)	-1.41	0.7	3.87	-1.80	-1.66
Pd	Pd(111)	-2.16	0.9	2.78	-1.30	-1.47
Pd	Pd/Ru(0001)	-2.86	0.9	2.78	-0.98	
Ag	Ag(111)	-4.28	1.0	2.26	0.09	-0.28
Pt	Pt(111)	-2.75	0.9	3.90	-1.45	-1.50
Pt	Cu ₃ Pt(111)	-2.55	0.9	3.90	-1.51	
Au	Au(111)	-3.91	1.0	3.35	-0.04	
Al	Al(111)				-0.49	-0.21

and 1.5 eV^2 , respectively, have been used, both of which are of the right order of magnitude compared to DFT estimates [17] of 0.09 eV^{-1} and 2 eV^2 . We note that in Fig. 2 the slope of the least square fitted curve guiding the eye is close to one, which means that the adsorbate-metal d interactions described by our model can account for the main trends in CO bonding from one surface to the next. The curve is offset by $\sim 0.5 \text{ eV}$ on the vertical axis for $E_{d\text{-hyb}} = 0$, which fits with CO forming a bond of this strength on simple metal surfaces (see Table I). Outside the scale in Fig. 2 is the result for CO/Ru(0001). Here the simple model [Eq. (1)] estimates a d contribution of -2.08 eV to the chemisorption energy. Both the model and the full calculation thus show that the CO-Ru bond is very strong. However, in this case (as for most of the metals to the left in the transition metal series) where $\epsilon_{2\pi} - \epsilon_d$ is small and the d band width is large, the neglect of the latter in the two level model of Eq. (1) leads to a largely overestimated $E_{d\text{-hyb}}$.

As an important property of the model, we further note that it captures the shifts in the CO chemisorption energy from the single crystal surfaces to the overlayer structures. This is apparent in Fig. 2 as seen by the dashed lines. For Cu₃Pt, which is a very stable covalently bonded alloy, the Cu sites become less reactive than predicted by the model, while the Pt sites behave roughly as expected. As the model only describes the coupling of the metal d states to the renormalized CO $2\pi^*$ and 5σ states we conclude

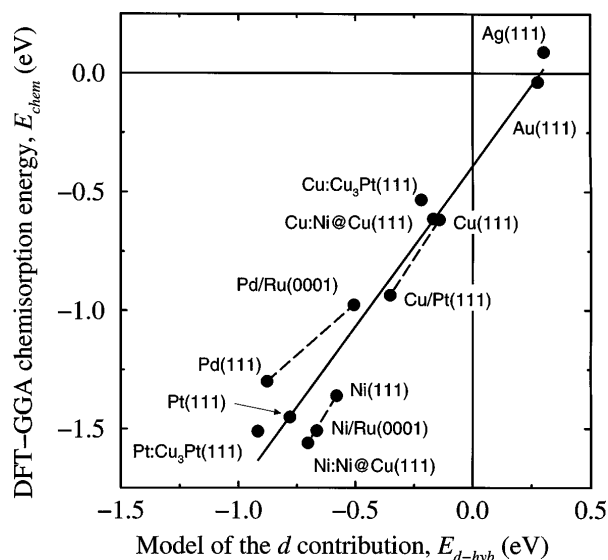


FIG. 2. Comparison of the model and the full DFT-GGA chemisorption energies for a number of metal systems.

that it is this interaction which is responsible not only for the gross trends in CO chemisorption energies over the wide range of the late transition metal surfaces considered but also for the details for metallic overlayers and alloy surfaces. We return to this below.

The *ab initio* DFT calculations [18] are performed using the local density approximation (LDA) [19] for finding self-consistent charge densities and densities of states (DOS) while using for the exchange-correlation energy in all reported total energy differences the GGA [20]. A quarter monolayer of CO is adsorbed on one side of slabs having six fcc (111)—for Ru: hcp (0001)—layers of metal atoms. Ionic cores are described with pseudopotentials [21]. The Kohn-Sham equations are solved in a basis of plane waves of kinetic energy up to 40 Ry (for Ni and Cu: up to 50 Ry) at 6 and 15 \mathbf{k} points in the C_{3v} and C_{2v} irreducible Brillouin zones, respectively.

For the present purposes, we choose to consider the DFT-GGA calculations as a computer experiment of CO adsorption on a number of metal surfaces with all ionic degrees of freedom kept fixed. Hereby we can concentrate on the ability of the model [Eq. (1)] to capture the trends caused by the *electronic* factors of the CO-metal bonding. Substrate relaxations are therefore not considered, but rather the truncated bulk geometries are used [22]. Further, for all surfaces, CO is put at the top position with a fixed metal-carbon distance of 1.94 Å and a CO bond length of 1.14 Å as reported from calculations for CO/Pd(110) [23]. The use of fixed CO coordinates is actually a good approximation. For CO on Pt(111) we find that the relaxed values are $r_{MC} = 1.88$ Å and $r_{CO} = 1.15$ Å and that this relaxation influences E_{chem} by less than 0.05 eV. That the DFT-GGA is capable of describing the CO itself and the CO-metal interaction well is suggested both by previous calculations

for Pd and Cu surfaces [23,24] and by our calculations: For CO in vacuum we get a bond strength of 10.88 eV and a vibrational frequency of 2162 cm^{-1} both of which compare well with the experimental values of 11.09 eV and 2169.8 cm^{-1} [25]. For CO adsorbed atop Pt(111) we calculate a downshift of the CO stretch frequency to 2120 cm^{-1} in agreement with an experimentally observed downshift to 2104 cm^{-1} [26]. Finally, considering the use of fixed site and coordinates, the chemisorption energies in our database agree well overall with the experimentally determined heats of CO adsorption included in Table I. In particular, for Ag(111) and Au(111) where the orthogonalization terms dominate we expect an outward relaxation of the CO to influence the chemisorption energies bringing theory and experiment in better agreement.

Our present model of the CO bonding is in complete agreement with the theoretical interpretations developed by Blyholder [15], Bagus [11], and others. The language of electron donation from the CO 5σ to the metal and backdonation from the metal to the CO $2\pi^*$ describes the concerted action of the coupling of the CO levels to the metal *sp* states and the *d* states. With the present division of the donation and backdonation into separate metal *sp* and *d* steps which follows the reasoning of Bagus and Pacchioni [11] we obtain a simple picture and a quantitative model of the electronic reason for the trends in the CO chemisorption energies over metal surfaces and overlayers.

We now return to the experimental observation by Rodriguez and Goodman of a strong correlation between the surface core level shift of different overlayers and the CO chemisorption energy [2]. Our analysis goes in two steps. First, we build on the extensive theoretical insight into the origin of the surface core level shifts by Weinert and Watson (WW) [3] and Hennig, Ganduglia-Pirovano, and Scheffler (HGS) [4]. WW show that the variation in surface core level shifts for metal overlayers is accompanied by a similar shift in the center of gravity of the *d* bands—at least towards the right in the transition metal series, while charge transfer effects are inadequate for explaining the shifts. The latter is confirmed by HGS, who also show that the *trends* in variations in the surface core level shifts for different overlayers are given by the initial state shift, that is, by the changes in the electronic structure of the unperturbed surface. From this we conclude that we can view the variation in the surface core level shifts as a measure of the variation in the *d* band center.

The second step in our analysis is then to use our model [Eq. (1)] to establish the relationship between variations in the *d* band center and the chemisorption energy. We note that in Eq. (1) the hybridization energy term related to the $2\pi^*$ dominates the expression. It therefore also dominates the differential change in E_{d-hyb} for a change $\delta\epsilon_d$ in the position of the *d*-band center, which may be caused by changes in the surroundings of the metal atom at which the CO bonds. We have

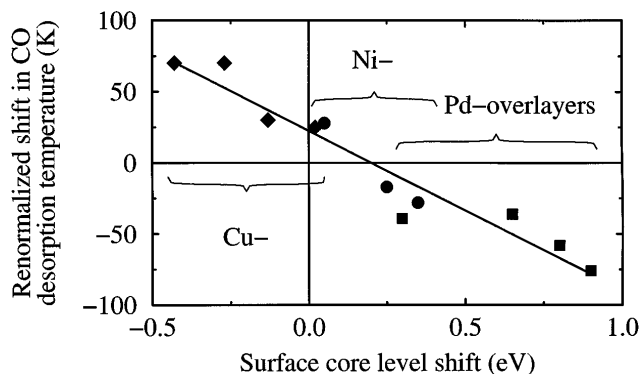


FIG. 3. The experimental data of Rodriguez and Goodman [2] for CO adsorption on Ni, Cu, and Pd overlayers plotted as the renormalized TPD peak shifts ($\Delta T/\gamma$) vs the measured shifts in surface core level positions. The reference systems for the CO TPD peaks and the surface core level positions for the overlayer structures are the respective single crystal surfaces.

$\delta E_{d\text{-hyb}} \approx [-4fV_{\pi}^2/(\epsilon_{2\pi} - \epsilon_d)^2]\delta\epsilon_d$. This suggests that different metal overlayers share the same direct proportionality relation between $\delta\epsilon_d$ and $\delta E_{d\text{-hyb}}/\gamma$, where γ is proportional to $fV_{sp}^2/(\epsilon_{2\pi} - \epsilon_d)^2$, which only depends on the overlayer metal and not on the substrate. For Cu, Ni, and Pd overlayers considered by Rodriguez and Goodman [2] γ has values of 1, 1.76, and 3.08 relative to Cu (cf. Table I). In Fig. 3 we examine if this analysis can be used on the original experimental data [2]. The figure strongly supports our thesis. Identifying the shift in the d -band center with (minus) the core level shifts and assuming the CO chemisorption energy shift proportional to (minus) the temperature-programmed desorption (TPD) peak shift ($\Delta T/\gamma$), a linear relation between the renormalized TPD peak shift and the core level shift is expected. Such a relationship is clearly present in Fig. 3, however, slightly offset from the origin. The offset might reflect a systematic offset in the core level shifts compared to the shifts of d -band centers.

The present work was in part financed by the Danish Research Councils through the Center for Surface Reactivity and Grant No. 9501775. Center for Atomic-scale Materials Physics is sponsored by the Danish National Research Foundation. JRCAT is supported by the New Energy and Industrial Technology Development Organization (NEDO) of Japan.

- [1] D. P. Woodruff and T. A. Delchar, *Modern Techniques of Surface Science* (Cambridge University Press, Cambridge, 1986).
- [2] J. A. Rodriguez and D. W. Goodman, *Science* **257**, 897 (1992).
- [3] M. Weinert and R. E. Watson, *Phys. Rev. B* **51**, 17 168 (1995).
- [4] D. Hennig, M. V. Ganduglia-Pirovano, and M. Scheffler, *Phys. Rev. B* (to be published).
- [5] D. M. Newns, *Phys. Rev.* **178**, 1123 (1969).

- [6] J. K. Nørskov, *Rep. Prog. Phys.* **53**, 1253 (1990).
- [7] T.-C. Chiang, G. Kaindl, and D. E. Eastman, *Solid State Commun.* **36**, 25 (1980).
- [8] B. I. Lundqvist, O. Gunnarsson, H. Hjelmberg, and J. K. Nørskov, *Surf. Sci.* **89**, 196 (1979).
- [9] S. Holloway, B. I. Lundqvist, and J. K. Nørskov, in *Proceedings of the International Congress on Catalysis, Berlin, 1984* (Verlag Chemie, Weinheim, Germany, 1984), Vol. 4, pp. 85–95.
- [10] R. Hoffmann, *Rev. Mod. Phys.* **60**, 601 (1988).
- [11] P. S. Bagus and G. Pacchioni, *Surf. Sci.* **278**, 427 (1992).
- [12] B. Hammer and M. Scheffler, *Phys. Rev. Lett.* **74**, 3487 (1995).
- [13] B. Hammer and J. K. Nørskov, *Nature (London)* **376**, 238 (1995); *Surf. Sci.* **343**, 211 (1995).
- [14] J. C. Campuzano, in *The Chemical Physics of Solid Surfaces and Heterogeneous Catalysis*, edited by D. A. King and D. P. Woodruff (Elsevier, New York, 1990), Vol. 3A, p. 389.
- [15] G. Blyholder, *J. Phys. Chem.* **68**, 2772 (1964).
- [16] M. A. van Daelen, Y. S. Li, J. M. Newsam, and R. A. van Santen, *Chem. Phys. Lett.* **226**, 100 (1994).
- [17] O. K. Andersen, O. Jepsen, and D. Glözel, in *Highlights of Condensed Matter Theory*, Proceedings of the International School of Physics “Enrico Fermi,” Course LXXXIX (North-Holland, Amsterdam, 1985), p. 59; J. K. Nørskov, *J. Chem. Phys.* **90**, 7461 (1989).
- [18] M. C. Payne, M. P. Teter, D. C. Allan, T. A. Arias, and J. D. Joannopoulos, *Rev. Mod. Phys.* **64**, 1045 (1992).
- [19] D. M. Ceperley and B. J. Alder, *Phys. Rev. Lett.* **45**, 566 (1980); J. P. Perdew and A. Zunger, *Phys. Rev. B* **23**, 5048 (1981).
- [20] J. P. Perdew, J. A. Chevary, S. H. Vosko, K. A. Jackson, M. R. Pederson, D. J. Singh, and C. Fiolhais, *Phys. Rev. B* **46**, 6671 (1992).
- [21] G. B. Bachelet, D. R. Hamann, and M. Schlüter, *Phys. Rev. B* **26**, 4199 (1982); N. Troullier and J. L. Martins, *ibid.* **43**, 1993 (1991).
- [22] When allowing for relaxations of the uppermost two layers of the clean Cu/Pt(111) surface the DFT-GGA values for the CO chemisorption change by less than 0.01 eV.
- [23] P. Hu, D. A. King, S. Crampin, M.-H. Lee, and M. C. Payne, *Chem. Phys. Lett.* **230**, 501 (1994).
- [24] P. H. T. Philipsen, G. te Velde, and E. J. Baerends, *Chem. Phys. Lett.* **226**, 583 (1994).
- [25] K. P. Huber and G. Herzberg, *Constants of Diatomic Molecules*, Molecular Spectra and Molecular Structure Vol. IV (Van Nostrand Reinhold, New York, 1979).
- [26] E. Schweizer, B. N. J. Persson, M. Tüshaus, D. Hoge, and A. M. Bradshaw, *Surf. Sci.* **213**, 49 (1989).
- [27] J. B. Miller, H. R. Siddiqui, S. M. Gates, J. N. Russell, Jr., and J. T. Yates, Jr., *J. Chem. Phys.* **87**, 6725 (1987).
- [28] P. Hollins and J. Pritchard, *Surf. Sci.* **89**, 486 (1979).
- [29] H. Pfnür, P. Feulner, H. A. Engelhardt, and D. Menzel, *Chem. Phys. Lett.* **59**, 481 (1978).
- [30] H. Conrad, G. Ertl, J. Koch, and E. E. Latta, *Surf. Sci.* **43**, 462 (1974).
- [31] G. McElhiney, H. Papp, and J. Pritchard, *Surf. Sci.* **54**, 617 (1976).
- [32] H. Steininger, S. Lehwald, and H. Ibach, *Surf. Sci.* **123**, 264 (1982).