Development of PEM Electrolysis at Elevated Temperatures

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Development of PEM Electrolysis at Elevated Temperatures


Presented by: Erik Christensen
Department of Energy Conversion and Storage
Technical University of Denmark
Projects completed 2011

**WELTEMP:** “Water Electrolysis at Elevated Temperatures”, European Commission, FP7

**HyCycle:** “Center for Renewable Hydrogen Cycling”, Danish Council for Strategic Research (DSF)

Projects ongoing 2012

**MEDLYS**

**MEDLYS:** “Medium temperature Water Electrolysers”, Danish Council for Strategic Research (DSF)

**PROCON**

**PROCON:** “Danish-Chinese Centre for Intermediate Temperature Proton conducting Systems”, Danish National Research Foundation (DNRF)
Concept of Research

Elevated temperature ($\geq 120^\circ$C)
- To obtain higher efficiency (Kinetics and thermodynamics!)
  (steam or liquid?)

Requires New Materials Development due to strongly increased demands to materials:

**Component:**
- Membrane
- Current collectors
- Bipolar plates
- Catalysts

**To replace/modify:**
- Nafion
- Titanium
- Titanium
- IrO$_2$, RuO$_2$, Pt: Are they stable?

MEAs $\rightarrow$ Electrolyser
The WELTEMP Project and the Partners
www.weltemp.eu

FP7, Collaborative Project, small or medium-scale focused research project
Duration: January 1st, 2008 - April 30th, 2011
Total costs: 3.2 million Euro  EC Funding: 2.4 million Euro

The Partners
Technical University of Denmark (Coordinator)          Denmark
Institute of Chemical Technology Prague              Czech Republic
Institute of Macromolecular Chemistry ASCR             Czech Republic
The Norwegian University of Science and Technology  Norway
IHT Industrie Haute Technologie SA                  Switzerland
Acta S.p.A.                                         Italy
Tantalum Technologies A/S                            Denmark
Danish Power Systems ApS                             Denmark
The partners of HyCycle
(Electrolysis and Photocatalysis, www.hycycle.dk)

Technical University of Denmark, DTU Energy Conversion (coordinator)

Center for Individual Nanoparticle Functionality (CINF), Department of Physics, Technical University of Denmark

Center for Atomic-scale Materials Design (CAMD)
Department of Physics, Technical University of Denmark

Department of Physics and Chemistry, University of Southern Denmark

Institute of Chemical Engineering, biotechnology and Environmental Technology
University of Southern Denmark

IRD Fuel Cells A/S, Denmark

Danish Power Systems ApS, Denmark

Tantaline A/S, Denmark
Objectives

(1) Membranes:
Temperature-resistant polymer membranes, *operational temperatures* $\geq 120^\circ$C
*Anion conducting (Alkaline)* membranes should be surveyed as well.

(2) Electrocatalysts:
Stability of IrO$_2$ based anodes and Pt cathodes at temp. $\geq 120^\circ$C should be demonstrated.
Low loadings!
New non-noble metal catalysts for use under alkaline conditions.

(3) “Construction materials”:
Development of current collectors and bipolar plates made in steel coated with tantalum, and
having excellent corrosion-, contact resistance-, and conductive properties.

(4) Membrane Electrode Assemblies (MEAs):
Methods for preparation of membrane-electrode assemblies (MEAs) with targets of fabrication
of MEAs single cell performance approaching 1.55 V at 1.0 A/cm$^2$ at a temperature
above $120^\circ$C

(5) Test Electrolysers:
Design, construction and testing of a prototype electrolysers
Concepts of Polymer Membrane Materials

1) PBI (polybenzimidazol)
   - Phosphoric acid doped (apparently not stable!)

2) PFSA (Perfluorosulfonic acid, Nafion, Aquivion)
   - Water is required to be present inside the structure -
     otherwise no proton conductivity
   - Water evaporates from the membrane at $T > 100^\circ$C
     then three ways to go:

   a) Modify Nafion/Aquivion by adding hygroscopic fillers (steam
      or liquid water)
   b) Doping with $\text{H}_3\text{PO}_4$! (only steam feeding)
   c) Pressurising the cell and working with liquid water

3) Anion conducting membranes (alkaline “PEM” electrolysis)
Nafion®

PBI

Aquivion™
“Short side chain”-
PFSA
Steam- or Pressurized Water Electrolysis?

STEAM electrolysis:
PFSA Membrane (Aquivion)  
doped with $\text{H}_3\text{PO}_4$

Pressurised LIQUID water electrolysis
PFSA Membrane (Aquivion)

Anode: 0.98 mg/cm² IrO₂, Cathode 0.34 g/cm² Pt, GDL 0.5mm Ta coated steel felt, Aquivion membrane, 63µm thick, Temperature 130°C, Atmospheric Pressure

Anode: 1.72mg/cm² IrO₂, Cathode 0.80 g/cm² Pt, GDL 0.5mm Ta coated steel felt, Aquivion membrane, 60 µm thick, Temperature 120°C, Pressure 3 bar
Ionic conductivities of membranes

Membrane materials:
(Mechanical) Reinforcement is important to conductivity

150°C, 6 bar, 100 % RH
Construction materials:

Tantalum coated steel to replace titanium

Flowplates and anodic current collectors ("GDL"):
Titanium not stable – replaced by tantalum coated steel

CVD: Chemical Vapor Deposition

\[
2\text{TaCl}_5 (g) + 5\text{H}_2(g) \rightarrow 2\text{Ta} + 10\text{HCl}(g) \uparrow \quad (800^\circ \text{C})
\]

A: Steel felt (uncoated)

B and C: Tantalum coated steel felts

D: Labscale flowplates coated with tantalum
Desired corrosion rate in 85 % H₃PO₄:
≤ 0.1mm/year

Corrosion rate of TITANIUM vs. temperature and the extent of polarization

<table>
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<th>$v_k$ [mm/a]</th>
<th>80 °C</th>
<th>120 °C</th>
<th>150 °C</th>
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<td>-500 mV</td>
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<td>$E_{kor}$</td>
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<td>485</td>
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<tr>
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<td>0.2</td>
<td>-</td>
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Corrosion rate of TANTALUM vs. temperature and the extent of polarization.

<table>
<thead>
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<th>$v_k$ [mm/a]</th>
<th>80 °C</th>
<th>120 °C</th>
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<td>$E_{kor}$</td>
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<td>&lt;0.01</td>
<td>&lt;0.01</td>
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<tr>
<td>500 mV</td>
<td>&lt;0.01</td>
<td>&lt;0.01</td>
<td>&lt;0.01</td>
</tr>
<tr>
<td>2000 mV</td>
<td>&lt;0.01</td>
<td>&lt;0.01</td>
<td>&lt;0.01</td>
</tr>
<tr>
<td>2500 mV</td>
<td>&lt;0.01</td>
<td>&lt;0.01</td>
<td>&lt;0.01</td>
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</tbody>
</table>
Catalyst Materials/Catalyst Support Materials:

**Cathode:** Pt/C

**Anode:** IrO$_2$ with or without *support*

**Supports:**
- SnO$_2$
- SnO$_2$–Sb$_2$O$_3$ (electronic conduct.)
- SnO$_2$–Sb$_2$O$_3$-”SnHPO$_4$” (elect. + proton!)
- TiO$_2$ (non-conductive supports: SiC - can also provide an improved performance !)

PEM steam electrolysis at 130 °C and atmosphere pressure. The anode loadings were 0.7 mg cm$^{-2}$ IrO$_2$, 1.4 mg cm$^{-2}$ IrO$_2$/SnO$_2$, 1.4 mg cm$^{-2}$ IrO$_2$/ATO, and 1.4 mg cm$^{-2}$ IrO$_2$/SnP-ATO, respectively. The cathode was made of 40 % Pt/C at a loading of 0.7 mg Pt cm$^{-2}$. Membranes used were PA doped Aquivion (0.05 mm).
Alternative concept: Alkaline MEAs

Durability issue?

Life test obtained with the ACTA alkaline MEA 475 mA/cm², $T_{\text{cell}} = 40 \degree C$. Now more than 6000 h!

Performance problem: Main issue is ionomer for catalyst layer preparation (Teflon was used). Active non-noble metal catalysts were developed for both anode and cathode!
Medium temperature/Intermediate temperature cell

Temperatures: 200-400°C

MEDLYS and PROCON projects:

Inorganic proton conducting membranes:

CsH$_2$PO$_4$

Nb-P

Bi-P

Nd-P

etc.

Various alternative (non-noble element) catalysts will be tested.
Achievements/Breakthroughs

• PEM Steam electrolysis can be carried out

• PFSA membranes can be made conductive at high temperatures by phosphoric acid doping.

• Reinforced membranes provide higher conductivity

• Pressurised cells reached higher performances than steam electrolysers at 130°C (until now...)

• Tantalum coated steel felt as anode GDL

• MEAs based on anion conductive/alkaline membranes can be prepared, high durability at temp. up to 60°C have been observed.

• Alkaline MEAs, working without noble metals.
Application perspectives

The research represents a survey of various types of electrolyser technologies:

**Acidic PEM, Alkaline PEM, Liquid water feeding, steam electrolysis (÷SOEC)**

**PEM: decentralized units**  **Alkaline: Centralized units**

For large scale use in the nearer future according to governmental plans, alkaline technologies will be important!

*(Denmark: 50% of total electricity consumption from sustainable sources in 2020=>
Large electricity storage capacity will be needed very soon!)*
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Thank You for your attention!