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Materials and Structures

Enhancing the efficiency of electrochemical desalination of stones: A proton pump approach --Manuscript Draft--

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1 **Enhancing the efficiency of electrochemical desalination of stones: A proton pump**
2 **approach**
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Abstract:

Soluble salts are among the most harmful alteration agents affecting the building materials. In recent years, several researches have been devoted to counteract alterations induced by soluble salts using electrokinetic techniques. However, the applicability of these techniques for conservation purposes remains limited due to adverse side effects, such as the extreme pH values occurring near the electrodes, which can affect the stone to be treated. The decrease in efficiency of the treatment caused by the dominant transport of H^+ and OH^- groups is also an undesired effect. The reduced duration of these treatments due to the drying of the material in contact with the anode also limits their practical use. To overcome these problems, a new electrokinetic design that includes a so called proton pump is presented in this report. This design is based on placing two electrodes in the anodic compartment in order to modulate the net amount of H^+ produced. The design was applied to desalinate sandstone samples contaminated with several soluble salts. The application of this new approach allowed us to establish an additional electroosmotic process at the anode, which was able to increase the duration of the treatment. Moreover, the new setup provided improved pH buffer ability due to the generation of OH^- in the anodic compartment, which increased the effectiveness of the treatment by hindering the entrance of H^+ in the porous structure.

Keywords: Electro-migration, electro-osmosis, desalination, sandstone, soluble salts

1. Introduction

The damage caused by the action of the soluble salts puts at risk the perdurability of the buildings that are affected, mainly those that are part of the cultural heritage. To solve this problem, various techniques have been developed to remove or reduce –as far as possible– the soluble salts present inside of the porous materials used in their construction. Among these techniques, the electrokinetic method stands out for the high efficiency in the extraction of soluble salts, both in the laboratory [1-8] and in situ [9-12]. Using this methodology, salt reduction percentages close to 100% can be reached, even at a considerable depth into the material, as demonstrated in [5], where 6-cm thick granite samples were completely desalinated. However, the electrokinetic treatments have several disadvantages, which must be overcome for their application, including:

- Extreme changes in pH around the electrodes, due to water hydrolysis. These extreme pH values can lead to chemical alteration of the materials [1, 5-8, 13-15].
- Colour changes in the treated material due to oxidation, hydrolysis, and solubilisation of some minerals, mainly those containing iron [13, 16].
- Generation of microfractures due to the stresses caused by the electric field across the pores of the materials [17], especially those materials containing piezoelectric minerals [18-20], such as quartz.

The main solutions so far suggested to mitigate these adverse effects consist of the use of a buffer electrolyte and in the interposition of a poultice between the electrodes and the material to be treated. The poultice contains calcium carbonate in the anode, as stated in some studies [5-7, 10, 15], and kaolin in the cathode [5-7].

In the traditional electrokinetic methods, there is also an important factor that affects efficiency, defined as the percentage of ion content after passing a certain charge. During this process, the generation of a high quantity of OH^- and H^+ decreases the efficiency because these ions have a mobility higher than that of the ions to be removed. This makes OH^- and H^+ the

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main responsible for the transport of the electrical current. Consequently, the mobilization of the ions of interest decreases and their transport number decreases, thereby decreasing the removal efficiency of the method [11, 15, 21-24].

The solution proposed to this problem lies in retaining –as far as possible– the OH⁻ and H⁺ ions within the poultices that are in contact with the electrodes [5-7, 10, 15], thus avoiding their entrance in the material. Such a solution increases the transport number of the ions to be removed.

Moreover, in this kind of treatment, it is important to maintain a high humidity level both in the poultices and in the porous structure to be treated, to allow a homogeneous distribution of the electrical current. During the treatment, and following several types of researches [16, 25-27], an electroosmotic process takes place due to the higher intrinsic solvation of cations in the majority of porous materials, which tend to accumulate water in the cathode side. However, at the anode the opposite phenomenon occurs, which causes drying of the material located at this electrode.

The subsequent drying of the material at the anode becomes a limiting factor regarding the time of application of the technique, because the drying worsens the contact between the electrode and the material located at the anode and hinders the current flow [1, 10-11, 16]. No technical solution to solve this problem has been developed. Only two solutions seem feasible: frequent replacement of the material in contact with the anode, or periodical wetting of this material by addition of fresh electrolyte.

This work focuses on the evaluation of a new electrokinetic configuration based on the use of two electrodes in the anodic compartment between which different potentials are applied. Using such an electrokinetic configuration, an increase of the duration of treatment is expected. This is due to the establishment of an electroosmotic process in the anode that compensates the loss of water. Moreover, the pH buffering capacity of the system will be increased, which will increase the effectiveness of the ion removal.

2. Materials and Methods

2.1. Porous material used

An ashlar of yellow sandstone extracted from the Kronborg castle located at the Oresund's Strait entrance (Denmark) was employed for this study. Ten stone cubes ($5 \times 5 \times 5 \text{ cm}^3$) and 12 stone prisms ($6 \times 6 \times 3 \text{ cm}^3$) were obtained from this ashlar.

Using the cubic-shaped samples, accessible porosity to water, water absorption coefficient, capillary porosity, and capillarity coefficient were calculated following the various European standards [28-29], and procedures established in different researches [30]. Mercury accessible porosity was determined using a Micromeritics Autopore IV9500 porosimeter that works with two pressure ranges, 0.20 Mpa-225 MPa (high pressure) and 14 kPa-0.20 MPa (low pressure). The pore network charge (zeta potential) was determined using a Zetasizer Nano Z, from Malvern Instruments.

The prismatic shaped samples were used for the evaluation of the desalination treatment. The prisms were contaminated with a saline solution composed of 9.9% NaNO_3 , 6.6% NaCl , and 0.66% CaSO_4 in weight. The contamination cycles consisted of 3 days of capillary absorption followed by a drying process of 2 days at room temperature. Before starting the next contamination cycle, the prisms were turned over to put the opposite face in contact with the solution, trying to achieve a distribution of salts as homogeneous as possible. Five contamination cycles were conducted.

At the end of the contamination cycles, three prisms were used to calculate the initial ion content (the reference value). For this, the prisms were segmented into four pieces (previously labelled on all the sandstone prisms) by hammer and chisel, weighed, and crushed using a pulverizer ring mill. The powder obtained from each piece was homogenized. Three powder samples (10 g each) were taken from each piece. The powder samples were maintained under stirring in 25 mL of ultrapure water for 24 h, and filtered through $0.45\text{-}\mu\text{m}$ nitrocellulose filters. The anion content (Cl^- , SO_4^{2-} , and NO_3^-) and the cation content (Ca^{2+} and Na^+) were

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obtained from the extracts using high resolution chromatography (Dionex ICS-100) and inductively coupled plasma optical emission spectrometry (Varian 720-ES inductively coupled plasma optical emission spectrometry), respectively. Also, conductivity (EC-Metro BASIC 30), using the normalized equation of Unhruh [31] and pH (pH-meter BASIC 20) were measured in each extraction.

The remaining nine stone prisms were used to perform each of the programmed desalination tests (three stone samples per test).

2.2. Description of the Electrochemical Setups.

Two different assemblies were prepared for this study. One of these, the ‘Traditional Setup’, has been widely used in several types of research [2, 11], which here we refer to as T-S. This assembly consists of two cells: one is the anodic compartment (called in this study as medium 2, Figure 1), which hosts electrode 2, and the other is the cathodic compartment, which hosts electrode 3. The second assembly corresponds to the ‘Double Anode Setup’ which here we refer to as DA-S; this assembly has one more cell in the anodic compartment (medium 1, Figure 1), to host another electrode (electrode 1). The electric field established between electrode 1 and electrode 2 works as a proton pump, i.e., a system that allows for modulating of the net amount of H^+ produced at the anode (medium 2 in Figure 1b) by pumping unwanted protons from medium 2 to medium 1.

INSERT FIGURE 1

In the T-S setup, one power supply was used (Power supply 2, Figure 2). In the DA-S setup, two power supplies connected in series (Figure 2) were necessary to regulate the current flowing through each cell.

In the two setups (T-S and DA-S), a constant voltage of 8V was established between electrode 2 and electrode 3 using the Power Supply 2 (Figure 2). In the DA-S setup, electrode 1

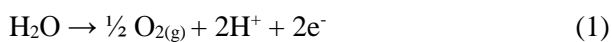
1 was polarized using the Power Supply 1 (Figure 2). With this setup, two different assays were
2 performed, applying two different potentials between the electrodes located in the anodic
3 compartment (electrodes 1 and 2): DA-S2 with $\Delta V_{1-2} = 2V$ and DA-S3 with $\Delta V_{1-2} = 4V$. These
4 different potentials were applied in order to evaluate the magnitude of the effects that the proton
5 pump can generate.
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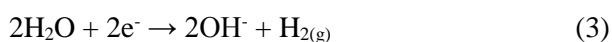
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14 In the DA-S setup, electrode 1 (Figures 1 and 2) is set at potential $V_1 + V_2$; therefore, it
15 can be considered as the effective anode to which the anions to be removed from the porous
16 material will be forced to migrate. In the electrode 2, a lower potential V_2 with respect to the
17 cathode was established (Figures 1 and 2). In this way, this electrode was employed as a kind of
18 bipolar electrode because it behaves as an anode with respect to electrode 3, located in the
19 cathode, and hence it allows the mobilization of the anions present inside the porous material
20 towards its vicinity. Furthermore, this electrode behaves as a cathode with respect to electrode
21 1, which forces the anions retained by poultice PA2 (located in the medium 2) to migrate
22 towards the medium 1.
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35 Titanium bars were used as electrodes due to its inertness with respect to the main
36 reactions in the anode compartment (oxidation reactions). The main reactions occurring at both
37 electrodes are:
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42 Anodic electrochemical reactions:



51 Cathodic electrochemical reactions:



57 According to previous studies [4-7, 15], it was necessary to place between the stone
58 samples and the electrodes a material able to retain the ions forced to migrate to the vicinity of
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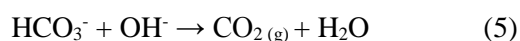
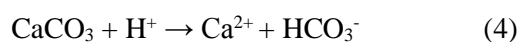
1 the respective electrode. Three different poultices made with different compounds were used for
2 this purpose.
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5 In the anodic compartment, medium 1 (corresponding to the DA-S, Figure 1) was filled
6 with a poultice made of cellulose housing a citric acid/sodium citrate buffer solution, pH 6 (here
7 referred to as PA1). The efficacy of this buffer solution had been demonstrated in previous
8 studies [5-6]. The weight cellulose/solution ratio employed was 1:7.6
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14 Cellulose was selected because of its high liquid storage capacity, which is due to its
15 large pores with an access pore size diameter greater than 10 μm [32-33]. As cellulose has a
16 zero buffering capacity, the H^+ generated at electrode 1, in contact with this poultice, are free to
17 migrate to the vicinity of the electrode 2, forming water. This effect corresponds to the “proton
18 pump” which allows maintaining of the necessary water content for the proper functioning of
19 the treatment.
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28 The medium 2 (corresponding to the anode of both setups, Figure 1) was filled with a
29 poultice made of kaolin, calcium carbonate, and the same buffer electrolyte mentioned above,
30 following a weight ratio 1:2:1.3, respectively. We refer to this poultice as PA2.
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36 In previous work, the addition of calcium carbonate to this poultice was effective at
37 buffering the acidification of the anodic compartment [4-7, 10, 15-16]. The existence of the
38 proton pump would allow buffering of the acidification in the vicinity of electrode 2. However,
39 if for any reason the pH of this medium becomes acid, the calcium carbonate would begin to
40 decompose, decreasing the pH buffering, according to the Eq. 4 and 5:
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54 In the DA-S setup, the two media of the anodic compartment were separated using a cellulose
55 filter.
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1 In the cathodic compartment a poultice made of kaolin and citrate/citric acid
2 electrolyte buffered at pH 6 was used (which we refer to as PC). The weight ratio
3 kaolin/electrolyte employed was 1.6:1. Kaolin was selected by its ion exchange capacity, which
4 allows the retention of the OH⁻ generated at electrode 3, thereby hindering its entrance into the
5 system and increasing, in consequence, the efficiency of the treatment [5-7].
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11 The stone prisms were positioned in the same direction (i.e., the face previously labelled
12 with the numbers 1 and 2 was put in contact with the anodic compartment and the face labelled
13 with the numbers 3 and 4 with the cathodic compartment). To reduce as much as possible the
14 evaporation of water through the stone prisms, these were wrapped with plastic film.
15 Furthermore, to hinder the evaporation through the holes of the cells, these were closed with
16 rubber stoppers, allowing a minimal space for gases to escape.
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29 **2.3. Operation of the DA-S setup**

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31 As stated before, this DA-S setup contains two electrodes in the anodic compartment that work
32 as a proton pump. The operation of the proton pump and its efficacy was evaluated under two
33 different conditions, by establishing two different potential drops between the anode electrodes
34 (DA-S2, 2V; DA-S3, 4V).
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41 Taking into account the Faradays' law and the Eq. 6 and 7, the variation in
42 concentration of protons, $\Delta[H^+]$, at each electrode is related with the circulating charge and,
43 therefore, with the current intensity.
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49 At electrode 1: $\Delta[H^+] = Q_1/F \omega = I_1 t / F \omega$ (6)
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51 At electrode 2: $\Delta[H^+] = Q_2/F\omega = (I_2-I_1) t / F \omega$ (7)
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54 where Q corresponds to the electrical charge (in C) circulated; F is the Faraday constant (96485
55 C/mol), ω corresponds to the volume of the compartment, I corresponds to the circulating
56 current (A), and t is the time elapsed (s)
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2 In this setup, it is expected that the higher potential difference established between the
3 electrodes located at the anodic compartment (ΔV_{1-2}) causes an increase in the intensity I_1 . This
4 should lead to increased H^+ concentration in the vicinity of the electrode 1, and decreased H^+
5 concentration in the vicinity of electrode 2.
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9 Following this hypothesis, the following effects at electrode 2 located close to the porous
10 material (Figure 1b) are expected:
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- 12 • Generation of water molecules by neutralization ($H_3O^+ + OH^- \rightarrow 2H_2O$). These water
13 molecules will keep, for a longer period, the moisture level needed for the flowing of
14 electric current through the circuit.
15
- 16 • The buffering of pH in the vicinity of the electrode 2. The drastic drop of the pH that
17 occurs in the anodic compartment will occur only nearby the electrode located the
18 furthest from the stone sample, which is the electrode 1. In the closest electrode
19 (electrode 2), the generation of OH^- will allow buffering of the acidic pH values.
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- 21 • The reduction of the H^+ amount entering into the porous material.
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33 The operation of the DA-S setup would be similar to the use of a bipolar electrode in the anode;
34 the technique used in [34]. However, the DA-S setup presents the novelty of being able to
35 regulate the potential difference established between the electrodes located in the anodic
36 compartment, which allows for controlling the magnitude of the effects indicated above.
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46 **2.4. Methods for evaluating the effectiveness**

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48 Three tests were performed: T-S setup, DA-S2 setup and DA-S3 setup. The three tests lasted 11
49 days and were performed in triplicate using three stone prisms (6 x 6 x 3 cm) in each of the
50 tests.
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55 The potential drops of voltage between the electrodes (T-S: electrode 2 vs. electrode 3;
56 DA-S: electrode 1 vs. electrode 2, electrode 1 vs. electrode 3 and electrode 2 vs. electrode 3)
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1 were recorded during the tests. For this, for 8 h per day, the potential drops were measured
2 every 30 min. Also, at the same time interval, the total intensity flowing through the stone
3 prisms was calculated, considering the potential drop recorded in the 1 kΩ resistance (Figure 2).
4 Moreover, the intensity that flows through the different media in DA-S was obtained directly
5 from the power supplies.
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11 At the end of each experiment, each stone prism was segmented into four pieces
12 (indicated in blue in Figure 1) by hammer and chisel: pieces 1 and 2 correspond to the surfaces
13 closest to the anode while pieces 3 and 4 correspond to those closest to the cathode. Poultrices
14 were segmented in three slices: a slice close to the anode (A), a medium slice (B), and a slice
15 close to the cathode (C) (Figure 1).
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23 In all these stone pieces and poultice slices, the following parameters were evaluated:
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- 25 • Water content. This parameter was measured as the amount of water present in the
26 material (stone or poultice) at the end of the test with respect to the dry weight obtained
27 after drying the samples of rock and poultrices at 105°C, with the exception of the
28 samples of the poultice PA1 (made of cellulose) that were dried at 50 °C. The
29 calculation of the water content allows us to assess the existence of electroosmotic
30 process at the anodic compartment and to determine whether the water content inside
31 the stone was sufficient to allow a homogeneous flow of the electric current.
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33 • pH, conductivity, and ion content of the aqueous extractions were obtained following
34 the procedure above explained. These analyses were made in triplicate for all samples
35 of rock and poultrices. With the data obtained for each piece of the stone prisms, the
36 desalination efficacy was calculated as the percentage of reduction of each ion (E%-
37 ion), which is established following the expression used in [35].
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$$53 \quad E\%-\text{ion}_{1-4} = 100 \times (\text{Initial ion content}_{1-4} - \text{Final ion content}_{1-4}) / \text{Initial ion content}_{1-4}.$$

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59 **3. Results and discussion**

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3.1. Stone's properties

Table 1 summarizes the main properties of the sandstone. The results show that this rock is of high porosity (in terms of accessible porosity to water and mercury). The difference between these parameters ($\approx 4\%$) indicates that this rock has a certain pore volume hardly accessible to water. Also, practically all the pores seem to fit within the range of the capillary pores, because the difference between the accessible porosity to water (13.24%) and capillary porosity (11.47%) is very low. Pore size distribution data were reported elsewhere [7]. This stone has two main pore families: Pore diameter between 1 and 10 μm (74%) and pore diameter higher than 100 μm (5%). The remaining porosity corresponds to pores that lie in the nanometer range. Thus, the majority of pores (80% of the volume) lie in the capillary range (i.e., access size diameter between 0.1–100 μm) [36].

INSERT TABLE 1

The surface charge of this stone was negative (-24 mV). This fact represents hindered flow of anions but favoured transport of cations (with water) towards the cathode. This behaviour is typical of almost all the porous materials used in building constructions [16, 25-27, 37].

The initial ionic content of the prisms shows that 1) the nitrate and chloride content was higher than that of the rest of the ions, in accordance with the composition of the contaminating solution; and 2) the deviation between the ion content for the three reference stone prisms was very low, which demonstrates that this content is representative for the rest of the stone prisms.

3.2. Intensity and resistance measurements

Figure 3a shows that the resistance to current flow was greater for the stone prisms desalinated with the T-S setup than for the samples desalinated with the DA-S setups (DA-S2 and DA-S3). Between DA-S setups no difference was found regarding the resistance of the stone prisms.

1 Regarding the flow of electric current (Figure 3b), it was seen that the current using the T-S
2 setup dropped dramatically after the first day. This was attributed to the shrinkage of the
3 poultice PA2 caused by the loss of water which prevented a proper electrode-poultice contact
4 [1]. In fact, electrode 2 has to be shaken up periodically to improve the ionic contact, facilitating
5 the current flow. Conversely, in the DA-S2 and DA-S3 setups, the current intensity flowing
6 through the stone prisms was much higher than that of the T-S setup, also remaining constant
7 during the tests. Shrinkage of the poultices at the anodes did not occur in the DA-S tests, which
8 granted a good permanent contact between the electrodes 1 and 2 and the respective poultices.
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11 Comparing DA-S setups, it was found that the electric current was higher in the DA-S3
12 test than in DA-S2. This result is expected considering that, for a similar resistance of the stone
13 prisms and poultices (Fig. 3a), the application of a greater potential between anode and cathode
14 allow a greater electric current to flow (Ohm's law: $\Delta V = I \times R$).
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34 3.3. Water Content

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36 Table 2 shows the water content of the poultices and stone prisms after the tests. The results
37 show differences in water content in poultices depending on the setup. Using the T-S setup, a
38 greater loss of water at the cathode with respect to the anode occurred, while with both DA-S
39 setups the results obtained were opposite (i.e., the greater loss of water occurred in the anodic
40 compartment with respect to the cathode). The results of the DA-S were in concordance with
41 previous research findings [16, 25-27], where an electroosmotic process was responsible for the
42 hydration of the material located in the cathode compartment, as long as the surface charge of
43 the material remains negative, such as in this case (Table 1).
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55 The results achieved using the T-S setup could be related, as mentioned in [2, 4, 16, 25],
56 to a decline of the electroosmotic process due to higher ion content in the stone prisms.
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1 Comparing both setups, the T-S setup leads to a greater reduction in water content in the
2 medium 2 and the cathode with respect to the same media under the DA-S setup. The water loss
3 was higher than 20% for the T-S setup in medium 2, while it ranged between 10% and 15%
4 using the DA-S setup. The loss of water at the cathode using the T-S setup was around 23%,
5 while using the DA-S setup this parameter ranges between 8% and 13%. These results suggest
6 that the DA-S setup was able to maintain sufficient moisture content in the vicinity of the stone
7 prisms due to the action of the proton pump, allowing the extraction process to proceed without
8 interruption.
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11 For the DA-S setups, regardless of the potential established (DA-S2 or DA-S3), the loss
12 of water was higher in the PA1 poultice than in the PA2. This fact suggests a hydration process
13 occurring during the treatment due to the establishment of an electroosmotic process. The
14 generation of an electroosmotic process in clays is a well-known phenomenon [38]. At the
15 cathode (PC poultice), the loss of water was still lower than at the anodes, which again suggests
16 electroosmosis.
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19 Comparing DA-S2 and DA-S3, it was noted that, under higher potential (DA-S3),
20 higher water decreasing occurs in medium 1, in accordance with the electroosmotic process
21 caused by the proton pump. Moreover, it was noticed that the water contents of medium 2 and
22 the cathode are more similar under the DA-S3 condition than DA-S2. This fact suggests that the
23 electroosmotic processes established in the medium 2 and in the cathode were similar, so that
24 the drying of both compartments occurred at similar rates.
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28 Regarding the water content in the stone prisms, there was a significant difference
29 between the setups evaluated. With respect to the T-S setup, the DA-S2 setup caused an
30 increase in the water content of about 24%, and the DA-S3 setup caused an increase of about
31 43%. These results showed that the proton pump generated with the DA-S setup allowed an
32 increase in the water content inside the material.
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2 Therefore, as indicated in the hypothesis, the DA-S setup allowed us to solve the
3 problem related to the water loss in the anode compartment, thereby moisturizing the anode
4 material located close to the stone.
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10 **3.4. pH measurements**

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12 Figure 4 shows the different pH values measured in the aqueous extractions of the different
13 slices (A, B and C) of the poultices that fill each of the cells (PA1 and PA2 at the anode and PC
14 at the cathode), as well as the pH measured on the faces of the stone prisms close to the anode
15 and cathode.
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23 Regarding the T-S setup, the pH values of medium 2 were slightly basic. These results
24 show, as expected based on other studies [5-7, 15], that the amount of CaCO_3 that was added to
25 the PA2 poultice was sufficient to buffer the acidification processes. The pH values reached
26 ranged from 10 to 11 at the cathode, values characteristic of this medium. The pH values of the
27 stone prisms were high (ranged from 9 to 11 approximately), especially in the surface closest to
28 the anode. This result could be related to the amount of Ca^{2+} ions released during the pH
29 buffering. Ca^{2+} can enter into the rock during treatment, react with the atmospheric $\text{CO}_{2(g)}$ after
30 the treatment, during drying, and, therefore, carbonation processes can generate CaCO_3 [39].
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44 Regarding the DA-S setup, the pH values reached in the medium 1 ranged from 2.5 to
45 3.7, indicating heavy acidification. Using the DA-S3 setup, an increase in the acidification
46 process took place, due to the higher current I_1 between electrodes 1 and 2 (see Eq. 6 and Fig.
47 3). In the medium 2, using the DA-S2 setup, the pH values in the slices A and B were similar to
48 those reached with the T-S setup. In the slice C, in which electrode 2 was placed, the pH values
49 became more basic (8.75) due to the generation of OH^- at the surface of the electrode, which
50 helped to buffer pH (see Eq. 7). Using the DA-S3 setup, the results show a similar trend,
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although the pH values reached were higher than those with DA-S2, especially at the surface of the electrode 2.

At the cathode, the pH values reached using the DA-S setups ranged from 11.2 to 11.6, which were on average more basic than those obtained using T-S setup. However, taking into account the deviation of the results obtained after the application of the T-S in slices B and C, the differences are not significant. The reason for this slightly higher pH could be related to a decreased releasing of calcium ions from the PA2 poultice during the pH buffering process or to increased OH⁻ retention by the PC poultice. The calcium released can flow from the anode to the cathode, where it will react with the OH⁻ ions forming Ca(OH)₂ precipitate [4]. This process, the fixation of the OH⁻ with Ca ions through calcium hydroxide production, is a normal aspect of the T-S setup.

Using the DA-S setup, the proton pump favours the buffering of pH at the anode, reducing CaCO₃ dissolution from the PA2 poultice and the subsequent calcium release. Therefore, the pH values measured were higher in the poultice PA2, especially with the application of the DA-S3 setup, due to the higher working intensity of the proton pump. Therefore, using the DA-S setup, it is possible to reduce the amount of CaCO₃ used in the preparation of the PA2 poultice. The reduction depends on the potentials applied to each of the electrodes because the buffering capacity of the setup will increase with ΔV_{1-2} .

The pH values reached in the stones desalinated using DA-S setups were different as a function of the voltage applied to electrode 1. With the DA-S2, the pH values were, on average, higher than those reached with the DA-S3. Using the DA-S2, the pH values reached were lower on the surface closest to the anode than those reached on the surface closest to the cathode. Using the DA-S3, pH values were similar on both rock surfaces (being in some cases a bit higher on the surface near the cathode), and were close to the characteristic pH of the untreated stone (7.15 ± 0.62).

1 The results obtained with both DA-S setups confirm that the proton pump allows for
2 reducing the risk of exposing the material to an acid pH. This is because the proton pump
3 guarantees a permanent buffer system that does not run out over time. While the traditional
4 solution, based on the exclusive use of a poultice made of CaCO_3 , is not permanent since it
5 depends on the presence of remaining CaCO_3 . The protection afforded by the proton pump can
6 be controlled through ΔV_{1-2} . Also, the incorporation of the proton pump allows a reduction in
7 the amount of H^+ penetrating the rock and competing in the transport of ions to be extracted
8 (i.e., the efficiency of the desalination process should be increased).
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21 **3.5. Conductivity measurements**

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24 Figure 5 shows the conductivity values of aqueous extractions of the slices of the poultices and
25 the stone pieces. These values give an idea of the ionic content of the media analysed; the higher
26 the conductivity, the greater the ionic content.
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31 The conductivity reached in the stone pieces, both in the anode and in the cathode, was
32 always higher with the T-S setup than with the DA-S setups. These results are indicative that the
33 ionic content in the stone prisms desalinated by this setup is higher than in those desalinated
34 with DA-S setups.
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41 Comparing the DA-S setups, the trend is different in the anode and in the cathode. In the
42 surfaces closest to the anode, the lowest conductivity values are obtained with the DA-S2 setup
43 and in the surfaces closest to the cathode, with the DA-S3 setup.
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48 At the cathode, the conductivity values were higher, in most cases, using the T-S setup
49 than using any of the DA-S setups. This might be due to the higher amount of Ca^{2+} ions released
50 in the T-S setup, since the buffering of the pH in the anode using this setup is exclusively due to
51 the action of the PA2 poultice. The lowest conductivity in this compartment was achieved using
52 the DA-S3 setup.
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2 In medium 2, the highest conductivity values were achieved with the T-S setup and with
3 the DA-S3 setup, and we detected no significant differences between the setups. In medium 1,
4 the highest conductivity was reached using the DA-S3 setup, which could indicate a greater
5 mobilization of anions towards this medium.
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11 12 13 14 15 **3.6. Ion content in the poultice** 16

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18 Table 3 shows the total average amount of the ions retained in each of the compartments (anode
19 and cathode) and also the ion content retained in each of the media that form the anodic
20 compartment (medium 1 and medium 2).
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25 The results in Table 3 show that the ion content in the anode is higher than in the
26 cathode. Moreover, the ion content is higher in the DA-S setups anodes than in the T-S setup,
27 especially of the DA-S3. These facts are true for all ions except for Na⁺. The sodium content is
28 higher at the cathode than at the anode; and at the anode, it is higher in the T-S setup than in the
29 DA-S setups. A similar behaviour is observed for the anions.
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34 Regarding the chloride ion, the results show that the DA-S setup allows extracting
35 higher amount of chloride than the T-S setup (Table 3), mainly with the DA-S3 setup. This is
36 the expected result considering the higher current recorded.
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41 Analysing each of the anode media that form the different setups (Table 3), it was found
42 that the amount of chloride retained in medium 2, the one closest to the rock, was slightly higher
43 when the T-S setup was used. Using the DA-S setup, part of the chloride extracted from the
44 stone was mobilized towards medium 1, in which the electrode with the highest potential was
45 located. This mobilization was greater using the DA-S3 setup. Concerning the cathode, the
46 quantity of chloride was higher when a T-S setup was used, followed by the DA-S2 setup. The
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existence of this ion in this compartment is due to the establishment of diffusive and advective processes in the PC poultice.

The results obtained by analysing the nitrate and sulphate ions were similar to those obtained by analysing the chloride ion (i.e., higher extraction in the anode compartment with DA-S setup than T-S setup, mainly with DA-S3 setup). In this DA-S3 configuration, the highest mobilization of both ions towards the medium 1 was obtained. At the cathode, the highest amount retained was obtained with the T-S setup and the lowest with the DA-S3 setup, especially for sulphate.

These results suggest that the drying of the anodic compartment, when the T-S setup was used hinders the process of forced ion migration under the influence of the electric field. This fact facilitates the extraction of anions present in the stone via advection and diffusion processes [40]. This would explain why using the T-S setup the anion content at the cathode is higher than that obtained using the DA-S setup (especially the sulphate content). On the contrary, using the DA-S setup, which reduces the drying rate of the anode's poultice due to the action of the proton pump, the migration process of these anions under the influence of the electric field can be held for a longer time.

In the case of the cations, two different behaviours are observed for sodium and calcium. The sodium content is higher than the calcium content, and it is also higher at the cathode than at the anode. This might be because sodium is present in the electrolyte employed to prepare each of the poultices. Moreover, migration pushes the cations towards the cathode. In the cathode, the sodium content was similar in the three setups while, in the anode, the content was higher in the T-S setup and lower in the DA-S setups. This fact could be related to a different migration ability of this ion in each of the setups.

Concerning the calcium content, it is higher at the anode because calcium is part of the PA2 composition. Moreover, the calcium content at the anode is higher in the DA-S setups, especially in the DA-S3 while at the cathode the highest content of calcium is obtained for the

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T-S setup. These results point directly to the dissolution of CaCO_3 in poultice PA2 during the pH buffering at the anode, and the subsequent migration of calcium ions released from the anode to the cathode.

The higher calcium content present in the anode of the DA-S setup, especially in the DA-S3, demonstrates that the proton pump allows pH buffering through the action of the electric field, thus reducing the amount of calcium released compared to if the poultice were performing the buffering function.

Using the DA-S setups, the release of calcium occurs mainly in the contact zone between poultices PA1 and PA2, which explains the presence of calcium ions in medium 1. The higher calcium content in medium 2 of the DA-S3 can be explained by the fact that the buffering capacity of the pH in this setup is higher than in the DA-S2 (this is consistent with the discussion in section 3.4). In the T-S setup, the buffering function is performed by the calcium carbonate present in the PA2 poultice (as seen in Eqs. 4 and 5). During this buffering process, part of the calcium is released and migrates towards the cathode. Thus, the T-S setup makes the calcium content increased in the cathode and decreased in the anode.

3.7. Ion content in the stone

Figure 6 shows the desalination efficacy percentages (E%-ion) of the anion extraction (chloride, nitrate and sulfate) for the four sampling points considered in the rock (two of the anode and two of the cathode).

In general, regardless of the anion, extraction rates in most cases were higher in the DA-S setups. This is because the desalination process in the DA-S setups worked through the whole period thanks to the proton pump, while in the T-S setup the drying of the materials (poultices and rock) discontinued the process. With the T-S setup enrichments were produced which were manifested with negative percentages of efficacy and in general the dispersion of the values obtained with this test was much higher.

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Regarding the results for the DA-S setups, differences between the anode and the cathode for the three ions were found. With respect to chloride and nitrate, the average percentages of efficacy achieved using the DA-S2 setup were practically constant (approximately 40%) throughout the entire sample. With the DA-S3 setup, the percentages of efficacy were lower in the areas closer to the anode (approximately 30%) and higher in the areas closer to the cathode (approximately 50% for chloride and above 60% for nitrate), surpassing in most cases in this zone the percentage of extraction achieved with the DA-S2. This indicates that the DA-S3 is more efficient because it allows a greater mobilization of the anions (i.e. an extraction from the cathode towards the anode and consequently a higher accumulation of both ions at the anode with respect to the DA-S2). This higher accumulation would not have occurred if the test had not been stopped.

INSERT FIGURE 6

Using the T-S setup, sulfate enrichments were much higher than those of Cl^- and the NO_3^- , especially in the zones closer to the anode. The sulfate enrichments are related to the lower mobility of this ion, especially when there are more mobile ions in the media, as indicated in [7]. Higher sulfate extraction efficiencies are reached (slightly higher than 60%) with the DA-S2 than those achieved for chloride and nitrate; this occurs in practically the whole rock, with the exception of level 3. The extraction percentages of sulfate obtained with DA-S3 are lower than those of chloride and nitrate, but they show a similar tendency, i.e., higher at the cathode (30–50%) and lower at the anode (being, at some level, negative). This fact, together with the contents of salts retained in the anode (Table 3), shows that DA-S3 is still producing a mobilization of the sulfate ions from the cathode to the anode, a mobilization much more effective than that obtained by the DA-S2.

Figure 7 shows the percentage of efficacy in the extraction of calcium and sodium for the four sampling points of the rock (two of the anode and two of the cathode). In general, for both calcium and sodium, extraction efficiencies were higher in the DA-S setups than in the T-S

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setup. Again, the negative influence of the drying of the materials (poultices and rock) with the T-S setup in the mobilization of ions is evident.

INSERT FIGURE 7

Regarding calcium ion, the highest efficacy was achieved with the DA-S3 setup because, by operating the proton pump with greater intensity, the input of calcium into the rock is lower. Calcium comes from the dissolution of CaCO_3 of the PA2 poultice. It is necessary to take into account that according to previous results [5, 39], the calcium entering the stone can precipitate in the form of CaCO_3 and, as a consequence, can collapse some percolating pores. Therefore, the application of the DA-S setups allows for control of the possible structural changes occurring in the rock.

With respect to sodium (Figure 7b), the percentages of efficacy achieved with the application of the DA-S2 and DA-S3 setups were very similar in almost all of the depth levels sampled. These percentages were higher than those achieved with the T-S setup, where high enrichments were seen.

In general terms, the DA-S setups and TS setup differ in the content of salts remaining in the rock, the amount of salts that are extracted to the poultices (Table 3), the intensity circulating during the 11 days of treatment (Figure 3b), and the water content in both poultices and rocks (Table 2). These differences confirm that the DA-S setups have been operating correctly during the 11 days of the test and could even run for longer, which would increase the differences found with the T-S setup.

4. Conclusions

The main conclusions derived from this study are:

- 1) The design of the DA-S setup is adequate to make a proton pump work at the anode.

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2) The proton pump of the DA-S setup allows an electro-osmotic process at the anode, bringing water that increases the duration of the treatment and ensures a correct contact between the electrodes and the poultices.

3) The proton pump of the DA-S setup gives to the system a self-buffering capacity of pH in the anodic compartment since the system can generate at the relative anode OH^- ions, which can counteract the acidification usually occurring at the anode.

4) The proton pump of the DA-S setup allows increasing of the effectiveness of the conventional desalination treatment. This is because the operation of the pump hinders the entrance of the H^+ ions into the porous material that reduces the mobilization of the ions to be extracted.

5) All of these effects of the proton pump are increased when the potential between the electrodes located in the anodic compartment increases. Thus, this parameter which is easily adjustable.

6) The proton pump buffering effect decreases the calcium released by decomposition of the CaCO_3 of the PA2 poultice. The lower release of calcium decreases the probability of blocking percolating pores in the rock. However, the lower release of Ca^{2+} ions causes a disadvantage with respect to the pH of the cathode: the lower the amount of calcium available, the less the ability to neutralize OH^- . So, although the proton pump substantially improves the performance of a traditional setup, it would be advisable to conduct a study focused on the analysis of whether an improved cathodic compartment (developed following the same theoretical principles) could solve this problem.

Conflict of Interest: The authors declare that they have no conflicts of interest.

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FIGURE CAPTIONS

Figure 1. Scheme of the two assemblies evaluated in this research (Traditional setup (T-S, a); Double Anode setup (DA-S, b)) and electrokinetic phenomena that take place in each of the tested setups. The different slices in which the poultices and the stone samples were divided for their analyses are also shown.

Figure 2. Scheme of the electric circuit of the two setups Traditional setup (T-S) and Double Anode setup (DA-S). The solid line corresponds to the electric circuit used in the T-S setup. The dashed lines correspond to the addition necessary to apply the DA-S setup. The scheme contains the electrodes and the variable resistance (poultices, stone and a calibrated resistance of 1 k Ω employed for accurate current measurements). All of the elements are connected in series.

Figure 3. a) Average measurements of the resistance (in k Ω); and b) the total current intensity (in mA) that flows through the samples during the 11 days of the tests.

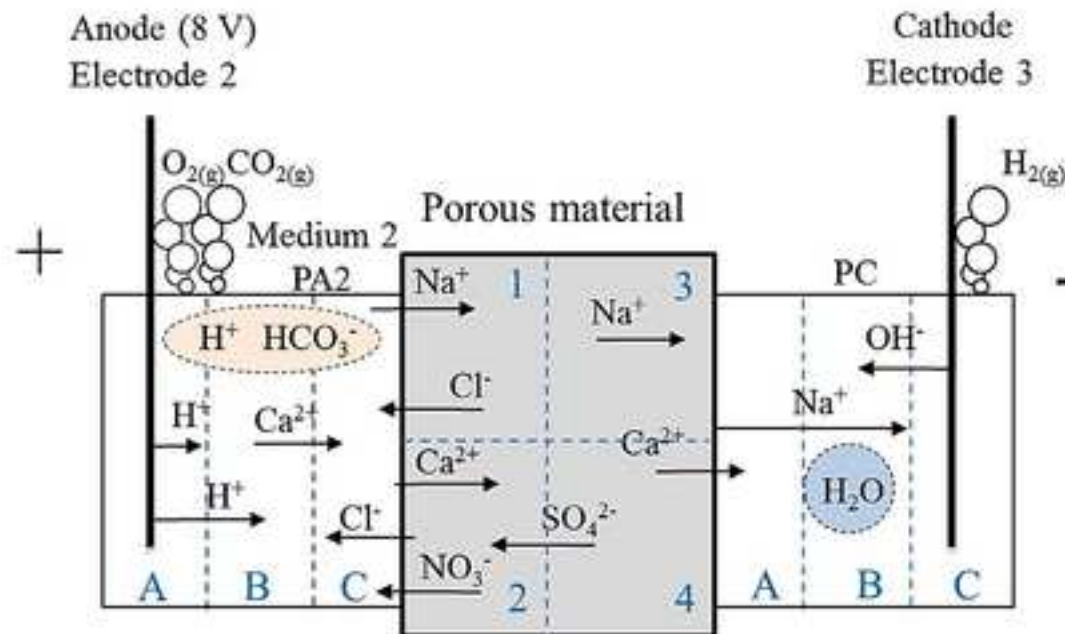
Figure 4. pH average measurements reached in the different parts of the setup (poultice slices and stones pieces).

Figure 5. The average conductivity measurements reached in the different parts of the setup (poultice slices and stone pieces).

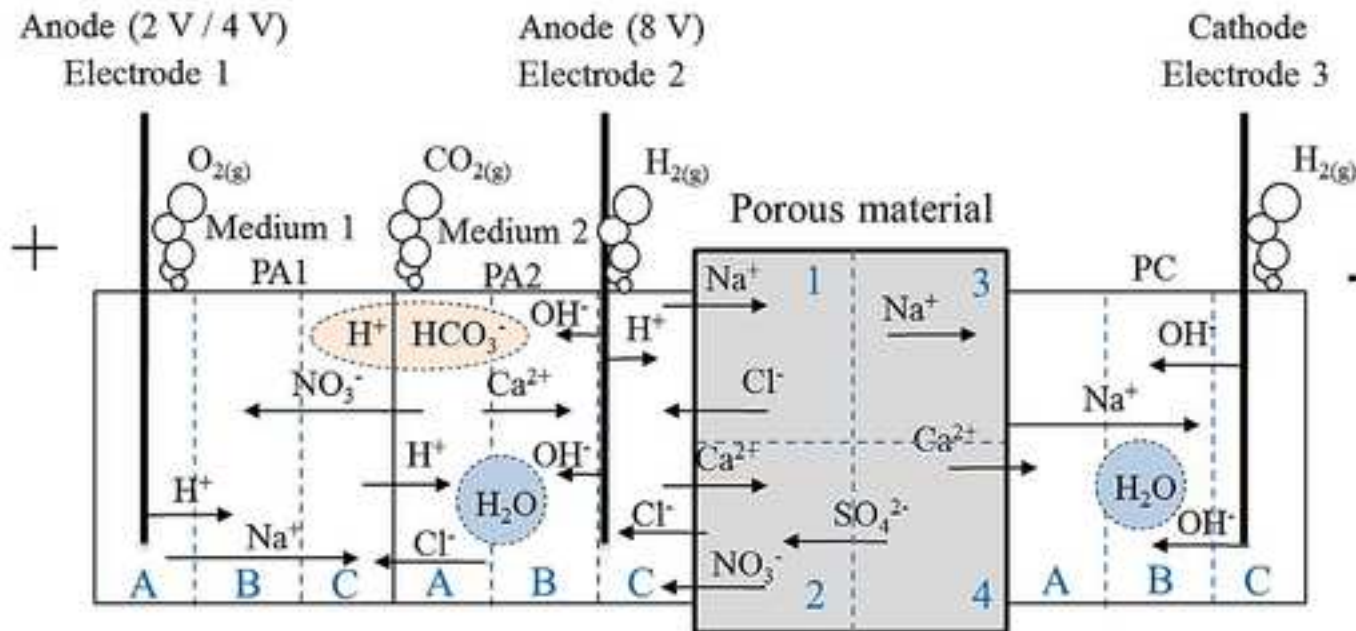
Figure 6. The average percentage of reduction (n = 3 samples of stone per setup) of Cl⁻ (a), NO₃⁻ (b) and SO₄²⁻ (c) in the different pieces of the stone.

Figure 7. The average percentage of reduction (n = 3 samples of stone per setup) of Ca²⁺ (a) and Na⁺ (b) in the different pieces of the stone.

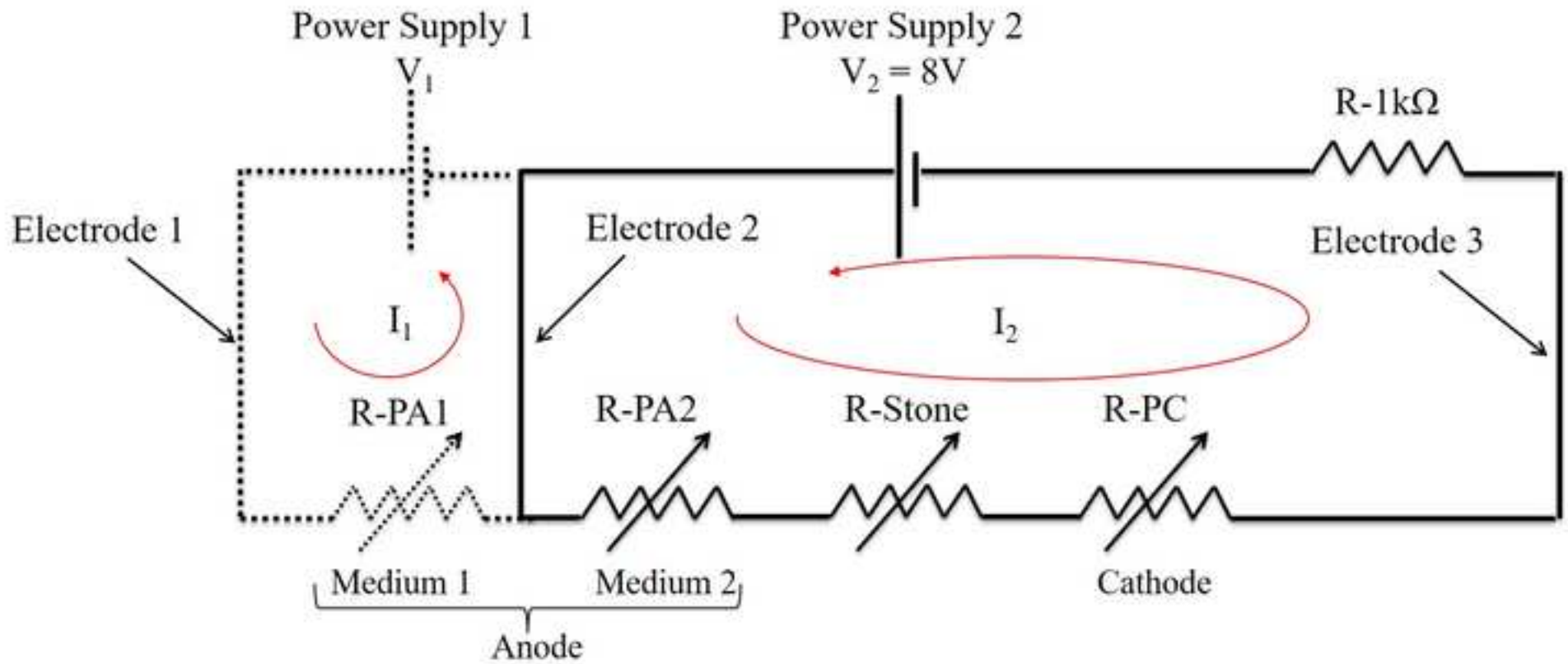
a) T-S



b) DA-S



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Double Anode Setup:

DA-S2: $V_1 = 2V$

DA-S3: $V_1 = 4V$

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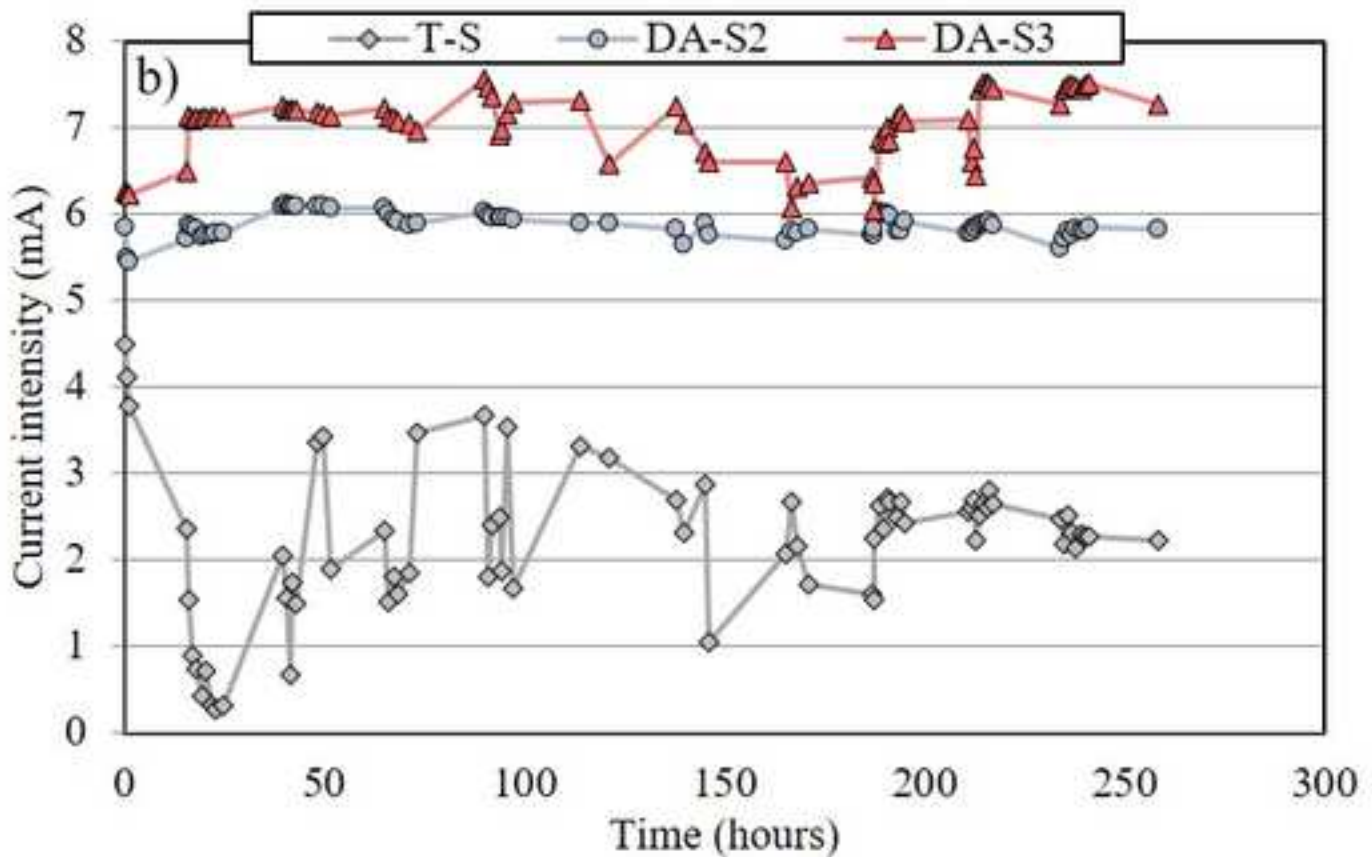
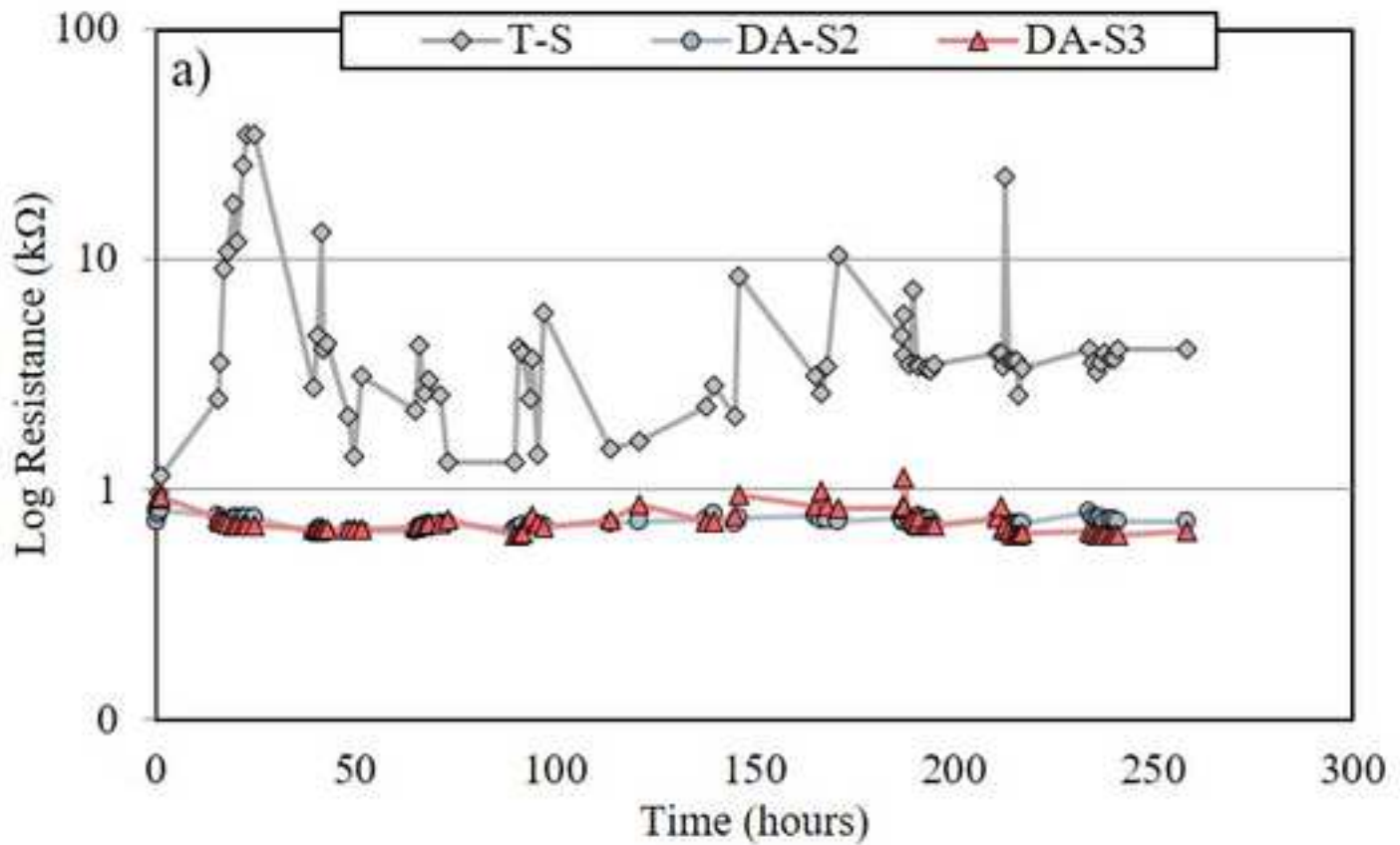


Figure 4

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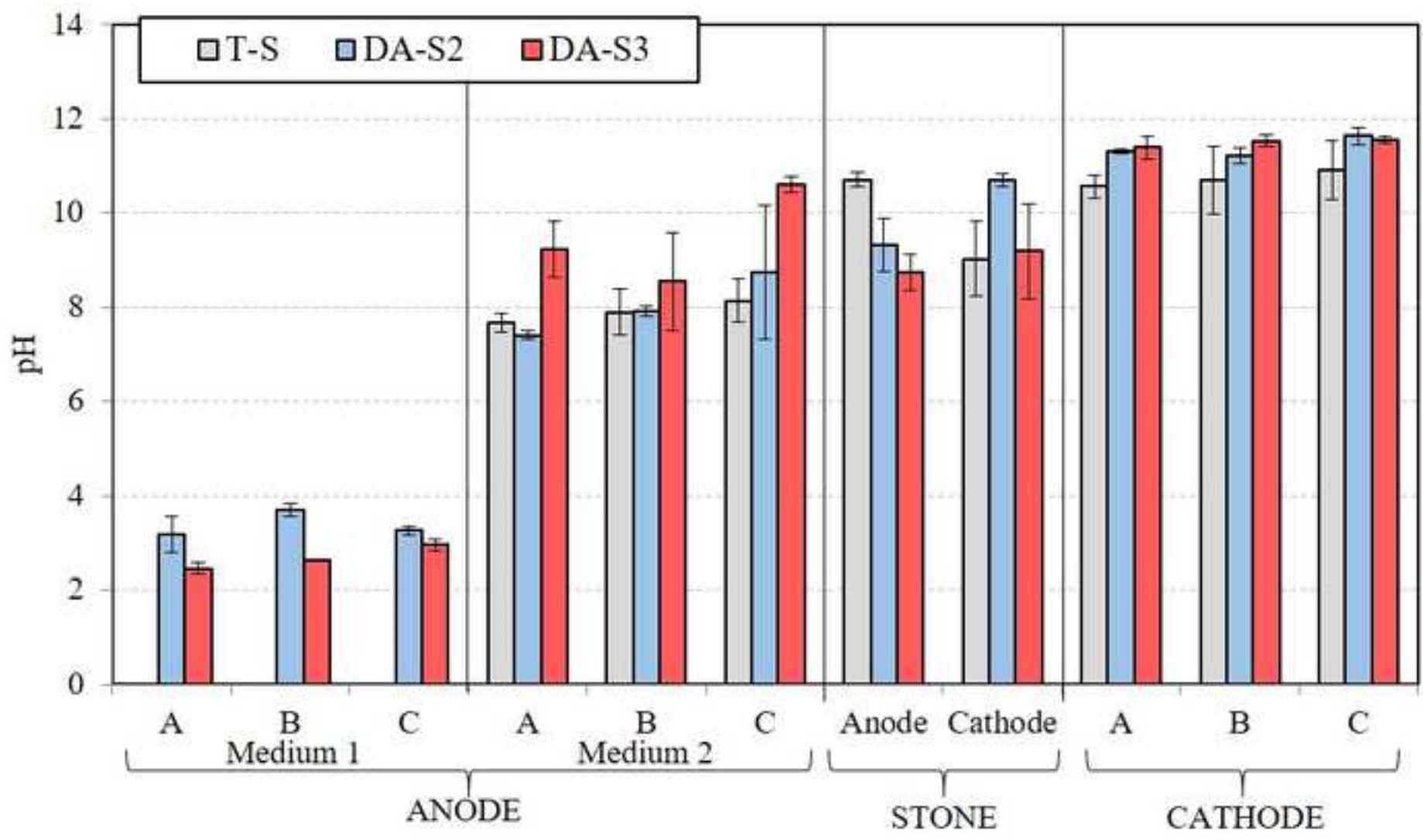
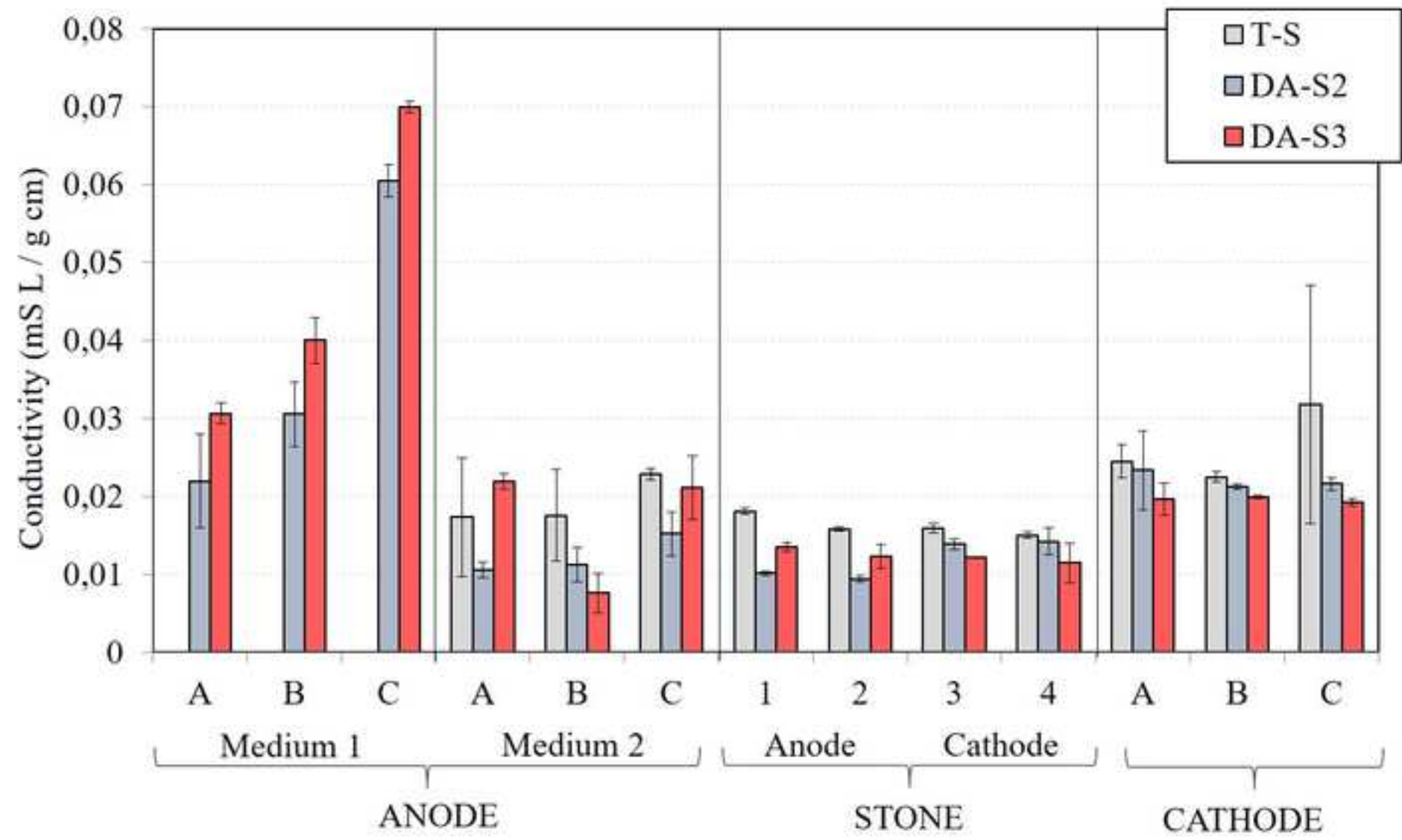
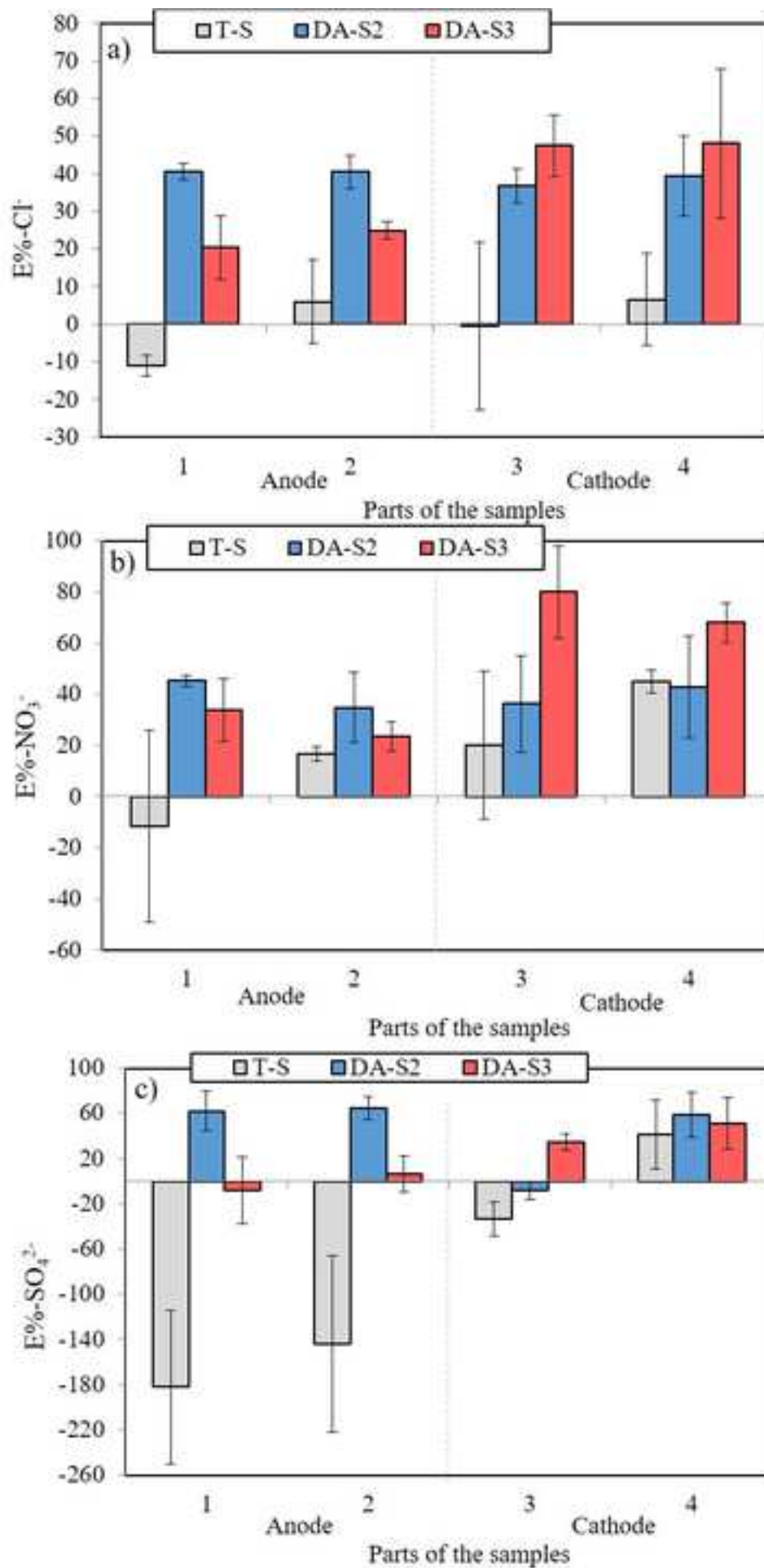


Figure 5

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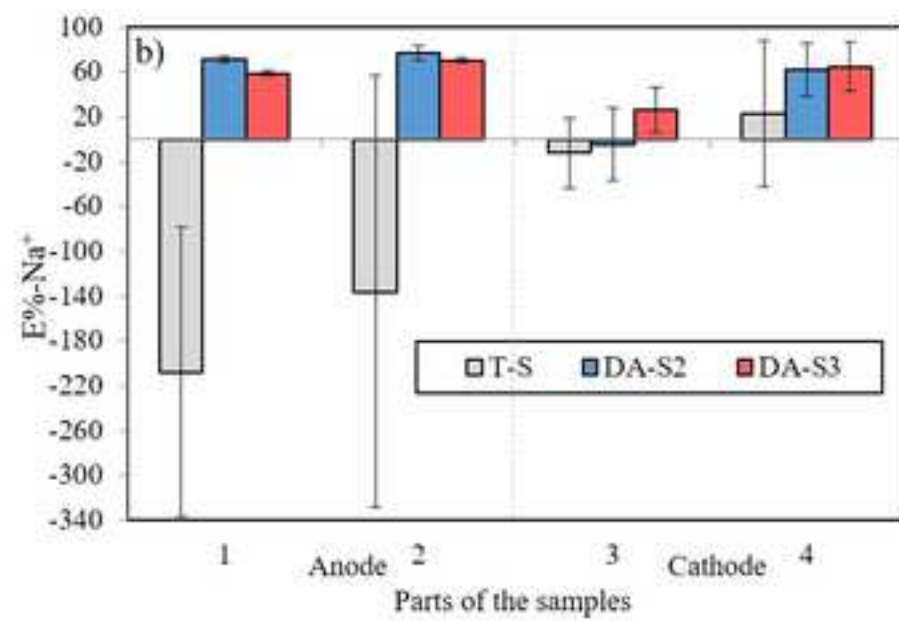
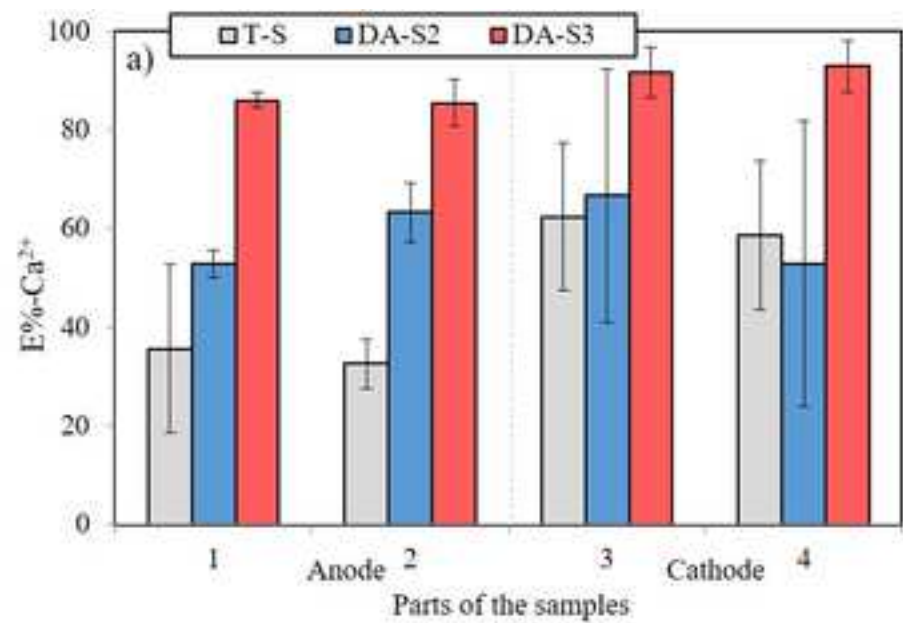
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Table 1. Main properties of yellow sandstone calculated following different European standards and researches. **Average initial ion content in the sandstone prisms.**

PARAMETERS	Yellow Sandstone	According to:
Water accessible porosity [%]	13.24±0.31	RILEM 1980 [28]
Mercury accessible porosity [%]	17.30±0.15	
Capillary porosity [%]	11.47±0.25	[29]
Capillarity coefficient [kg/m ² s ^{0.5}]	0.046±0.001	ICR-CNR 1985 [30]
pH	7.15±0.62	
Conductivity (µS L/ g cm))	0.014±0.00	[31]
Zeta potential (mV)	-24±9.60	
Initial ion content (% g/g)		
Cl ⁻	0.248±0.009	
SO ₄ ²⁻	0.021±0.003	
NO ₃ ⁻	0.335±0.017	
Ca ²⁺	0.021±0.003	
Na ⁺	0.018±0.005	

Table 2. Water content in % (average value and standard deviation) of the poultice and of the stone before and after the treatments applying the T-S setup and DA-S setups (DA-S2 and DA-S3). The reduction of the water content (in %) with respect to the initial one is also shown

Water Content in the Poultice (%)	Initial	Final	Reduction (%)
T-S			
• Medium 2 (PA2)	41.26±2.28	32.83±0.27	20.45±0.65
• Cathode (PC)	58.31±0.72	44.99±0.15	22.84±0.26
DA-S2			
• Medium 1 (PA1)	580.51±7.09	398.83±6.15	31.30±1.06
• Medium 2 (PA2)	41.26±2.28	36.38±0.68	11.84±1.64
• Cathode (PC)	58.31±0.72	53.16±1.27	8.82±2.18
DA-S3			
• Medium 1 (PA1)	580.51±7.09	389.02±5.76	32.99±2.91
• Medium 2 (PA2)	41.26±2.28	34.99±0.12	15.21±0.30
• Cathode (PC)	58.31±0.72	50.33±0.26	13.67±0.44
Water Content in the Stone Prism (%)	T-S	DA-S2	DA-S3
	1.94±0.13	2.42±0.06	2.77±0.08

Table 3: Total average amount (n = 3 samples of stone per assembly) of Cl⁻, NO₃⁻, SO₄²⁻, Ca²⁺ and Na⁺ retained in the different compartments (anode and cathode) during the Traditional setup and Double Anode setups (DA-S2 and DA-S3). Ion content retained in each of the different mediums that form the anodic compartment is shown. The total amount is the sum of the total average amount analyzed in each of the slices of the poultice (3 samples per layer representing a total of 9 samples analyzed by poultice).

Ion content in each compartment (mg/compartment)						
Anode				Cathode		
Ion	T-S	DA-S2	DA-S3	T-S	DA-S2	DA-S3
Cl ⁻	241.66±25.3 2	281.28±40.0 6	327.06±11.1 6	26.01±10.10	24.04±1.98	20.28±0.73
NO ₃ ⁻	287.78±91.8 8	412.65±108. 88	517.57±20.9 3	27.42±11.87	16.27±6.55	9.76±2.92
SO ₄ ²⁻	19.17±1.55	26.63±3.89	31.01±10.06	12.98±6.75	2.66±1.46	2.23±0.74
Ca ²⁺	44.95±16.12	65.61±8.74	125.01±55.9 8	1.62±1.10	0.80±0.42	0.35±0.12
Na ⁺	269.83±107. 29	193.05±69.6 2	183.31±84.9 6	324.41±128. 99	324.72±154. 94	309.04±3.81
Ion content in each medium of the anodic compartment (mg/medium)						
T-S		DA-S2		DA-S3		
Ion	Medium 1	Medium 2	Medium 1	Medium 2	Medium 1	Medium 2
Cl ⁻	-	241.66±25.3 2	63.33±3.56	217.95±24.7 6	101.43±5.69	225.63±2.20
NO ₃ ⁻	-	287.78±91.8 8	145.61±47.9 0	267.03±29.0 8	251.63±12.9 4	265.94±1.86
SO ₄ ²⁻	-	19.17±1.55	9.40±0.10	17.23±2.65	17.82±5.35	13.19±1.76
Ca ²⁺	-	44.95±16.12	33.72±5.98	31.89±0.20	53.92±14.56	71.10±25.02
Na ⁺	-	269.83±107. 29	33.83±8.68	159.22±40.5 5	1.40±0.83	181.82±59.2 4