Trends in the exchange current for hydrogen evolution

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The hydrogen evolution reaction is one of the most important electrochemical reactions, whereby protons from solution combine with electrons at an electrode to form first hydrogen atoms chemisorbed at the electrode surfaces and then H₂ gas. The ability of a given metal to catalyze the hydrogen evolution reaction is usually measured by the exchange current density, which is the rate of hydrogen evolution per surface area at the electrode potential, where the reaction is at equilibrium.

Different materials have widely different exchange current densities. In this paper, we investigate the origin of these variations. We calculate the chemisorption energy for hydrogen for a number of metals. We also collect a number of measured exchange current densities. In this paper, we investigate the origin of these variations. We show that we are able to understand the trends presented here, and was not included. The plane wave cutoff was 350 eV for H, and 450 eV for O adsorption calculations.

**Experimental**

In Table I we include a collection of experimental data for exchange currents for hydrogen evolution. We have included both the polycrystalline data as well as data from single crystal experiments. There are some variations from one measurement to the next on the same metal, but clearly the trends are the same no matter which measurement is used.

**Volcano Curve**

We are now in a position to generate a plot of the experimental exchange currents as a function of the calculated hydrogen chemisorption energies. We use the low coverage chemisorption energies for all metals. We will refer to the effect of coverage later. The result is shown in Fig. 1. Clearly, the data order quite nicely into a volcano. Note that all the reactive metals are on the left “leg” of the volcano and the unreactive ones, Cu, Ag, and Au, are on the right.

**Kinetics and the Origin of the Volcano**

The total hydrogen evolution reaction can be written

\[
\text{H}^+ + e^- \rightarrow \frac{1}{2}\text{H}_2 \tag{2}
\]

It takes place at an electrode supplying the electrons, and providing an intermediate state of the process:

\[
\text{H}^+ + e^- + ^* \rightarrow \text{H}^* \tag{3}
\]

\[
2\text{H}^* \rightarrow \text{H}_2 \tag{4}
\]

where the * denotes a site on the surface (so an * by itself denotes a free site and H* denotes a hydrogen atom adsorbed on the surface). The final hydrogen evolution step may also be

\[
\text{H}^+ + e^- + \text{H}^* \rightarrow \text{H}_2 \tag{5}
\]

We now consider hydrogen evolution at a metal electrode, which is part of an electrochemical cell. During hydrogen evolution a current \(i\) will be running

\[
i = -er \tag{6}
\]

where \(r = r_+ - r_-\) is the net rate of (Eq. 2).

The exchange current is the forward (and backward) rate when \(i = 0\) and \(E_{\text{th}} = \Delta G^0\) is in equilibrium.
The free energy diagrams for the reaction at equilibrium over Pt(111), Ni(111), Mo(110), and Au(111) are shown in Fig. 2. At equilibrium the free energy per H atom (the chemical potential) of the initial and final states of reaction (Eq. 2) are the same. In Fig. 1 and 2 we have also included the free energy of the adsorbed state calculated as

\[ \Delta G_{H^2} = \Delta E_H + \Delta E_{ZPE} - T\Delta S_H \]

where \( \Delta E_H \) is the hydrogen chemisorption energy from Table I and \( \Delta E_{ZPE} \) is the difference in zero point energy between the adsorbed and the gas phase. We exploit the fact that the vibrational entropy in the adsorbed state is small meaning that the entropy of adsorption of \( \Delta S_H \) for \( H_2 \) is \( \Delta S_H = -1/2\Delta E_{ZPE} / T \), where \( \Delta E_{ZPE} \) is the entropy of \( H_2 \) in the gas phase at standard conditions. \( \Delta E_{ZPE} \) is calculated to be 0.04 eV for H/Cu(111), see the Appendix, and we take this value to be representative for all the metals we study here. This means that \( \Delta G_{H^2} = \Delta E_H + 0.24 \) eV.

In Fig. 2 we have chosen to show data for low coverage (1/4 monolayer). As indicated in Table I, trends will not change if we include the possibility of a higher coverage. The figure indicates that over Pt the reaction is essentially thermo-neutral where as this is not the case for metals on either side of the volcano, e.g., Ni, Au, and Mo. This in itself indicates why Pt is a much better electrocatalyst than the other metals for hydrogen evolution at the equilibrium potential.

We now turn to the kinetic modelling. We will concentrate on the trends in exchange rates from one metal to the next and on the contribution to these trends originating directly from the variations in the thermochemistry, Fig. 2 and Table II. We therefore neglect contributions to the rate constants from any reaction barriers other than those coming from the thermochemistry. This means that the variations from one metal to the next in the rate constants for getting out of the chemisorption state, \( H^2 \) in Fig. 2, are the same. Combined with the fact that at equilibrium the rates involved in the forward and backward steps must be the same, it means that the exchange rates of the two consecutive steps are the same. We choose here to focus on the rate of the proton reduction, Eq. 3. The forward rate of Eq. 3 can be written

\[ r_1 = k_1 (1 - c_{H^2}^1) \]

where \( k_1 \) is the rate constant, 0 is the H coverage (measured in fraction of a monolayer with respect to the number of metal atoms in the surface layer), and \( c_{H^2}^1 \) is the concentration of protons in the electrolyte. The coverage of hydrogen can be calculated from Eq. 4, and hence

\[ i_0 = i_{H^2} = \frac{1}{2} i_e = -n r_i \]

Table I. Calculated hydrogen adsorption energies (\( \Delta E_{H^2} \)) and measured exchange currents (\( i_e \)) on different transition and noble metals. The hydrogen chemisorption energies are calculated both for low (1/4 monolayer) and high (1 monolayer) coverage. The references to the experimental data are shown in the right column.

<table>
<thead>
<tr>
<th>Metals</th>
<th>( \Delta E_{H^2} ) (eV)</th>
<th>( \Delta E_{H^2} ) (eV)</th>
<th>(log(( i_e )/A cm(^{-2} ))</th>
<th>Ref.</th>
</tr>
</thead>
<tbody>
<tr>
<td>Au(111)</td>
<td>0.21</td>
<td>0.39</td>
<td>-6.6</td>
<td>7</td>
</tr>
<tr>
<td>Au(111)</td>
<td>0.01</td>
<td>0.39</td>
<td>-6.8</td>
<td>8</td>
</tr>
<tr>
<td>Au</td>
<td>-0.4</td>
<td>-0.33</td>
<td>-3</td>
<td>9</td>
</tr>
<tr>
<td>Ag(111)</td>
<td>0.27</td>
<td>0.34</td>
<td>-5.0</td>
<td>10</td>
</tr>
<tr>
<td>Ag</td>
<td>-0.38</td>
<td>-0.33</td>
<td>-3</td>
<td>11</td>
</tr>
<tr>
<td>Pt</td>
<td>-0.33</td>
<td>-0.27</td>
<td>-3</td>
<td>9</td>
</tr>
<tr>
<td>Pt</td>
<td>-2.63</td>
<td>0.34</td>
<td>12</td>
<td></td>
</tr>
<tr>
<td>Rh</td>
<td>-0.34</td>
<td>-0.30</td>
<td>-6.6</td>
<td>9</td>
</tr>
<tr>
<td>Rh</td>
<td>-0.21</td>
<td>-0.16</td>
<td>-3.7</td>
<td>9</td>
</tr>
<tr>
<td>Ir</td>
<td>-0.51</td>
<td>-0.47</td>
<td>-5.2</td>
<td>9</td>
</tr>
<tr>
<td>Ni</td>
<td>-0.51</td>
<td>-0.49</td>
<td>-5.3</td>
<td>9</td>
</tr>
<tr>
<td>W</td>
<td>-0.67</td>
<td>-0.33</td>
<td>-5.3</td>
<td>9</td>
</tr>
<tr>
<td>W</td>
<td>-0.51</td>
<td>-0.49</td>
<td>-5.3</td>
<td>9</td>
</tr>
<tr>
<td>Mo</td>
<td>-0.67</td>
<td>-0.77</td>
<td>-7.0</td>
<td>9</td>
</tr>
<tr>
<td>Re</td>
<td>-0.56</td>
<td>-0.45</td>
<td>-2.8</td>
<td>1</td>
</tr>
<tr>
<td>Nb</td>
<td>-0.80</td>
<td>-0.80</td>
<td>-6.8</td>
<td>9</td>
</tr>
</tbody>
</table>

\( ^a \) bcc(110)

\( \Delta G_{H^2} = \Delta E_H + \Delta E_{ZPE} - T\Delta S_H \)

\( \Delta E_{H^2} \) is the hydrogen chemisorption energy from Table I and \( \Delta E_{ZPE} \) is the difference in zero point energy between the adsorbed and the gas phase. We exploit the fact that the vibrational entropy in the adsorbed state is small meaning that the entropy of adsorption of \( 1/2 H_2 \) is \( \Delta S_H = -1/2\Delta E_{ZPE} / T \), where \( \Delta E_{ZPE} \) is the entropy of \( H_2 \) in the gas phase at standard conditions. \( \Delta E_{ZPE} \) is calculated to be 0.04 eV for H/Cu(111), see the Appendix, and we take this value to be representative for all the metals we study here. This means that \( \Delta G_{H^2} = \Delta E_H + 0.24 \) eV.

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where \( k_1 \) is the rate constant, 0 is the H coverage (measured in fraction of a monolayer with respect to the number of metal atoms in the surface layer), and \( c_{H^2}^1 \) is the concentration of protons in the electrolyte. The coverage of hydrogen can be calculated from Eq. 4, and hence

\[ i_0 = \frac{1}{2} i_e = -n r_i \]
Table II. Calculated reaction energies for the reactions Eq. 16
and Eq. 17 (ΔE = E_{ads} + n/2E_{H2} + E_{OH}), where n is 1
and 2 for Eq. 16 and Eq. 17 respectively. This is the dissociative
adsorption energy of water on different transition and noble met-
als given per water molecule. The calculated values are for the
fcc(111) surfaces and do not include spin-polarization. We also
show the adsorption free energies calculated as ΔG = ΔE
+ ΔE_{ZPE} + ΔE_{ads} for O and OH adsorption, where the zero
point energy and entropy differences are given in the Appendix.

<table>
<thead>
<tr>
<th>Metal</th>
<th>ΔE_{EH} (0.25 ML)/eV</th>
<th>ΔG_{OH} (0.25 ML)/eV</th>
<th>ΔE_{O} (0.25 ML)/eV</th>
<th>ΔG_{O} (0.25 ML)/eV</th>
</tr>
</thead>
<tbody>
<tr>
<td>Au</td>
<td>1.49</td>
<td>1.84</td>
<td>2.75</td>
<td>2.80</td>
</tr>
<tr>
<td>Ag</td>
<td>0.72</td>
<td>1.07</td>
<td>2.12</td>
<td>2.17</td>
</tr>
<tr>
<td>Pd</td>
<td>0.92</td>
<td>1.27</td>
<td>1.53</td>
<td>1.58</td>
</tr>
<tr>
<td>Pt</td>
<td>1.05</td>
<td>1.40</td>
<td>1.57</td>
<td>1.62</td>
</tr>
<tr>
<td>Rh</td>
<td>0.34</td>
<td>0.69</td>
<td>0.44</td>
<td>0.49</td>
</tr>
<tr>
<td>Ir</td>
<td>0.63</td>
<td>0.98</td>
<td>1.00</td>
<td>1.05</td>
</tr>
<tr>
<td>Ni</td>
<td>0.13</td>
<td>0.48</td>
<td>0.34</td>
<td>0.39</td>
</tr>
<tr>
<td>W</td>
<td>-0.80</td>
<td>-0.45</td>
<td>-2.06</td>
<td>-2.01</td>
</tr>
<tr>
<td>Co</td>
<td>-0.08</td>
<td>0.27</td>
<td>-0.22</td>
<td>-0.17</td>
</tr>
<tr>
<td>Cu</td>
<td>0.37</td>
<td>0.72</td>
<td>1.20</td>
<td>1.25</td>
</tr>
<tr>
<td>Mo</td>
<td>-0.61</td>
<td>-0.26</td>
<td>-1.62</td>
<td>-1.57</td>
</tr>
</tbody>
</table>

\[ \theta = \frac{K}{1 + K} \]

where

\[ K = \exp(-\Delta G_{H^+/H} / kT) = \exp(-\Delta E_H + \Delta E_{ZPE} + 1/2\beta \Delta S_{H^+/H}) / kT) \]

Let us first consider the limit where the proton transfer is exo-
ermic (ΔG_{H^+/H} < 0). In this case we would expect the rate con-
tant \( k_1 \) to be large and independent of ΔG_{H^+/H}. The rate con-
tant \( k_0 \) includes all effect relating to the reorganization of the
solvent during the proton transfer to the surface, and we are thus
assuming this to be independent of the metal at the equilibrium
potential. This leads to the following expression for the exchange
current at pH 0

\[ i_0 = -e k_0 \frac{1}{1 + \exp(-\Delta G_{H^+/H} / kT)} \]

For the other case where the proton transfer is endothermic
(ΔG_{H^+/H} > 0, e.g., Au in Fig. 2), we would expect the proton transfer
to be activated by at least ΔG_{H^+/H}. Under the same assumptions as
above, the rate constant for this case is

\[ k_1 = k_0 \exp(-\Delta G_{H^+/H} / kT) \]

and the exchange current is

\[ i_0 = -e k_0 \frac{1}{1 + \exp(-\Delta G_{H^+/H} / kT)} \exp(-\Delta G_{H^+/H} / kT). \]

The result of the model, with the single unknown parameter
\( k_0 = 200 \text{ s}^{-1} \text{ site}^{-1} \), fitted to give a reasonable overall magnitude of the
rate is included in Fig. 1. The trends are well described by the
simple model, including in particular the position of the maximum.
Clearly ΔG_{H^+/H} = 0 separates the two legs of the volcano. Such a
picture has been suggested before, and our calculated database of
chemisorption energies puts the physical picture on a firm footing.

Note that the model seems to underestimate the current density
for the metals furthest away from the maximum. This can be under-
stood in the following way. For the most noble metals Ag and Au,
steps and other low coordination number defects will have a stron-
ger metal-hydrogen bond than the close packed surfaces. These
defects will therefore be closer to the maximum and may give an
extra contribution to the current. For the least noble metals, Mo, W,
and Nb, the surface is covered by an oxide (see later) and the mea-
sured values are most probably not representative of the metal in the
metallic state.

The results indicate that we can understand the two branches of
the volcano curve in the following way: To the left of the maximum,
the rate decreases with decreasing ΔE_H due to a lack of available
sites for H + H recombination at the surface. Here hydrogen bonds
too strongly. To the right of the maximum, the rate decreases with
increasing ΔE_H because proton transfer becomes more and more
difficult as hydrogen becomes more and more unstable on the sur-
face. Here hydrogen bonds too weakly. Pt is very close to optimum,
because all reaction steps of the hydrogen evolution process on this
metal are thermo-neutral.

There has been other suggested volcano curves for the hydrogen
evolution reaction. Our volcano distinguishes itself from these in
using a systematic database for the intrinsic H-metal interaction ener-
gy as the descriptor (or x-axis). This provides an extremely simple
picture of the origin of the volcano, as outlined above.

In Fig. 1 and 2 we have used the low coverage adsorption ener-
gies from Table I. Clearly, the high coverage energies follow exactly
the same trends and thus give the same picture. However, one point
about the coverage dependence deserves further comments. In Fig. 3
we show the free energy diagram for two different coverages of
hydrogen on Pt(111). We use a low coverage value, which we take
from the \( \theta = 0.25 \) result in Table I. At \( \theta = 1 \) we use the differential
heat of adsorption as defined in Eq. 15.

\[ \Delta E_{H^+/H} (1) = 2 \Delta E_{H^+/H} (0.25) \]

Figure 3 suggests the following physical picture. At equilibrium,
the low coverage state is fully occupied. The addition of a new
proton and electron must add a hydrogen atom in the high coverage
state. Both states can turn over, but the overall activation energy is
smaller for the high coverage state, making this the most likely
candidate for the catalytically active species. This could be an
explanation of the two types of hydrogen usually invoked to explain
experimental data in the electrochemistry literature.

Water Dissociation on the Surface

One question in connection with the experimental data in Table I
and Fig. 1 is whether reaction products from the adsorption and

dissociation of water might cover some of the metals. To investigate this we have first calculated the effect of molecular water on the adsorption. The result is that adding a water overlayer changes the adsorption energy of hydrogen on Pt(111) by less than 0.02 eV. Even though the effect could be a little larger for the more reactive metals it seems justified to neglect it in the present trend study.

In addition we have calculated the reaction energies for the reactions

\[
\text{H}_2\text{O} + \text{*} \rightarrow \text{OH}^* + \text{1/2H}_2 \quad [16]
\]

and

\[
\text{H}_2\text{O} + \text{*} \rightarrow \text{O}^* + \text{H}_2 \quad [17]
\]

The results are included in Table II. In general one should also consider the possibility of forming solvated H\(^+\) and electrons at the electrode rather than H\(_2\), but for the exchange current at standard conditions this makes no difference since the solvated hydrogen is in equilibrium with H\(_2\)\,(gas). The values in Table II are relative to gas phase water. We also include the adsorption free energies, calculated using the zero point energies and entropies collected in the Appendix. The result is that of the metals considered in Table II, only Mo and W should be covered by surface oxygen under the conditions of interest here. This means that the measured exchange currents for these metals in Fig. 1 should not be taken to be representative of the metal, but rather of the surface oxide.

**Conclusions**

The results presented above clearly demonstrate the correlation between the hydrogen chemisorption energies and the exchange current for hydrogen evolution. The volcano curves in Fig. 1 also indicate that Pt is a better electrocatalyst than other metals for hydrogen evolution, primarily because the evolution reaction is thermo-neutral on Pt at the equilibrium potential. These findings can potentially be useful in predicting bimetallic electrocatalysts for hydrogen evolution, as well as for hydrogen electro-oxidation for fuel cell applications.

**Acknowledgments**

Useful discussions with J. Meier and A. K. Pal are appreciated. This work was funded in part by U.S. Army Research Office (ARO), USA, under grant DAAD19-02-1-0311 (U.S. and S.P.). The DFT calculations have been performed with support from the Danish Center for Scientific Computing through grant no. HDW-1101-05. A.L. acknowledges financial support from the European Union through contract no. ENK5-CT-2001-00572.

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**References**


**Table A-1. Zero point energy corrections and entropic contributions to the free energies.**

<table>
<thead>
<tr>
<th></th>
<th>(T\delta) (eV)</th>
<th>(T\Delta S) (eV)</th>
<th>(E_{ZPE}) (eV)</th>
<th>(\Delta E_{ZPE}) (eV)</th>
</tr>
</thead>
<tbody>
<tr>
<td>(\text{H}_2\text{O}) (0.035 bar)</td>
<td>0.67</td>
<td>0</td>
<td>0.56</td>
<td>0</td>
</tr>
<tr>
<td>(*\text{OH} + 1/2\text{H}_2)</td>
<td>0.20</td>
<td>-0.47</td>
<td>0.44</td>
<td>-0.12</td>
</tr>
<tr>
<td>(*\text{O} + \text{H}_2)</td>
<td>0.41</td>
<td>-0.27</td>
<td>0.34</td>
<td>-0.22</td>
</tr>
<tr>
<td>(1/2 \text{O}_2 + \text{H}_2)</td>
<td>0.73</td>
<td>0.05</td>
<td>0.32</td>
<td>-0.24</td>
</tr>
<tr>
<td>(\text{H}_2)</td>
<td>0.41</td>
<td>...</td>
<td>0.27</td>
<td>...</td>
</tr>
<tr>
<td>(1/2 \text{O}_2)</td>
<td>0.32</td>
<td>...</td>
<td>0.05</td>
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<td>(\text{O}^*)</td>
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<td>0.07</td>
<td>...</td>
</tr>
<tr>
<td>(\text{OH}^*)</td>
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<td>...</td>
<td>0.30</td>
<td>...</td>
</tr>
<tr>
<td>(\text{H}^*)</td>
<td>0</td>
<td>...</td>
<td>0.17</td>
<td>...</td>
</tr>
</tbody>
</table>

**Appendix**

In Table A-1 we show the entropies and zero point energies (ZPEs) used in the construction of the free energies in Fig. 2 and 3 and in Table II. The gas phase values are from Ref. 15, while the values for the adsorbed species are taken from DFT calculations for O and OH adsorbed on Cu(111) by Gokhale et al.\(^{13}\) We use gas phase \(\text{H}_2\text{O}\) at 0.035 bar as the reference state because at this pressure gas phase \(\text{H}_2\text{O}\) is in equilibrium with liquid water at 300 K.\(^{14}\) We use the same values for the adsorbed species for all the metals, as vibrational frequencies have been found to depend much less on the metal than the bond strength.\(^{15}\)