Dissolution of metal oxides in molten carbonates relates directly to the stability of materials for electrodes and construction of molten carbonate fuel cells (MCFC).

The earlier investigations on MCFC, up to 1970s, were concentrated on silver for both anode and cathode. It was shown, however, that silver was unsuitable for long-term use due to its high solubility in the carbonate melt under cathode-side conditions.\(^1\)\(^,\)\(^2\)

NiO, in the lithiated form for high conductivity, still seems to be the most promising material for the MCFC cathode. It is, however, unlikely to attain a lifetime more than 40,000 h because of its solubility in the carbonate melt. Several investigations have been carried out to measure the solubility of NiO in individual alkali carbonates,\(^3\)\(^,\)\(^4\) binary mixtures,\(^5\)\(^,\)\(^6\) ternary mixtures,\(^7\)\(^,\)\(^8\) and alkali/alkaline earth carbonate mixtures.\(^9\)\(^,\)\(^10\)

In order to obtain other materials as possible substitutes for the lithiated NiO cathode, extensive research has been conducted. Among the numerous candidates of pervoskite-type compounds and mixed metal oxides, the chemistry of CuO\(^11\)\(^,\)\(^12\), ZnO\(^12\)\(^,\)\(^13\), LiFeO\(_2\)\(^14\)\(^,\)\(^15\), (Li,K)\(_2\)CrO\(_4\)\(^15\), LaNiO\(_3\)\(^12\)\(^,\)\(^13\), Li\(_2\)MnO\(_3\)\(^13\)\(^,\)\(^17\), LiCoO\(_2\)\(^16\)\(^,\)\(^19\), and La-ZrO\(_2\)\(^20\) has been investigated in the alkali carbonate mixtures in the temperature range between 550 and 950°C.

Dissolution of metal oxides in molten alkali carbonates is believed to be controlled by the acid/base chemistry of the solvent. The acid/base concept, postulated for the oxyanion-containing molten salt systems by Lux\(^21\) and Flood and Forland,\(^22\) is analogous to the Brønsted concept for aqueous solutions. The oxygen ion, O\(^2-\), is considered to be the basic component of the solvent in the same sense that the hydrogen ion defines acidity in aqueous solutions. Alkali metal carbonates dissociate according to the following equilibrium

\[
\begin{align*}
M_2CO_3(s) & \leftrightarrow 2M^{+} + CO_3^{2-} \\
CO_3^{2-} & \leftrightarrow O^{2-} + CO_2(g)
\end{align*}
\]

where CO\(_3^{2-}\) is identified as the basic component and the gaseous CO\(_2\) (which is presumed to be in equilibrium with that in the melt) as the acidic component. The basicity/acidity of the melt can therefore be defined by a function of either -log \(aO_2^-\) or -log \(pCO_2\).

Transition metal oxides, e.g., nickel oxide NiO, are found to dissolve in the carbonate melt either in an acidic mode or in a basic mode. When the oxygen ion activity in the melt is very low, corresponding to a high partial pressure of CO\(_2\), the acidic dissolution is supposed to occur

\[
\text{NiO} \leftrightarrow \text{Ni}^{2+} + O^{2-}
\]

or

\[
\text{NiO} + \text{CO}_2 \leftrightarrow \text{Ni}^{2+} + \text{CO}_3^{2-}
\]

At a high oxide ion activity or a low partial pressure of CO\(_2\), the basic dissolution occurs by reacting with oxide ions to form complex anions such as NiO\(_2\)\(^2-\).

\[
\text{NiO} + O^{2-} \leftrightarrow \text{NiO}_2^{2-}
\]

or

\[
\text{NiO} + CO_3^{2-} \leftrightarrow \text{NiO}_2^{2-} + CO_2
\]

NiO\(^-\) or NiO\(_2\)\(^-\), as suggested by Orfield and Shores,\(^3\) may also be the products of the basic dissolution, where the formal oxidation states are +3 and +4, respectively. The mechanism of basic dissolution has not been elucidated since little is known about the stability of various nickelate ions. It is however the acidic dissolution that governs the process under the cathodic conditions of an MCFC.

If the activity coefficient of the dissolved nickel species is assumed to be constant, a plot of the logarithm of the NiO solubility as a function of the melt basicity, e.g., -log \(pCO_2\), would show two distinct regimes of dissolution. According to reaction 3, the acid dissolution will be represented as a straight line with a slope of +1.

For single molten carbonates of alkali metals, it is reported that the solubility in both acidic and basic regimes increases in an order from Li\(_2\)CO\(_3\) to Rb\(_2\)CO\(_3\).\(^3\) The slopes of the basic dissolution lines decrease from -1.0 for Li\(_2\)CO\(_3\) and Na\(_2\)CO\(_3\) to -0.5 for Rb\(_2\)CO\(_3\). A corresponding trend in the slopes of the acidic solubility also exists but does not appear to follow the chemistry of the solvent. The values of the slopes are found to be +1.0 for Na\(_2\)CO\(_3\), +1.4 for K\(_2\)CO\(_3\), and +1.2 for Rb\(_2\)CO\(_3\). For Li\(_2\)CO\(_3\), however, no acidic dissolution is observed.

In a mixture of two or more alkali carbonates, the melt basicity is a function of both the \(pCO_2\) above the melt and the mole fraction of the alkali cations. In a binary mixture of Li\(_2\)CO\(_3\) and K\(_2\)CO\(_3\), for example, the basicity is a function directly proportional to the \(pCO_2\) when the composition is fixed. If the partial pressure of CO\(_2\) is fixed, the basicity is, however, a nonlinear function of the mole fractions of the melt components.\(^6\)

In binary mixtures of Li-K, Li-Na, and Na-K carbonates, both acidic and basic dissolution of NiO has been observed. Orfield and Shores\(^6\) reported that the Li\(_2\)CO\(_3\)-rich mixtures (higher than 50% mol Li\(_2\)CO\(_3\) at 910°C) showed only the basic dissolution, while the results from Ota et al.\(^7\) and Doyon et al.\(^10\) for the Li/K = 62/38 melt showed a well-defined acidic dissolution in the temperature range from 650 to 750°C. The temperature difference may partly account for this by considering the fact that the free energy of dissociation of an alkali carbonate depends upon the temperature of the melt and the dissolution of metal oxides depends on the basicity.

\* Electrochemical Society Active Member.
The assumption of a simple acid dissolution (reaction 3) seems to be valid, since Doyon et al.\textsuperscript{10} and Ota et al.\textsuperscript{7} obtained slopes of acid dissolution lines around +1. Furthermore Doyon et al.\textsuperscript{10} and Plomp et al.\textsuperscript{11} demonstrated that the addition of basic oxides (MgO and SrO) into the Li/K = 62/38 melt led to a decrease in solubility of NiO. At different compositions of the Li/K binary carbonates, the NiO solubility was found to decrease with an increase in the content of lithium salt. The similarity of the lithium effect to that of MgO and SrO led some workers\textsuperscript{6,7,10} to conclude that Li\textsubscript{2}CO\textsubscript{3} is the most basic salt among the three alkali carbonates (basic Li\textsubscript{2}CO\textsubscript{3} > Na\textsubscript{2}CO\textsubscript{3} > K\textsubscript{2}CO\textsubscript{3} acidic).

On the other hand, in addition to the variant slopes for single carbonates,\textsuperscript{3} Orfield and Shores\textsuperscript{6} reported acid dissolution slopes between 0.65 and 0.80 for all compositions of Na\textsubscript{2}CO\textsubscript{3}-K\textsubscript{2}CO\textsubscript{3} and Li\textsubscript{2}CO\textsubscript{3}-K\textsubscript{2}CO\textsubscript{3} with an exception of the 0.96 K\textsubscript{2}CO\textsubscript{3}-0.04 Li\textsubscript{2}CO\textsubscript{3} melt, where the slope was found to be close to 1.

Orfield and Shores\textsuperscript{6} suggested that the variant slopes of the acid dissolution lines may best be correlated with the difference between the temperature of experiments and the melting point of the melts. When the temperature difference is larger than 10°C, the slope was found to be constant for all mixtures regardless of components or compositions, but being around 0.7 instead of the expected 1.0. They calculated the basicity of the melt for different temperatures and replotted the solubility data as a function of this calculated basicity [at constant \(p_{\text{CO}_2}\)]. The slope, however, less than 0.2, being very different from those of the solubility vs. basicity lines at a constant temperature but different \(p_{\text{CO}_2}\).

Moreover, when Orfield and Shores\textsuperscript{6} replotted the solubility data of NiO at different Li\textsubscript{2}CO\textsubscript{3} contents as a function of the calculated basicity (at a fixed \(p_{\text{CO}_2}\) pressure), they found that, although the basicity varies over four orders of magnitude with an addition of 10% Li\textsubscript{2}CO\textsubscript{3} to K\textsubscript{2}CO\textsubscript{3}, the solubility remains almost identical to that of pure K\textsubscript{2}CO\textsubscript{3}. With larger additions of Li\textsubscript{2}CO\textsubscript{3}, as the composition approaches pure Li\textsubscript{2}CO\textsubscript{3}, the basicity changes by only one order of magnitude, but the solubility of NiO decreases considerably.

The above arguments evidently simplify the mechanism of the acid/base dissolution of metal oxides in alkali carbonates. As Orfield and Shores\textsuperscript{6} proposed, the formation of complex ions might be one of the explanations. This paper is devoted to an investigation of the solubility of metal oxides from the viewpoint of the complex formation.

**Experimental**

The solubilities of PbO, NiO, Bi\textsubscript{2}O\textsubscript{3}, and Fe\textsubscript{2}O\textsubscript{3} were measured in (Li\textsubscript{x}K\textsubscript{1-x})\textsubscript{2}CO\textsubscript{3} melts with \(x\) between 0.43 and 0.62. The alkali carbonates, Li\textsubscript{2}CO\textsubscript{3} (Riedel-De Haen, analytic) and K\textsubscript{2}CO\textsubscript{3} (Merck, analytic), were ground at room temperature and mixed for 72 h before melting at 650°C. Powdered of PbO (99.9% Aldrich) and Bi\textsubscript{2}O\textsubscript{3} (99.8% Johnson Matthey) were first melted in air, cooled, and crushed to obtain oxide pellets. The Fe\textsubscript{2}O\textsubscript{3} and NiO powders (both 99% from Aldrich) were sintered as pellets at 800°C in air. The metal oxide pellets were then placed in an alumina crucible (id 4 cm and height 6.5 cm) and covered with about 100 g of the premixed carbonates. The mixture was heated to 650°C under an atmosphere of pure CO\textsubscript{2}. The gas (either pure CO\textsubscript{2} or CO\textsubscript{2}-Ar mixture) was then bubbled through the melt by means of an alumina inlet tube to ensure that the melt was saturated with the gas. Gaseous CO\textsubscript{2}, argon, and a premade mixture of 0.50 (±0.005%) CO\textsubscript{2} in Ar were provided by Hede Nielsen A/S, and from these the mixed gas of other compositions was obtained by means of mass-flow meters and a controller (Bronkhorst, HI-TEC E-5514). Initial tests were conducted for each oxide to determine the time needed for reaching the equilibrium at the studied experimental conditions.

The alumina inlet tube was lifted just above the melt for an hour before the sampling took place in order to allow any particles to settle. Melt samples (5 to 6 g) were taken by means of a quartz tube from the melt just under the upper surface. The samples containing different oxides were then dissolved in diluted acetic acid. After evaporation, the remains were dissolved in 0.1 M nitric acid and analyzed by atomic absorption spectroscopy (Perkin Elmer 2100). Due to its high solubility, PbO was analyzed gravimetrically by precipitation as PbCrO\textsubscript{4}.

An electrochemical cell

\[
\text{Pb \bigg| CO}_2, \text{PbO}_{(\text{sat.})}, \text{Li-K}_2\text{CO}_3(1) \parallel \text{Li-K}_2\text{CO}_3(2) \bigg| \text{PbO}_{(\text{sat.})}, \text{Ar-CO}_2, \text{Pb}
\]

was constructed for EMF measurements. Here the melt composition in the left side was kept at \(x_{\text{Li}} = 0.43\), while the melt composition in the right side varied from \(x_{\text{Li}} = 0.43\) to 0.62. Both melts were saturated with lead oxide under variant partial pressures of carbon dioxide. Liquid lead was used as electrodes in both chambers with pure iron as the connection wire. The apparatus is shown in Fig. 1. The connection between the two half-cells was made by cutting the side wall of both tubes and blocking the resulting holes with porous alumina. The temperature in the melt was measured with a Pt/Pt-10% Rh thermocouple protected by an alumina tube.

Results and Discussion

**Solubility measurements.**—Figure 2 shows the dissolution curves of PbO under an atmosphere of pure CO\textsubscript{2} at 650°C in lithium-potassium melts with lithium mole fractions of 0.43, 0.51, and 0.62,
The solubility of metal oxides in molten carbonates, as previously believed for NiO, the lithium effect will merely change the activity of the oxide in the melt, and the oxide activity, in turn, affects the dissolution of metal oxides via an acid/base equilibrium. This, however, does not seem to be the case, especially for the Fe₂O₃ and Bi₂O₃ dissolution.

It is also interesting to notice that the solubility of the two bivalent metal oxides (NiO and PbO) behaves the same way with respect to the lithium partial pressure as the Fe₂O₃ and Bi₂O₃ dissolution. This can be either coincidental or indicating a connection, for example, to the oxidation states of the metals. However, no further investigation was carried out in the present work.

Orfield and Shores⁶ have made a measurement of NiO solubility in melts of different Li/K compositions at 910°C. For the acidic solubility (pCO₂ = 0.75) in melts containing lithium carbonate less than 50 mol %, they obtained a straight line when plotting log(x_NiO) against log(aₖCO₃). It is, however, interesting that a straight line is also obtained when the log(x_NiO) is plotted against the mole fraction of lithium carbonate in the range from 0.04 mol % to 50 mol % Li₂CO₃. Ota et al.⁷ made a measurement of NiO solubility in melts of different Li/K compositions under pCO₂ = 1 atm in a temperature range from 600 to 750°C. One set of their solubility data (from 750°C) is replotted as a function of the CO₂ partial pressure, together with the solubility data for PbO at 650°C from the present study (Fig. 4). The dependence of PbO and NiO solubility upon the CO₂ partial pressure is very much the same, i.e., the solubility decreases with decreasing pCO₂ in the high-pressure range but increases with decreasing pCO₂ in the low pressure range. The minimum solubility for NiO and PbO appears at nearly the same partial pressure of carbon dioxide.

For the acidic dissolution at pCO₂ higher than 0.1 atm, the slopes of the log(x_NiO) vs. log (pCO₂) are claimed to be around unity by Ota et al.⁷ and Doyon et al.¹⁰ as well. For PbO from the present work, however, a linear regression of log(x_NiO) vs. log(pCO₂) gives the acidic slopes between 0.88 and 0.91 for three compositions of the melt. The slopes of the basic solubility lines for PbO, when the CO₂ pressure is lower than 0.01 atm, are found to be about −0.3, similar to that for NiO solubility reported by Ota et al.⁷ and Doyon et al.¹⁰

**EMF measurements.**—The EMF results are shown in Fig. 5. With the melt of x_Li = 0.43 and pCO₂ = 1 as reference in the left chamber of the cell, the measured EMF values vary with the mole fraction of lithium carbonate in the melt and the CO₂ partial pressure of the gas phase of the right chamber. For a lithium mole fraction from 0.43 to 0.51 and 0.62 under pCO₂ = 1, the EMF increases by 22 and 39 mV, respectively.

When, however, the partial pressure of carbon dioxide decreases, the measured EMF decreases (see Fig. 6). In the x_Li = 0.51 melt, for example, a decrease in pCO₂ from 1 atm to 0.25 and 0.1 atm leads to a decrease of the EMF by 39 and 85 mV, i.e., from 22 to −17 mV and −63 mV, respectively.

For an approximate evaluation of the liquid-junction potential, Andersen²³ suggested an expression using the ionic mobility instead of the diffusivity.
of the transference number, the former being less sensitive to concentration. Consider the liquid junction potential

$$\text{Li}_2\text{CO}_3 - \text{K}_2\text{CO}_3(1) \parallel \text{Li}_2\text{CO}_3 - \text{K}_2\text{CO}_3(2)$$

Assuming that the system mainly consists of Li$^+$, K$^+$, and CO$_3^{2-}$, and their mobilities are invariable over the studied concentration range, and taking the carbonate anion as the velocity reference $u(CO_3^{2-}) = 50$, one has

$$\Delta\varphi = \frac{RT}{nF} \int \left( u_1 dX_1 + u_2 dX_2 \right) / \left( u_1 X_1 + u_2 X_2 \right)$$

where $u_1$, $u_2$ are the mobilities, and $X_1$, $X_2$ are the mole fractions of Li$^+$ and K$^+$, respectively. The mole fraction range studied in the present work is from $x_{Li} = 0.43$ to 0.62. At 923 K for $u_1 = 2u_2$, one has $\Delta\varphi < 5$ mV, small compared with the measured EMF values.

**Model for complex formation.**—As seen from the solubility and EMF measurements, any decrease in carbon dioxide partial pressure results in a decrease in the EMF value and in the solubility of the metal oxide. Keeping the solvent equilibrium (reaction 2) in mind, one expects a higher oxide activity as carbon dioxide partial pressure decreases. This, in turn, corresponds to a more negative value of EMF.

On the other hand, an increase in the lithium mole fraction of the melt leads to a decrease in the solubility of the metal oxide but an increase in the EMF value. This contradiction, compared with the effect of the carbon dioxide partial pressure, may suggest a more complicated mechanism than the previously assumed acid/base chemistry for metal oxide dissolution in molten carbonates. We try to understand the phenomena by applying the assumption of complex formation to the dissolution of metal oxides.

In order to relate the solubility and EMF data to the assumption of complex formation, the following model is suggested by considering the mechanism of the dissolution of a divalent metal oxide

$$\text{MO} + \text{CO}_2 + (n-1) \text{CO}_3^{2-} \rightleftharpoons [\text{M(CO}_3]_{n-2}^{2n+2}$$

where M is a divalent metal and n is a number to be determined. With $n = 1$, the formula corresponds to the formation of a simple carbonate of lead, while $n > 1$ corresponds to the formation of negatively charged complex ions.

Here it is assumed that the metal oxide is not involved in redox reactions as far as the acidic dissolution is concerned, even though this may not be true when the basic dissolution is present. Several measurements$^{6,7,10}$ have shown that the solubility of transition metal oxides (e.g., NiO) is independent of the partial pressure of oxygen. Therefore the involvement of gaseous oxygen, being either produced or consumed during the dissolution of metal oxides, is not considered.

Furthermore, the formation of any LiMO$_x$ compound from the reaction between metal oxides and Li$_2$CO$_3$ will only affect the model when the CO$_2$ pressure is varied.

Considering the linear relationship between the metal oxide solubility and carbon dioxide partial pressure in the acidic region, one would expect the formed complex in a monomeric form. In addition...
the coordination number, $n$, should be expected to be of low value so that an unlikely high charge number ($-2n+2$) of the complex ion can be avoided.

By further assuming the activity coefficient of the dissolved metal oxide as unity, a relationship between the oxide solubility ($x_{PbO}$), carbon dioxide pressure ($p_{CO_2}$), and carbonate activity can then be expressed as

$$K_p = \frac{x_{PbO}^2}{p_{CO_2} a_{CO_3}^{-n}}$$  \[9\]

With consideration of the solvent equilibrium (reaction 2), with the dissociation constant

$$K_d = \frac{p_{CO_2} a_{CO_3}}{a_{PbO}}$$  \[10\]

the metal oxide solubility can also be related to $p_{CO_2}$ and oxide activity

$$K'_p = \frac{x_{PbO}^2}{p_{CO_2} a_{CO_3}^{-n}}$$  \[11\]

At a fixed melt composition ($x_j$), the oxide activity and carbon dioxide partial pressure will be in inverse proportion when the carbonate activity stays constant. To estimate the anion concentration in the mixture, the dissociation constant, $K_d$, for the Li/K mixed carbonate was taken from Andersen,\textsuperscript{23} who obtained it by means of thermodynamic EMF measurements. For the Li/K = 50/50 melt at 650°C he gives log $K_d = -7.51$. This corresponds, at $p_{CO_2} = 1$ atm, for example, to the activity ratio of oxide ion to carbonate anion of $10^{-8}$.

In the EMF cell, the lead electrode is in equilibrium with the dissolved metal ions, i.e., the complex Pb(CO$_3$)$_2$-$n$+2. The cell reactions can therefore be written as

anodic reaction: $Pb + nCO_3^{2-}_{(M)} \rightleftharpoons PbCO_3^{2-}_{(R)} + 2e^-$  \[12\]

and cathodic reaction: $PbCO_3^{2-}_{(M)} + 2e^- \rightleftharpoons Pb + nCO_3^{2-}_{(M)}$  \[13\]

overall reaction: $PbCO_3^{2-}_{(M)} + 2e^- \rightleftharpoons Pb + nCO_3^{2-}_{(R)} + 2CO_2$  \[14\]

The EMF of the cell can be expressed as

$$E = (RT/2F) \ln \left[ \frac{x_{PbO}^2}{x_{PbO}^{M}} \right] \frac{a_{CO_3}^{2-}_{(M)} / a_{CO_3}^{2-}_{(R)}}{p_{CO_2}}$$  \[15a\]

Here subscripts $R$ and $M$ refer the reference and measurement chambers, respectively. As can be seen, when both melts are saturated with lead oxide, the cell EMF is directly related to the oxide solubility, ($x_{PbO}^R$), and the carbonate activity, $a_{CO_3}^{2-}$, the latter with a power of $n$. By combining the solvent dissociation constant $K_d$, the EMF can also be expressed as

$$E = \left[ \frac{RT}{2F} \ln \left( \frac{x_{PbO}^2}{x_{PbO}^{M}} \right) \frac{p_{CO_2}^{3/2}}{p_{CO_2}^{M/2}} \right] \frac{a_{CO_3}^{2-}_{(M)}}{a_{CO_3}^{2-}_{(R)}}$$  \[15b\]

From the measured oxide solubility and EMF values at different $CO_2$ partial pressures and lithium contents ($x_j$), using Eq. 15, the variation of carbonate ion activity or oxide ion activity can be calculated for different $n$ values.

Figure 7 shows the calculated ratio of oxide ion activity ($a_{CO_3}^{2-}_{(M)}/a_{CO_3}^{2-}_{(R)}$) in a melt of fixed Li/K (= 43/57) composition but different $CO_2$ partial pressures. It should be remarked that the reference side of the EMF cell is with a melt of Li/K = 43/57 and $p_{CO_2} = 1$ atm. The oxide activity ratio of the measuring-side melt to the reference-side melt is unity when $p_{CO_2}$ of the measuring side is also 1 atm. As the $CO_2$ partial pressure decreases, the oxide ion activity in the measuring-side melt increases accordingly. The solubility of PbO will also decrease, as seen from the plots in Fig. 4. The difference of the measured EMF value and the term $(RT/2F) \ln \left[ x_{PbO}^2/x_{PbO}^M \right]$ will be accounted for by the pressure and activity terms in Eq. 15b. By assuming the number $n = 1, 2, 3$, the corresponding oxide activity ratio is obtained as a function of the $CO_2$ partial pressure, as shown in Fig. 7. The solid line is calculated from the solvent equilibrium (Eq. 10). The value of the coordination number used is indicated in the figure.

Fig. 7. The calculated oxide activity ratios $a_{CO_3}^{2-}_{(M)/a_{CO_3}^{2-}_{(R)}}$ as a function of carbon dioxide partial pressure for different coordination number. The reference chamber (R): Li/K = 43/57, $p_{CO_2} = 1$ atm; the measuring chamber (M): Li/K = 43/57, varying $p_{CO_2}$. Temperature: 650°C. The solid line is calculated from the solvent equilibrium (Eq. 10). The value of the coordination number used is indicated in the figure.

Evidence for complexes of transition metal and other metal ions in molten sulfate, chloride, and fluoride melts is well known.\textsuperscript{24} In molten alkali carbonates, there is obviously a competition between the alkali metal and the transition metal ions for carbonate anions. The result of the competition will of course depend on the relative attracting power of the metal ion. Lithium ion with the smallest
radius, is of the greatest polarizing power and would render the complex ions unstable. This will help to explain the lithium effect on the solubility of metal oxides.

From the very similar solubility behavior of NiO and PbO, one might expect a similar complex chemistry for both nickel oxide and lead oxide. As discussed above, the lithium effect on the NiO solubility can be virtually understood with the help of the assumed complex formation. Moreover, the stability of complex ions is not affected only by the acid/base chemistry of the melt, and therefore the variant slopes of solubility-basicty lines, as found in measurements by Orfield and Shores, are not completely surprising.

Additionally, the temperature coefficient of the NiO solubility was reported to be negative for the acidic dissolution in a wide range of the melt composition. When plotting the slopes of acid dissolution lines against the temperature difference between the melt temperature and the melting point of the solvent, Orfield and Shores found that the slope decreases from about 1.5 to about 0.7 as the temperature difference increases up to 100°C. When the temperature difference exceeds 100°C (in the range from 100 to 400°C), the slope becomes constant with a value of about 0.7, regardless of components or composition of carbonate melts. This behavior is also expected by considering the fact that the thermal stability of the complex ions is generally temperature dependent. The physical structure of the ionic melt and complex ions will probably change as the temperature increases.

During the operation of molten carbonate fuel cells, the dissolution product of the nickel oxide cathode is assumed to migrate toward the negative anode and to be reduced in the vicinity of the anode, resulting in the so-called short-circuiting. In this connection, a simply positively charged metal ion (Ni$^{2+}$) from the dissolution of a NiO cathode seems more likely if only electromigration is considered. However, in light of the Nernst-Planck flux equation, a migration of a negatively charged ion toward the negative anode can occur when the concentration gradient is large enough. This is well known for electrochemical deposition of negatively charged complex species in, e.g., aluminum and magnesium electrolysis.

**Conclusions**

Solubilities of PbO, NiO, Fe$_2$O$_3$, and Bi$_2$O$_3$ in molten Li/K carbonates have been measured at 650°C under different pressure of carbon dioxide. As the lithium molar fraction increases from 0.43 to 0.62 in the melt, the solubilities of NiO and PbO decrease while those of Fe$_2$O$_3$, and Bi$_2$O$_3$ remain approximately constant. At a fixed composition of the melt, NiO and PbO dissolve in the carbonates either in an acidic or a basic mode, as the partial pressure of carbon dioxide varies.

Both decreasing CO$_2$ pressure and increasing lithium content result in a reduction in the PbO solubility in the melt. On the contrary, the EMF measurement shows opposite effects, i.e., decreasing CO$_2$ pressure leads to more negative EMF values but increasing lithium content gives more positive EMF values. This contradiction is explained by the assumption of complex formation upon the dissolution of metal oxides in alkali carbonate melts.

By means of a model of complex formation, the following reactions are suggested for the dissolution of lead oxide in molten carbonates

$$\text{PbO} + 2\text{CO}_2 + \text{CO}_3^{2-} \rightleftharpoons [\text{Pb(CO}_3)_2]^{4-}$$

and/or

$$\text{PbO} + 2\text{CO}_2 + 2\text{CO}_3^{2-} \rightleftharpoons [\text{Pb(CO}_3)_3]^{4-}$$

A similar chemistry for nickel oxide might be expected, which explains the unusual solubility behavior of NiO in carbonate melts, i.e., the decreased solubility with increasing lithium content.

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**References**


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Figure 8. The calculated oxide activity ratios $a_{\text{O}_2}/a_{\text{O}_2}$ as a function of the coordination number. The reference chamber (R): Li/K = 43/57, $p_{\text{CO}_2} = 1$ atm; the measuring chamber (M): Varying Li/K composition, $p_{\text{CO}_2} = 1$ atm. Temperature: 650°C. The solid lines ( ) are from the solubility data, and the dashed lines ( ) are from the solubility and EMF measurements. The lithium molar fraction of melts is indicated in the figure.