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Electrochemical Promotion of Oxidative Coupling of Methane on Platinum/Polybenzimidazole Catalyst

I. M. Petrushina,* V. A. Bandur, N. J. Bjerrum,* F. Cappeln, and L. Qingfeng*

Materials Science Group, Department of Chemistry, Technical University of Denmark, Lyngby DK-2800, Denmark

The electrochemical promotion of catalytic methane oxidation was studied using a (CH₄, O₂, Ar), Pt|polybenzimidazole (PBI)-H₃PO₄|Pt, (H₂, Ar) fuel cell at 135°C. It has been found that C₂H₂, CO₂, and water are the main oxidation products. Without polarization the yield of C₂H₂ was 0.9% and the yield of CO₂ was 7.3%. This means that C₂ open-circuit selectivity was approximately 11%. Open-circuit voltage was around 0.6 V. It has been shown that the CH₄ → C₂H₂ catalytic reaction can be electrochemically promoted at negative polarization and exhibits a clear “volcano-type” promotion behavior, meaning that there was a maximum promotion effect at a polarization of -0.15 V, or 0.45 V catalyst potential vs. a hydrogen electrode (3.8% C₂H₂ yield). The catalytic rate enhancement ratio, $r(\text{C}_2)/r_0(\text{C}_2)$, at this maximum was 4.2. There was no C₂H₂ production at polarization ≥ 0.1 and ≤ -0.3 V. The yield of C₂H₂ decreased with decreasing temperature. Dependence of CO₂ yield on polarization also showed a “volcano-type” behavior with maximum yield of 8.3% at -0.15 V polarization. The catalytic rate enhancement ratio for CO₂ production, $r(\text{CO}_2)/r_0(\text{CO}_2)$, at this maximum was 1.1, which means that this catalytic reaction is only slightly affected by the electrochemical polarization. This indicates that polarization especially affects the C₂ selectivity of the catalyst. The obtained data was explained by the electrochemical production of Pt-H active centers at the electrolyte-catalyst-gaseous reactant interface ($\lambda \gg 1$).

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The fact that a catalytic system often consists of an electronic and an ionic conductor makes it possible to use electrochemical techniques to study the nature of the promotion effect in catalysts. Application of galvanic cells with solid electrolytes for the study of catalytic phenomena was suggested by Wagner.¹ It was found later, that the rate and selectivity of a catalytic reaction can be remarkably and reversibly changed by electrochemical polarization of the electronic conductive catalyst on the ionic conductive support.² This phenomenon was discovered by Vayenas *et al.* and named the “NEMCA effect” (non-faradaic electrochemical modification of catalytic activity).² The term “electrochemical promotion” was first proposed by Pritchard³ in his response to Vayenas’s paper.²

The NEMCA effect has been demonstrated for more than 50 heterogeneous³⁻¹³ and one homogeneous¹⁴ catalytic reactions. Thus, for example, it was shown that CH₄ oxidation to CO₂ on the Pt/yttria-stabilized zirconia (YSZ) catalyst can be electrochemically promoted up to five times.⁴ In the case of CH₄ oxidative coupling on Ag/YSZ catalyst (with C₂H₄, C₂H₆, and CO₂ as the products), the CH₄ conversion increased about 30% after positive polarization of the catalyst, while C₂ selectivity increased at negative polarization.⁴

According to Vayenas *et al.*,⁴⁻⁶ the NEMCA effect can be described by the following parameters
The rate enhancement ratio

$$\rho = r/r_0 \quad [1]$$

and the enhancement factor

$$\lambda = (r - r_0)/(I/2F) \quad [2]$$

where r is the catalytic rate at current I , r_0 is the open-circuit catalytic rate, and F is Faraday’s constant. In all studied cases of the electrochemical promotion $|\lambda| \gg 1$, and that was the reason why this effect has been called “non-faradaic.”

Vayenas *et al.*⁴⁻⁶ explained the electrochemical promotion of a catalyst as an effect of changes in the work function, $e\Delta\Phi$ (e is the electron charge, $\Delta\Phi$ is change of the electron extraction potential), of the catalyst under polarization. The change of the electron extraction potential has been assumed to be equal to the applied polarization, *i.e.*

$$\eta = \Delta\Phi \quad [3]$$

It was also suggested that this change in the work function is caused by spillover of the products of the charge transfer from the three-phase boundary to the catalyst-gas interface. The change in the work function was measured using the Kelvin probe technique.⁴⁻⁶

However, it was found later⁹⁻¹¹ that the work function change of catalysts with different morphology, measured by the use of a Kelvin probe, may be only a part of the applied polarization^{9,10} or does not change at all.¹¹ Summarizing the results of Ref. 4-6 and 9-11, one can agree with the assumption of Metcalfe¹⁰ that

$$\Delta\Phi = \xi\eta \quad [4]$$

where ξ is a coefficient. Furthermore, one can also assume $0 \leq \xi \leq 1$.

It is known that the electrochemical potential of an electron in a metal $\bar{\mu}_e$ (Fermi level) is a sum of the chemical potential of the electron in the metal μ_e and $e\phi$, where ϕ is the inner or Galvani potential. The Galvani potential is a sum of an outer (or Volta) potential Ψ and a surface potential χ . The work function of the electron, $e\Delta\Phi$, is equal to the electrochemical potential of uncharged metal ($\Psi = 0$),¹⁵ *i.e.*

$$\Delta\Phi = \Delta\mu + \Delta\chi \quad [5]$$

and $\Delta\Psi \neq 0$ is the reason for the difference between $\Delta\Phi$ and η (Eq. 4).

Taking into account the considerations described and the published data on the electrochemical promotion,²⁻¹⁴ we can define three types of electrochemical promotion depending on values of the faradaic current (I) and λ

1. The electrochemical promotion effect is the charge-induced (electrostatic) change of the strength of chemisorptive bonds at the catalyst-reactant interface (CI-effect). The CI-effect is characterized by $I = 0$. Charging of the electric double layer at the catalyst-support interface is the only result of the polarization of the catalyst. $\Delta\Psi \neq 0$ and $\Delta\Phi \neq \eta$.

2. The electrochemical promotion effect is induced by the electrochemical production of the catalyst promoters (EPP-effect). $I \neq 0$; $\lambda \gg 1$; $\Delta\Psi \approx 0$; and $\Delta\Phi \approx \eta$.

* Electrochemical Society Active Member.

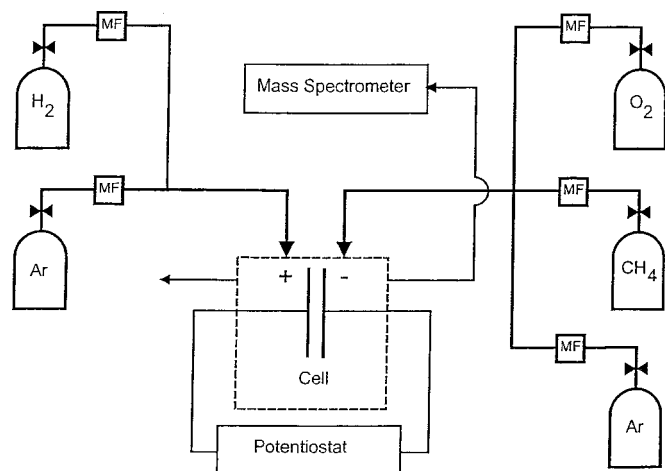


Figure 1. Setup for study of electrochemical promotion of catalytic methane oxidation.

3. The case of electrocatalysis is when reactant (or reactants) of a catalytic reaction are oxidized or reduced electrochemically. $I \neq 0$, $\lambda \approx 1$.

The CI-effect usually plays an important role in the electrochemical promotion of homogeneous catalytic reactions using non-catalytic electrodes.¹⁴ Most of the published cases of electrochemical promotion²⁻¹³ can be ascribed to the EPP-effect or mixed CI- and EPP-effects. Electrocatalytic reactions are used in many important technologies, *e.g.*, fuel cells, water electrolysis, etc.

In the present research the electrochemical promotion of catalytic methane oxidation was studied using a



fuel cell.¹⁶ The aim of this study was to create electrochemically a new alternative to the $\text{CH}_4\text{—CO}_2$ route for the catalytic reaction through the electrochemical production of H-Pt active centers, *i.e.*, using the EPP-effect.

The PBI- H_3PO_4 electrolyte was chosen because it is stable at temperatures higher than 100°C , at which CO poisoning of the Pt catalyst is of little importance.¹⁶

Experimental

The setup for study of electrochemical promotion of the catalytic methane oxidation is given in Fig. 1. A carbon-supported Pt catalyst with Pt load around 0.5 mg/cm^2 was used in the present investigation.¹⁶ The morphology of the catalyst and the structure of the three-phase (Pt-electrolyte-gas) boundary are given in Fig. 2. The electrolyte was PBI doped with H_3PO_4 . The techniques of preparing the PBI(H_3PO_4) membrane electrolyte and the Pt, C/PBI (H_3PO_4)/Pt,C assembly (the working electrode area is 5 cm^2) were developed in our group as well as by others.¹⁷ Graphite plates with gas channels were used as holders and current collectors. Two aluminum end plates with attached heaters were used to clamp the graphite plates. Temperature was controlled by a homemade controller. Mass flowmeters (5850 S Brooksmart) and 0154 Brooks Instrument controllers were used to control the inlet gas composition. The outlet gas composition was measured by an on-line quadrupole mass spectrometer (QMS 421, Pfeiffer) with a secondary electron multiplier detector and a cross-beam ion source with a rhenium filament. The ionization voltage was 70 V. The outlet gases were admitted to the mass spectrometer from the fuel cell through a 0.8 mm stainless steel capillary.

The mixture of methane and oxygen, diluted by Ar ($\text{CH}_4/\text{O}_2/\text{Ar} = 34/17/70 \text{ mL/min}$), was used as a working mixture at one electrode, and hydrogen or hydrogen/argon mixture (usu-

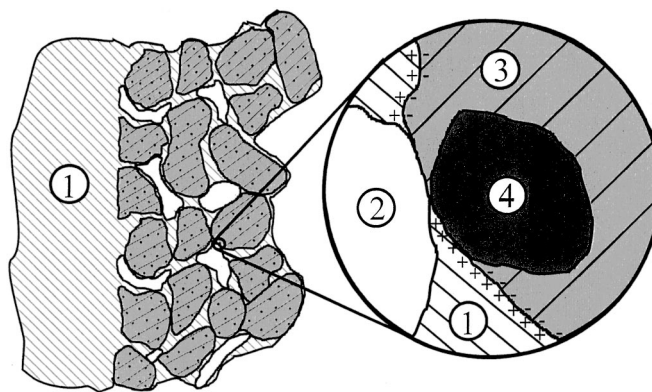


Figure 2. Reactant-catalyst-electrolyte boundary: (1) PBI- H_3PO_4 electrolyte; (2) $\text{CH}_4/\text{O}_2/\text{Ar}$ gas phase; (3) carbon support of the Pt catalyst; and (4) Pt catalyst.

ally H_2/Ar was 50/50 vol%) was used as reference and counter gas at the other electrode. The gas flow rate at both electrodes was 120 mL/min . The temperature was $135\text{--}137^\circ\text{C}$.

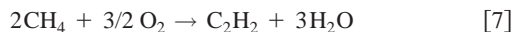
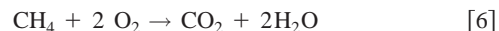
Commercial gases CH_4 (99.95%), H_2 (99.9%, $\leq 10 \text{ ppm O}_2$, $\leq 15 \text{ ppm H}_2\text{O}$), O_2 (99.8% $\text{O}_2 + 0.2\% \text{ N}_2$ and Ar), and N_2 ($< 40 \text{ ppm O}_2 + \text{H}_2\text{O}$) were used.

An EG&G Instruments (Princeton Applied Research) 283 potentiostat/galvanostat and homemade potentiostat controlled by 352 SoftCorrTM III software and homemade software were used for electrode polarization and steady-state voltammetric measurements. The initial (at zero polarization) catalytic activity was measured after the yield of gas products had stabilized, *i.e.*, under steady-state conditions. For each value of polarization the measurements were performed after stabilization of the yield of gas products (1-1.5 h). Turnover frequency *vs.* polarization curves were reproduced for each studied product and temperature.

Results and Discussion

The results of the investigation of the electrochemical promotion of the catalytic methane oxidation are given in Fig. 3-7.

It has been found that C_2H_2 , CO_2 , and water are the main oxidation products. Without polarization the yield of C_2H_2 was 0.9% (Fig. 3) and the yield of CO_2 was 7.3% (Fig. 4). This means that C_2 open-circuit selectivity was approximately 11%. The open-circuit voltage (OCV) was around 0.6 V. The reactions are assumed to be



The oxidative coupling of methane (OCM) is a relatively widely studied catalytic reaction.^{18,19} However, in most OCM cases the major products are C_2H_6 , C_2H_4 , and CO_2 . It is also known that direct pyrolysis of methane to acetylene takes place at temperatures close to 2000°C with 30% methane conversion.²⁰ An alternative route for direct conversion of methane into acetylene can be made through the activation in a two-stage burner by a hydrogen/oxygen flame.²¹ Thus, in an alumina reactor, acetylene can be produced from the preheated (700°C) $\text{CH}_4/\text{O}_2/\text{H}_2$ gaseous mixture with a yield of $\sim 17\%$. It was found in a recent investigation²² that the plasma catalytic conversion of methane produces acetylene with high selectivity and yield under atmospheric pressure in the temperature range $70\text{--}500^\circ\text{C}$. The highest yield of C_2 hydrocarbons (about 20% C_2 yield and more than 85% selectivity toward acetylene) was obtained in the hydrogen-containing plasma. Lower temperature favors the formation of acetylene, while no acetylene is formed at higher temperatures (more than 300°C).

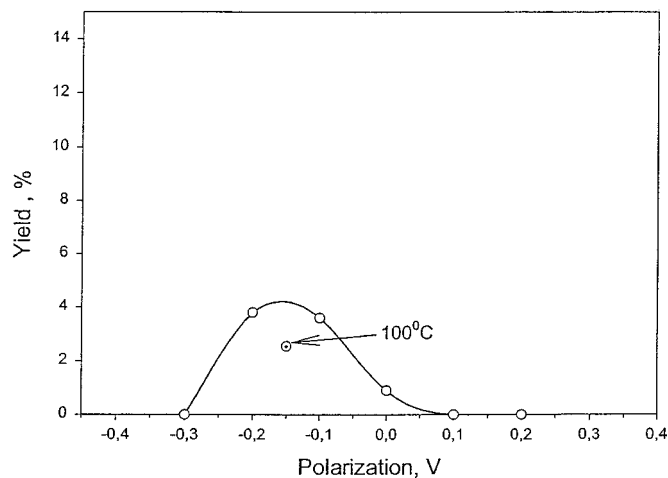


Figure 3. Dependence of C_2H_2 yield on polarization obtained in (CH_4, O_2, Ar) , Pt/PBI- H_3PO_4 /Pt, (H_2, Ar) fuel cell at $135^\circ C$.

It can be concluded that the presence of hydrogen in the reactant mixture at low temperature increases the C_2H_2 yield of OCM. Therefore, the most probable explanation of the unusual $CH_4 \rightarrow C_2H_2$ conversion in the present study is based on proton conductivity of the Pt catalyst support. It can be assumed that there is an $H^+ \rightleftharpoons H_{ads}$ equilibrium at the three-phase boundary of the catalyst, and this equilibrium is a source of the H_{ads} active centers for the $CH_4 \rightarrow C_2H_2$ conversion.

Dependence of the C_2H_2 yield on polarization is given in Fig. 3. It can be seen that the $CH_4 \rightarrow C_2H_2$ catalytic reaction can be electrochemically promoted at negative polarization and exhibits a clear “volcano-type” promotion behavior.⁴ This means that there is a maximum promotion effect at a polarization of -0.15 V, or 0.45 V catalyst potential vs. the hydrogen electrode (3.8% C_2H_2 yield). The catalytic rate enhancement ratio, $r(C_2)/r_0(C_2)$ (Ref. 4), at this maximum is 4.2. It can also be seen from Fig. 3 that there is no C_2H_2 production at polarizations ≥ 0.1 V and ≤ -0.3 V. The yield of C_2H_2 decreases with decreasing temperature. The value of λ calculated for the case of maximum promotion effect was 7.9×10^3 , i.e., $\lambda \gg 1$. This means that the electrochemical promotion for the CH_4 - C_2H_2 route has an EPP nature.

Dependence of CO_2 yield on polarization is given in Fig. 4. This

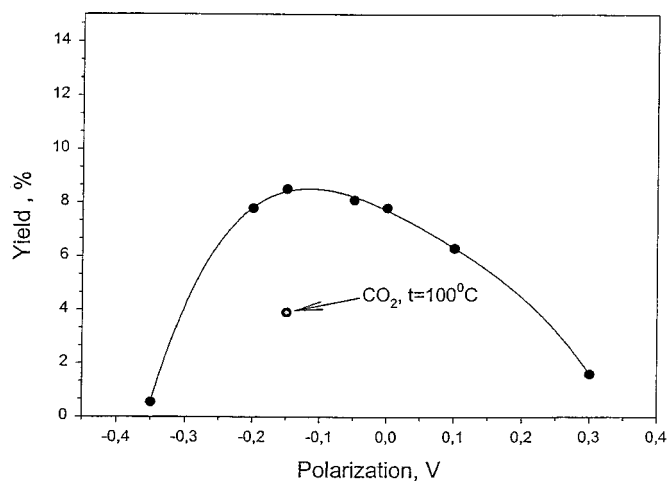


Figure 4. Dependence of CO_2 yield on polarization obtained in (CH_4, O_2, Ar) , Pt/PBI- H_3PO_4 /Pt, (H_2, Ar) fuel cell at $135^\circ C$.

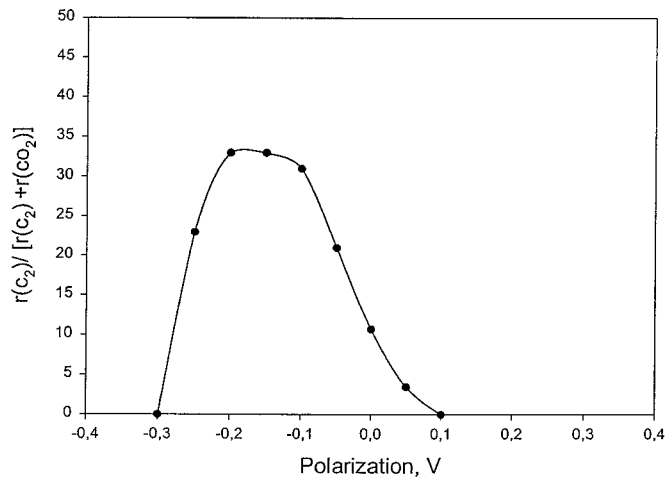


Figure 5. Dependence of $r(C_2)/[r(C_2) + r(CO_2)]$ ratio on polarization obtained in (CH_4, O_2, Ar) , Pt/PBI- H_3PO_4 /Pt, (H_2, Ar) fuel cell at $135^\circ C$.

dependence also shows a volcano-type behavior with maximum yield of 8.3% at -0.15 V, i.e., at the same potential as for C_2H_2 production. The catalytic rate enhancement ratio for CO_2 production, $r(CO_2)/r_0(CO_2)$, at this maximum is 1.1, which means that this catalytic reaction is only slightly affected by the electrochemical polarization. This also means that the polarization affects the C_2 selectivity more than the CO_2 selectivity. Figure 5 shows the “volcanic” nature of the dependence of the ratio between $r(C_2)$ and $r(C_2) + r(CO_2)$ with a maximum at -0.2 V. Similar to the C_2H_2 path, CO_2 production is inhibited at very positive and very negative polarizations and decreases with a temperature decrease from 135 to $100^\circ C$ (Fig. 4). The value of λ calculated for the case of maximum promotion effect was 2.7×10^3 , i.e., $\lambda \gg 1$, meaning that the electrochemical promotion for the CH_4 - CO_2 route (as in the case of the CH_4 - C_2H_2 route) has an EPP nature.

The same potential of maximum promotion for both C_2H_2 and CO_2 paths shows that polarization probably has the same effect on the rate-determining stages of these paths. It would therefore be useful to review the published data on electrochemical conversion of methane. Since the 1960s there have been many publications on the electrocatalytic oxidation of methane in connection with the development of phosphoric acid fuel cells.¹⁸ This reaction was mainly studied on platinum and platinum group metal catalysts. In most of the studied cases there was a complete electrocatalytic oxidation of methane to CO_2 .¹⁸ It has been shown that the rate-determining step of this reaction is dissociative chemisorption of methane at the catalyst-electrolyte interface. Moreover, it has been found that on a Pt-black catalyst in 4.3 M $HClO_3$ at $65^\circ C$ the maximum amount of methane is adsorbed at 0.3 V vs. the standard hydrogen electrode (SHE).²³

The only studied electrochemical system that provided the oxidative coupling of methane was a high-temperature fuel cell with zirconia-based solid electrolytes.¹⁸ The effect of polarization on the oxidative methane coupling in the $CH_4, O_2 / Ag / YSZ / Pt / O_2$ fuel cell at $800^\circ C$ was studied by Seimanides and Stoukides.²⁴ They have shown that “oxygen pumping” at positive polarization promotes methane conversion to CO_2 and that negative polarization slightly promotes the OCM to C_2H_4 and C_2H_6 , therefore promoting C_2 selectivity.

In order to explain similarities and differences in the electrochemical promotion of methane conversion in $CH_4, O_2 / Pt / PBI-H_3PO_4 / Pt, H_2$, and $CH_4, O_2 / Ag / YSZ / Pt / O_2$ fuel cell systems we should compare the open-circuit potentials of the Pt and Ag catalysts in these systems. In the present study for the Pt catalyst it was 0.6 V vs. the hydrogen electrode and in Ref. 24 it was

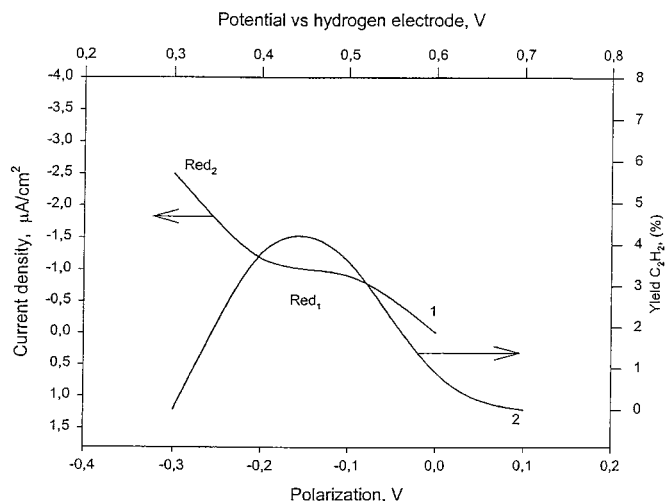


Figure 6. Steady-state voltammetric curve (1) obtained at Pt catalyst and dependence of C₂H₂ yield on polarization (2) obtained in (CH₄, O₂, Ar), Pt/PBI-H₃PO₄/Pt, (H₂, Ar) fuel cell at 135°C.

-0.45 V vs. the oxygen electrode, *i.e.*, approximately 0.6-0.8 V against the hydrogen electrode. Considering the value of the potential of the maximum CH₄ adsorption, *i.e.*, 0.3 V vs. SHE,²³ the catalyst potentials in both fuel cells should be moved in the negative direction in order to achieve this maximum. The data obtained in this work on OCM and C₂ selectivity is in agreement with this assumption. Moreover, the much stronger promotion effect in the CH₄, O₂/Pt/PBI-H₃PO₄/Pt, H₂ system can be explained on the basis of the difference in the current carriers in this system and the CH₄, O₂/Ag/YSZ/Pt/O₂ fuel cell, protons and oxide ions, respectively.

A steady-state voltammetric curve obtained at the Pt catalyst (curve 1) and C₂H₂ yield vs. polarization dependence (curve 2) are given in Fig. 6. It can be seen that the polarization range of the C₂H₂ production (0.3-0.6 V vs. the hydrogen electrode) is the potential range of the reduction wave Red₁. This wave has a limiting current plateau at around 1 $\mu\text{A}/\text{cm}^2$ at the potentials of the maximum promotion effect. It has been found²⁵ that during the electrochemical reduction of H⁺ in 0.1 M H₃PO₄ at 25°C at a Pt rotating disk electrode, covered by PBI thin film, so-called strongly adsorbed hydrogen is produced in the region of 0.2-0.4 V vs. the reversible hydrogen electrode. It is natural to assume that wave Red₁ is caused by the electrochemical production of strongly adsorbed hydrogen. The combination of the maximum CH₄ adsorption and electrochemical production of the Pt-H active sites (in the same potential region), which stabilize C₂ products, can provide a strong promotion effect at the CH₄, O₂/Pt/PBI/H₃PO₄ boundary. In the case of the CH₄, O₂/Ag/YSZ/Pt/O₂ fuel cell, this effect is much weaker because of the absence of the proton current carriers in the solid electrolyte.

It was found that electrochemical conversion of CH₄ into CO₂ proceeds through the formation of the so-called O-type products of the dissociative adsorption of methane, *i.e.*, ≡COH, -CHO.¹⁸ This means that in this case parameters of the oxygen adsorption should play an important role in the kinetics of the methane conversion. Positive polarization of the Ag catalyst²⁴ probably provides the optimum potential for the oxygen adsorption at the catalyst-gas interface.

The experimental data is in good agreement with our assumption about the EPP nature of the electrochemical promotion in the CH₄, O₂/Pt/PBI-H₃PO₄/Pt, H₂ system. It is also in agreement with

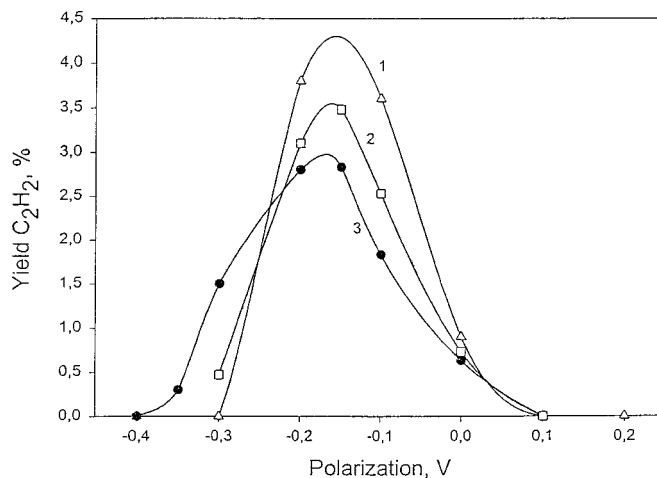


Figure 7. Dependence of electrochemical promotion on the partial pressure of hydrogen at the counter electrode obtained in (CH₄, O₂, Ar), Pt/PBI-H₃PO₄/Pt, (H₂, Ar) fuel cell at 135°C (1) 100 vol %; (2) 50 vol %; (3) 0 vol % H₂).

the assumption that through the reversible²⁵ electrochemical production of Pt-H sites we can create a new OCM route for catalytic CH₄ oxidation.

The experimental data given in Fig. 7 are also in agreement with the assumption about the EPP nature of the promotion of C₂H₂ production. Reversibility of the H⁺ ⇌ H_{ads} electrochemical reaction means that this reaction proceeds under diffusion control. Therefore, by changing the H⁺ concentration gradient across the PBI electrolyte we should affect the H⁺ diffusion rate and in this way the electrochemical H⁺ reduction rate and the rate of OCM. As can be seen from Fig. 7, the zero hydrogen pressure at the counter electrode, which should provide a lower H⁺ concentration gradient across the electrolyte, also provides a weaker promotion effect on the methane conversion into C₂H₂.

Conclusions

The possibility of creation of a new OCM route for catalytic CH₄ oxidation by the electrochemical production of Pt-H centers at the Pt-PBI(H₃PO₄)-gas boundary has been demonstrated.

It was shown that Pt catalyst activity and selectivity toward the CH₄ → C₂H₂ reaction can be electrochemically promoted with the maximum promotion effect at -0.15 V.

The nature of electrochemical promotion was discussed. The conclusion was made that there are three different types of electrochemical promotion depending on values of faradaic current (*I*) and λ .

1. The electrochemical promotion effect is the charge-induced (electrostatic) change of the strength of chemisorptive bonds at the catalyst-reactant interface (CI-effect). The CI-effect is characterized by $I = 0$. Charging of the electric double layer at the catalyst-support interface is the only result of the polarization of the catalyst. $\Delta\Psi \neq 0$ and $\Delta\Phi \neq \eta$.

2. The electrochemical promotion effect is induced by the electrochemical production of the catalyst promoters (EPP-effect). $I \neq 0$; $\lambda \gg 1$; $\Delta\Psi \approx 0$, and $\Delta\Phi \approx \eta$.

3. The case of electrocatalysis is when reactant (or reactants) of a catalytic reaction is oxidized or reduced electrochemically. $I \neq 0$, $\lambda \approx 1$.

The EPP nature of the electrochemical promotion of the CH₄ → C₂H₂ catalytic reaction was demonstrated ($\lambda \gg 1$).

Acknowledgments

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