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Determination of Oxygen Transport Properties from Flux and Driving Force Measurements

Bjarke Thomas Dalslet,^z Martin Søggaard, and Peter Vang Hendriksen

Risø National Laboratory, Roskilde 4000, Denmark

We demonstrate that an electrolyte probe can be used to measure the difference in oxygen chemical potential across the surface, when an oxygen flux is forced through an oxygen permeable membrane disk. The oxygen flux as well as the total oxygen chemical potential difference is carefully controlled by an oxygen pump. The developed method is tested on a $(\text{La}_{0.6}\text{Sr}_{0.4})_{0.99}\text{Co}_{0.2}\text{Fe}_{0.8}\text{O}_{3-\delta}$ membrane. An $\text{La}_{0.75}\text{Sr}_{0.25}\text{MnO}_3/\text{Y}_{0.16}\text{Zr}_{0.84}\text{O}_{1.92}/\text{La}_{0.75}\text{Sr}_{0.25}\text{MnO}_3$ oxygen pump was attached to one side of the membrane. A conical $\text{Ce}_{0.9}\text{Gd}_{0.1}\text{O}_{1.95}$ (CG10) electrolyte probe was pressed against the other side of the membrane. The voltage difference between the base and the tip of the CG10 probe was recorded with an applied oxygen flux through the membrane. This voltage was used to extract precise values of the surface exchange rate constant, k_{O} . Using these values of k_{O} , the vacancy diffusion factor, D_{v}^0 , could be extracted from data of the flux and the oxygen chemical potential difference across the membrane measured with the oxygen pump. Furthermore, upon a gas change, the transient voltage signals of the oxygen pump and the probe could be fitted to give values of D_{v}^0 and k_{O} .

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Mixed ionic and electronic conductors (MIECs) are interesting materials for cathodes in solid oxide fuel cells.^{1,2} Other uses of MIECs exist in the field of controlled oxidation, oxygen production, or reactors for partial oxidation of methane to synthesis gas.^{3,4} Current research focuses on identifying materials combining good mechanical, catalytic, and oxygen permeation properties. To rationalize this process, accurate measurement methods and models are needed to evaluate the catalytic properties and the ionic and electronic transport parameters. Numerous studies using methods like conductivity relaxation⁵⁻⁸ and isotopic exchange^{9,10} characterize both the ionic exchange process on the surface of the MIEC and the bulk transport of ions, based on the fitting of a single transient response. Other studies have investigated the oxygen permeability of membranes positioned between gas flows of different oxygen chemical potential.¹¹⁻¹³ These conventional methods lack precision, as the surface exchange process and the bulk transport process can be difficult or impossible to distinguish from each other. Several studies^{1,8,14,15} show, for instance, the difficulty in measuring the surface exchange rate of perovskite MIEC samples thicker than 0.1 mm at high temperature and oxygen partial pressure, where the surface exchange reaction is so fast that transport is completely dominated by the bulk. A possible solution is to investigate samples of different geometry (thick membranes to determine the bulk transport, thin films for surface sensitivity). This is, however, a cumbersome and not always satisfactory approach as the relaxations of thin films are very fast at high temperatures. Furthermore, the surface properties of thin films are not necessarily the same as those of bulk samples as they depend heavily on the substrate and preparation method.¹⁶

The objective of this study is to demonstrate a setup with improved measurement precision, combining a surface sensitive electrolyte probe and an oxygen pump. First, the required equipment is detailed; second, the basic theory of electrolyte sensors and mixed conductors is treated; third, the models needed to interpret the measurements are described; and finally, details of the data treatment and minor corrections will be discussed. The transport properties of the well characterized perovskite MIEC $(\text{La}_{0.6}\text{Sr}_{0.4})_{0.99}\text{Co}_{0.2}\text{Fe}_{0.8}\text{O}_{3-\delta}$ (LSCF) measured by the method are then presented, followed by a discussion of possible misinterpretations and inaccuracies of the models.

Oxygen pumps, i.e., electrically controlled electrochemical cells, have frequently been used to measure oxygen fluxes with high precision in titration and permeation studies.¹⁷⁻²⁰ In this study, the oxygen pump is used to measure the oxygen flux passing through a membrane; as the flux is directly proportional to the electronic cur-

rent applied between the pumping electrodes, the oxygen flux can be measured with very high accuracy. Furthermore, a set of reference electrodes on the pump can be used to measure the oxygen chemical potential, μ_{O_2} , in situ. Electrolyte probes have been used to measure properties of oxygen ion conducting materials by, e.g., Fouletier et al.²¹ and Wiemhöfer et al.²² Their potential for measuring the surface exchange kinetics was proposed by Bouwmeester.¹ The electrolyte probe in this study measures the difference in chemical potential of oxygen across the membrane surface as a voltage. The combination of a local probe and a precise flux measurement allows the precise determination of both the bulk transport and surface exchange parameters while only applying a small steady state oxygen chemical potential difference across the membrane ensuring an almost homogenous oxide ion distribution. Furthermore, relaxation measurements are made by application of a stepwise gas composition change on one side of the membrane. This causes a relaxation of the oxide ion distribution in the membrane. The resulting transient response in the driving force gradients provides another experimental route to determine the bulk transport and surface exchange parameters.

Experimental

Sample preparation.—LSCF powder was prepared by the glycine pyrolysis process.²³ The powder was calcined at 900°C for 12 h, and ballmilled using 5×5 mm cylindrical ZrO_2 balls in a 500 mL polyethylene container with ethanol for 24 h at 200 rpm. A flat cylindrical membrane was shaped from this powder using a uniaxial pressure of 70 MPa. This membrane was isostatically compressed in an evacuated latex container suspended in water at a pressure up to 325 MPa. The membrane was sintered at 1300°C for 12 h. The sintered membrane was polished with SiC polishing sheets and diamond suspensions (particle size down to 1 μm). The sintered and polished membrane had a diameter of 20.3 ± 0.1 mm, a thickness of 2.253 ± 0.005 mm and a density of 5965 kg m^{-3} (95% of bulk). X-ray diffraction revealed a single phase hexagonal perovskite with cell parameters $a = 5.487 \text{ \AA}$ and $c = 13.48 \text{ \AA}$.

Setup and measurement procedure.—The experimental setup is shown in Fig. 1. The MIEC membrane is placed between two 300 μm thick gold sealing rings. The inner diameter of the gold rings is 14 mm and defines the perimeter of the active surface of the membrane. The closed compartment is defined by the membrane, the lower gold sealing ring and an $\text{Y}_{0.16}\text{Zr}_{0.84}\text{O}_{1.92}$ (YSZ) oxygen pump. Two pairs of electrodes are painted on the oxygen pump. Each pair has one electrode in the closed compartment, and one electrode in the reference compartment, which is flushed with an air flow ($>2.10^{-6} \text{ m}^3/\text{s}$ or 120 mL/min at ambient conditions). The

^z E-mail: bjarkehingrumme@hotmail.com

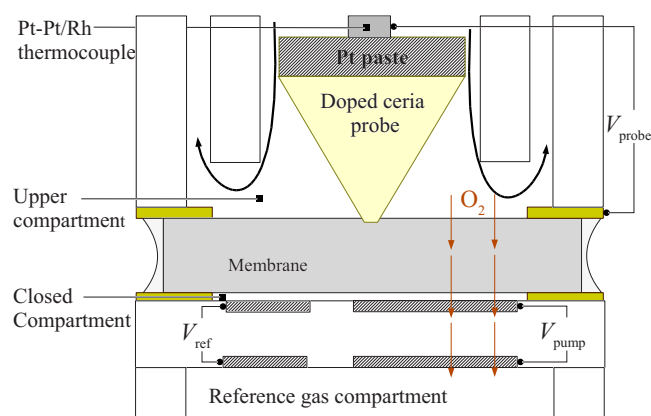
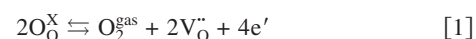


Figure 1. (Color online) The measurement setup.

pumping electrode pair is made of a mixture of porous 50% YSZ + 50% $\text{La}_{0.75}\text{Sr}_{0.25}\text{MnO}_3$ with porous Pt as current collector and is used to pump oxygen in and out of the closed compartment. The *reference* electrode pair is made of porous Pt (sintered Ferro Pt paste), and is used to measure the difference in oxygen chemical potential between the closed compartment and the reference compartment. The circumference of the membrane is sealed using a glass seal. Above the membrane an alumina cylinder defines an upper compartment. Two concentric alumina pipes terminate in the upper compartment. The inner pipe feeds gas (the “flow gas”) to the upper compartment while the outer pipe transports the exhaust gas to a Pt|YSZ|Pt p_{O_2} sensor. The gas flow through the upper compartment is $1.67 \times 10^{-6} \text{ m}^3/\text{s}$ (100 mL/min measured at ambient conditions) of air diluted with N_2 . Inside the inner pipe a spring loaded thermocouple applies an electrical contact and a downward mechanical force to a pointed electrolyte probe in contact with the membrane surface. The pointed probe was made from $\text{Ce}_{0.9}\text{Gd}_{0.1}\text{O}_{1.95}$ (CG10). CG10 was selected over YSZ to avoid SrZrO_3 formation at the probe tip. A cylinder was shaped using an uniaxial pressure of 70 MPa. It was then isostatically compressed in an evacuated latex container suspended in water at a pressure up to 325 MPa. After sintering at 1323 K the CG10 cylinder was mechanically sharpened and resintered at 1873 K. All electrical contacts were made using platinum and gold wires. A Solartron 1250/1286 setup was used as a programmable current source with a Keithley 2700 multimeter/data collector.

Defect chemistry and transport properties of LSCF.— In this section, the defect chemistry and transport properties of the test material, LSCF, are described. Two models for interpretation of steady state and relaxation measurements, respectively, are presented.

Defect chemistry.— The release and incorporation of oxygen in LSCF can be written in the Kröger-Vink notation



At equilibrium the chemical potential of oxygen, μ_{O_2} is the same in the gas and solid phase. This allows us to label the state of the LSCF MIEC with the state of the gas with which it is in equilibrium, and vice versa. The chemical potential of oxygen, $\mu_{\text{O}_2}^{\text{MIEC}}$, in a MIEC in equilibrium with a gas with a given chemical potential of oxygen $\mu_{\text{O}_2}^{\text{eq.gas}}$ can then be defined as

$$\mu_{\text{O}_2}^{\text{MIEC}} = \mu_{\text{O}_2}^{\text{gas}} = \mu_{\text{O}_2}^{\ominus} + RT \ln(p_{\text{O}_2}^{\text{eq.gas}}/p_{\text{O}_2}^{\ominus}) \quad [2]$$

where R is the gas constant and T is the temperature; $p_{\text{O}_2}^{\text{eq.gas}}$ and $p_{\text{O}_2}^{\ominus}$ are the oxygen partial pressures of the equilibrium and standard gas, respectively. The oxygen chemical potential of the selected standard gas $\mu_{\text{O}_2}^{\ominus}$ can be calculated using the IUPAC standards,²⁴ see Appendix A. Likewise, the equivalent oxide concentration in the gas, $C_{\text{O}}^{\text{gas}}$, is defined as the oxide concentration $C_{\text{O}}^{\text{MIEC}}$ in the MIEC with which it is in equilibrium. A difference in $\mu_{\text{O}_2}^{\text{MIEC}}$, $\Delta\mu_{\text{O}_2}^{\text{MIEC}}$ can then (for small differences) be written in terms of a difference in $C_{\text{O}}^{\text{MIEC}}$, $\Delta C_{\text{O}}^{\text{MIEC}}$

$$\Delta\mu_{\text{O}_2}^{\text{MIEC}} = RT \Delta \ln p_{\text{O}_2}^{\text{eq.gas}} \approx RT \frac{\Delta C_{\text{O}}^{\text{MIEC}}}{C_{\text{O}}^{\text{MIEC}}} \frac{\partial \ln p_{\text{O}_2}^{\text{eq.gas}}}{\partial \ln C_{\text{O}}^{\text{MIEC}}} = 2RT \frac{\Delta C_{\text{O}}^{\text{MIEC}}}{C_{\text{O}}^{\text{MIEC}}} \gamma \quad [3]$$

where $\gamma = \partial \ln p_{\text{O}_2}^{\text{eq.gas}} / 2 \partial \ln C_{\text{O}}^{\text{MIEC}}$ is the thermodynamical factor.

The equilibrium relation between μ_{O_2} and $C_{\text{O}}^{\text{MIEC}}$ for a given $p_{\text{O}_2}^{\text{eq.gas}}$ has been determined using thermogravimetry and coulometric titration. Thermogravimetry is a widely used method. coulometric titration measures $\mu_{\text{O}_2}^{\text{MIEC}}$ as function of oxygen pumped, via an oxygen pump, into a sealed chamber with a powder sample. A detailed description of this method is beyond the scope of this paper, and we refer to the paper by Zachau-Christiansen et al.¹⁷ The values for the thermodynamic factor, γ , and oxygen nonstoichiometry parameter, δ , are reproduced in Fig. 2. Further details will be made available in a future publication.

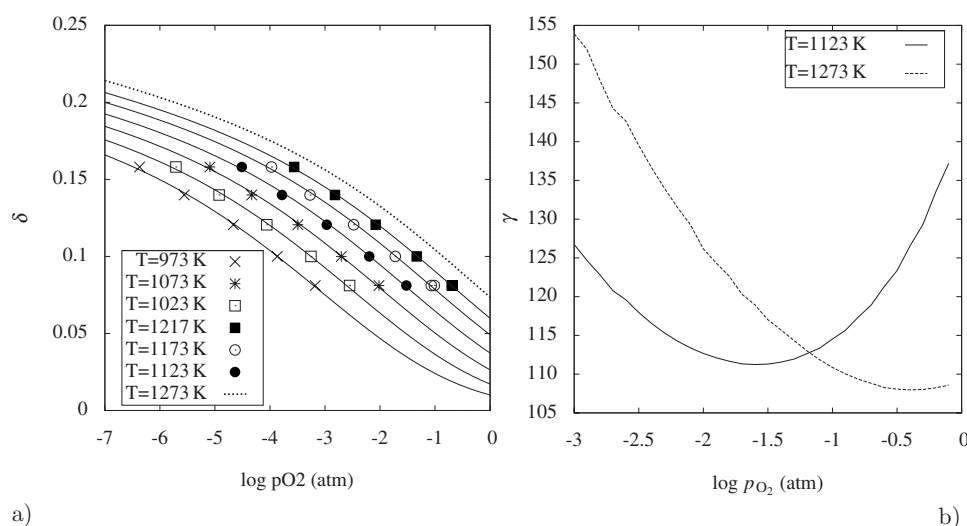


Figure 2. (a) δ as function of p_{O_2} . Points are measured values and the lines are from a model fit. (b) The thermodynamic factor, γ , as function of p_{O_2} derived from the model lines in (a).

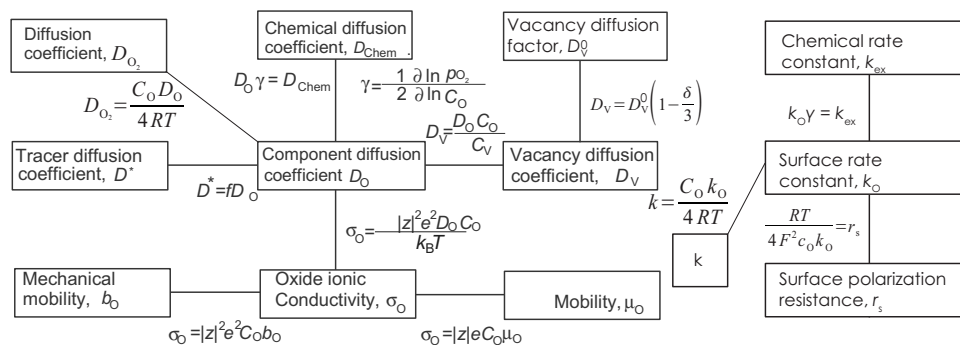


Figure 3. Relations between transport parameters in LSCF: f is the correlation factor (0.69 for perovskites as LSCF), γ is the thermodynamic enhancement factor, z is the charge number of the ionic species, i.e., -2 for O^{2-} , k_B is the Boltzmann constant, e is the electronic charge, and δ is the number of vacancies per formula unit. C_O is the concentration of oxide ions. C_V is the concentration of oxide vacancies. R is the gas constant and T is the temperature.

Transport properties.—In this study different parameters will be used to describe the oxygen transport in the bulk and across the surface. These parameters are described by Maier²⁵ and in the textbooks by Tilley²⁶ or Gellings and Bouwmeester.¹ In LSCF local electroneutrality causes the transport of electrons (e^-) and oxygen ions (O^{2-}) to be coupled. In LSCF, the ionic conductivity is controlling the transport of both electrons and ions as the electronic conductivity is much larger than the ionic conductivity. In the absence of external electrical fields, we are free to define the diffusing species in the material as “oxygen,” O_2 , consisting of two oxide ions O^{2-} and four electron holes. From Gellings and Bouwmeester¹

$$j_{O_2} = -\frac{1}{4^2 F^2} \frac{\sigma_e \sigma_O}{\sigma_e + \sigma_O} \nabla \mu_{O_2}^{MIEC} \approx -D_{O_2} \nabla \mu_{O_2}^{MIEC}, \quad D_{O_2} = \frac{\sigma_O}{4^2 F^2} \quad [4]$$

where F is the Faraday constant, j_{O_2} is the oxygen flux, σ_e is the electronic conductivity, and σ_O is the oxide ion conductivity. D_{O_2} is the diffusion coefficient and is defined under the assumption (true for LSCF) that $\sigma_O \ll \sigma_e$. Combining Eq. 3 with Eq. 4 we get Ficks first law of diffusion

$$j_{O_2} = -\frac{D_{chem}}{2} \nabla C_O^{MIEC}, \quad D_{chem} = \frac{4\gamma RT}{C_O^{MIEC}} D_{O_2} \quad [5]$$

where D_{chem} is the chemical diffusion coefficient of oxide ions. From D_{chem} we obtain the oxygen atom self-diffusion coefficient D_O as $D_{chem} \gamma^{-1}$. As most of the oxide sites in the structure are occupied, the diffusion of oxide ions will be controlled by the number of oxide vacancy sites available. For small vacancy concentrations the transport equation becomes

$$j_{O_2} = \frac{\gamma_V}{2} D_V^0 \left(1 - \frac{\delta}{3}\right) \nabla C_V, \quad D_V^0 = \frac{4RT}{C_V \left(1 - \frac{\delta}{3}\right)} D_{O_2} \quad [6]$$

$$\gamma_V = \gamma \frac{C_V}{C_O^{MIEC}}$$

where D_V^0 is the vacancy diffusion factor, C_V is the vacancy concentration, and γ_V is the thermodynamic factor associated with the oxygen vacancies. The term $(1 - \delta/3)$ was proposed by Lankhorst¹⁴ because a vacancy cannot diffuse to a vacant oxide site. We expect that D_V^0 is independent of C_V apart from second order effects (such as C_V influencing the crystal structure or immobilization of vacancies due to ordering). When interpreting our measurements D_V^0 will therefore be used as a fitting parameter.

When the surface is brought out of equilibrium with the surrounding atmosphere, a step in μ_{O_2} , $\Delta\mu_{O_2}$, exists at the surface. Assuming a linear response of the flux to the step in μ_{O_2} we can define a surface exchange coefficient, k , and an equivalent surface resistance r_s

$$j_{O_2} = -k \Delta\mu_{O_2} = -\frac{1}{4^2 F^2 r_s} \Delta\mu_{O_2} \quad [7]$$

The relations between the different transport parameters are summarized in Fig. 3.

Steady state model.—To fit data recorded during steady state permeation measurements a one-dimensional numerical model based on the Wagner equation²⁷ combined with oxygen exchange at the surface is used. The Wagner equation is obtained from integration of Eq. 4

$$j_{O_2} = -\frac{1}{4^2 F^2 L} \int_{\mu'_{O_2}}^{\mu''_{O_2}} \frac{\sigma_e \sigma_O}{\sigma_e + \sigma_O} d\mu_{O_2} \quad [8]$$

where μ'_{O_2} and μ''_{O_2} are the chemical potentials of oxygen at each interface (just inside the bulk, see Fig. 4).

The μ_{O_2} profile in Fig. 5 has been calculated by a combination of Eq. 7 at the surfaces and Eq. 8 in the bulk. The calculation was done with $i_{pump} = -100$ A/m² and all material properties as determined for LSCF at $T = 1248$ K and $p_{O_2} = 4$ kPa. The vertical parts are the discrete jumps in μ_{O_2} at the surfaces, and the apparently linear part is in the bulk. The linearity is achieved as very small gradients in μ_{O_2} are applied as the oxygen pump allows very precise control of the oxygen flux. The need for a large signal/noise ratio requires a

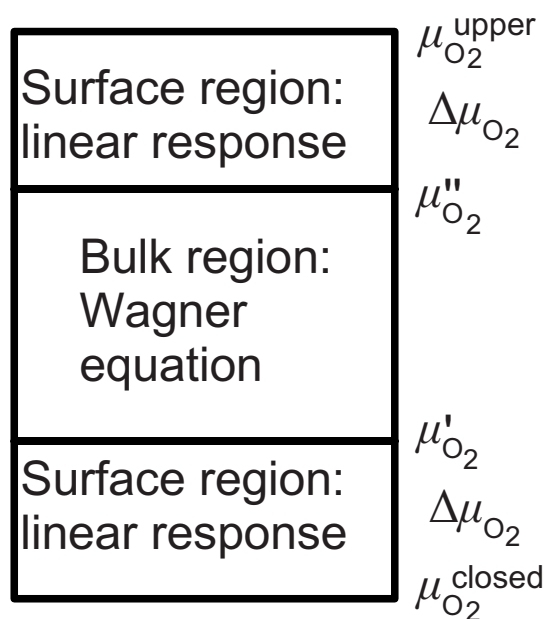


Figure 4. Sketch of the model used for analysis of steady state permeation measurements.

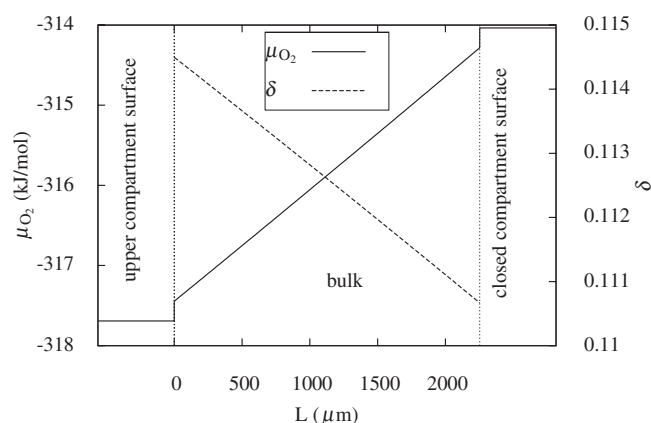


Figure 5. An example of the chemical potential profile, and the δ value of the membrane in the experiment shown in Fig. 9 calculated using the steady state model.

current density of more than 10 A/m², but Fig. 5 shows that even at $i_{\text{pump}} = -100$ A/m² ($I_{\text{pump}} = 15$ mA) only a very small variation in δ is imposed, and the transport parameters are therefore almost constant throughout the membrane.

The gold seals act as apertures for the surface exchange (see Fig. 1) but the membrane cross section is wider than these apertures, resulting in an overestimation of D_v^0 . D_v^0 can be corrected by multiplying with a membrane geometry dependent factor f_c (for the membrane of this study it has a value of 0.93).²⁸

Relaxation model.—When changing the p_{O_2} of the flow gas in the upper compartment the membrane will adjust its δ value to the value of μ_{O_2} . This relaxation can be fitted with an appropriate model. When the membrane is not in steady state, the time dependence of $\mu_{\text{O}_2}^{\text{MIEC}}$ will be governed by conservation of mass

$$\frac{\partial C_{\text{O}}^{\text{MIEC}}}{\partial t} = \nabla \cdot \mathbf{j}_{\text{O}} \quad [9]$$

where t is the time. Combining with Eq. 3 we get

$$\frac{\partial \mu_{\text{O}_2}^{\text{MIEC}}}{\partial t} = 2 \frac{RT\gamma}{C_{\text{O}}^{\text{MIEC}}} \nabla \cdot \mathbf{j}_{\text{O}_2} \quad [10]$$

Equations 4, 7, and 10 are solved using an iterative algorithm in a cylindrical finite element mesh. The finite element mesh consists of 20 segments in the longitudinal dimension and ten in the radial dimension. The model geometry and boundary conditions are outlined in Fig. 6.

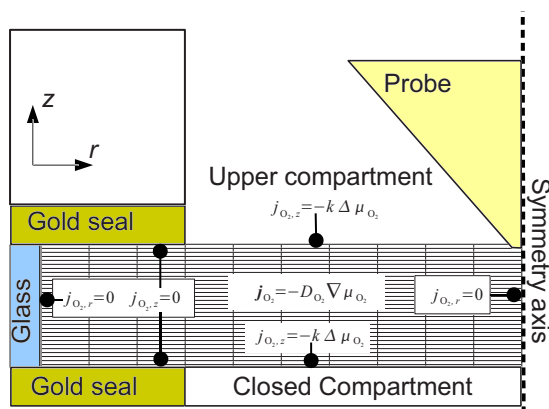


Figure 6. (Color online) Sketch of the model used for analysis of relaxation measurements.

When fitting the relaxation data, D_v^0 is used as the fitting parameter; D is calculated from D_v^0 for each spatial segment of the membrane. D_v^0 is assumed constant for the entire membrane.

Measurements using an oxygen pump and an electrolyte probe.—In this section the basic principle of using electrolytes to measure chemical potential differences, and the operation principle of the oxygen pump, is described. An area specific conductance of the membrane is introduced and a few necessary corrections are treated. Finally, the same principle is used to give an interpretation of the signal from an electrolyte probe along with a few corrections. **Oxygen chemical potential measurement.**—The basic principle behind both the probe and the oxygen pump is using oxide conducting electrolytes to measure differences in the chemical potential of oxygen. An oxide conducting electrolyte placed between two external phases with different oxygen chemical potential will approach an equilibrium between the oxide or oxygen in the two external phases and the oxide in the electrolyte surfaces. The relationship, at equilibrium, between the oxygen activity of the external phases and cell voltage, as measured between electrodes placed on the two interfaces, will be

$$\mu'_{\text{O}_2} - \mu''_{\text{O}_2} = 4FV_{\text{ele}} \quad [11]$$

where μ'_{O_2} and μ''_{O_2} are the oxygen chemical potential in each of the external phases, and V_{ele} is the voltage between the two electrolyte|external phase interfaces.

The oxygen pump.—The YSZ oxygen pump is an electrochemical cell with a YSZ electrolyte. It is assumed that the following reaction takes place at the pumping and reference electrodes



Combining Eq. 2 and 11, the Nernst equation governing the relationship between oxygen activity and cell voltage will be

$$p_{\text{O}_2}^{\text{gas}} = p_{\text{O}_2}^{\text{ref}} \exp\left(\frac{4FV_{\text{ref}}}{RT}\right) \quad [13]$$

$p_{\text{O}_2}^{\text{gas}}$ is the oxygen partial pressure of the gas in the closed compartment, $p_{\text{O}_2}^{\text{ref}}$ is the p_{O_2} of the gas in the reference compartment (air in these measurements) and V_{ref} is the voltage between the reference electrodes. As four electrons pumped in the external circuit will correspond to one oxygen molecule passing between the electrodes of the cup, the relationship between pumping current and oxygen transfer is

$$j_{\text{O}_2} = \frac{I_{\text{pump}}}{4F} \quad [14]$$

where I_{pump} is the pumping current.

The area specific conductance, G .—In the setup of Fig. 1 the ionic current through the membrane in steady state must equal the ionic current through the pump. In this case an area specific conductance, G , of the membrane can be defined as

$$G = \frac{i_{\text{pump}}}{\Delta V_{\text{Nernst}}}, \quad i_{\text{pump}} = I_{\text{pump}}/A_{\text{MIEC}} \quad [15]$$

$$\Delta V_{\text{Nernst}} = V_{\text{closed}} - V_{\text{upper}} = V_{\text{ref}} - V_{\text{ref}}^0 \quad [16]$$

where i_{pump} is the area specific current and A_{MIEC} is the exposed area of the MIEC membrane. ΔV_{Nernst} is the difference in the Nernst voltages between the gas in the upper compartment, V_{upper} , and the gas in the closed compartment, V_{closed} , with the reference compartment gas as a common reference. $V_{\text{closed}} = V_{\text{ref}}$ are measured by the reference electrodes of the oxygen pump, with the air in the reference compartment providing the reference μ_{O_2} . V_{ref}^0 is measured at zero flux, where the flow gas of the upper compartment is in equilibrium with the gas in the closed compartment (i.e., $V_{\text{upper}} = V_{\text{closed}}$). As the Nernst voltage measures the oxygen chemical potential, which is the driving force of both the bulk (Eq. 7) and the

surface oxygen ion transport (Eq. 8), G can be expected to be independent of the flux if σ_o , σ_e and the surface resistance, r_s , are independent of the flux; this is only expected true for very small fluxes. G can in this case be extracted with good accuracy as the slope of a plot of steady state values of i_{pump} as function of ΔV_{Nernst} . If the surface resistance is small G can be converted to an ionic conductivity

$$\sigma_o = \frac{Gf_c}{L} \quad [17]$$

by dividing with the membrane thickness L and multiplying with the geometric factor f_c .

Effect of gas diffusion in the closed compartment.— In the closed compartment V_{ref} is measured using two reference electrodes painted on either side of the pump. As the gas in the compartment is not stirred, an oxygen concentration gradient will be present between the reference electrode and the membrane when an ionic current runs through the membrane. The diffusion coefficient of O_2 in N_2 at 1273 K is $2.6 \times 10^{-4} \text{ m}^2/\text{s}$.²⁹ As the distance between the electrode and the membrane is 0.3 mm, the p_{O_2} difference needed to support a large 130 A/m^2 pumping current would be 5 Pa. This is much smaller than the difference measured with ΔV_{Nernst} which in this case is more than 100 Pa. The p_{O_2} gradient in the compartment is thus very small and G is not affected by gas diffusion in the compartment.

Effect of permeation flux on the flow gas.— The oxygen flux through the membrane will change the p_{O_2} above the membrane. Applying a simple mass balance approach, the change in p_{O_2} will be proportional to I_{pump} and inversely proportional to the volume flow of gas through the upper chamber, Φ_f . The gas flow direction can be approximated as a radial distribution spreading from the membrane center point and thus μ_{O_2} is not homogeneous across the membrane surface. The average Nernst voltage of the gas in the upper compartment, V_{upper} measured vs air is derived in Appendix B as

$$V_{\text{upper}} = \frac{RT}{4F} \left[\ln \left(\frac{p_{O_2}^{\text{in}}}{p_{O_2}^{\text{air}}} \right) + \left(\frac{p_{O_2}^{\text{in}}}{\Delta p_{O_2}} + 1 \right) \left(\ln \left(\frac{\Delta p_{O_2}}{p_{O_2}^{\text{air}}} + 1 \right) - 1 \right) \right] \quad [18]$$

$$\Delta p_{O_2} = \frac{I_{\text{pump}} RT \Phi_f}{4F \Phi_f} \quad (\text{steady state}) \quad [19]$$

$$\Delta p_{O_2} = \frac{J_{O_2} RT \Phi_f}{\Phi_f} \quad (\text{relaxation}) \quad [20]$$

where T_{Φ_f} is the temperature at which the volume flow is measured. $p_{O_2}^{\text{in}}$ is the p_{O_2} of the inlet gas, calculated from V_{ref}^0 . The deviation of V_{upper} from V_{ref}^0 results in an underestimation of G in the steady state measurement. Substituting V_{upper} for V_{ref}^0 in Eq. 15 gives a corrected area specific conductance G_c . In the relaxation measurements the permeation flux will cause a deviation from the usual step change of μ_{O_2} . If this deviation is not taken into account, k will be underestimated.

The electrolyte probe.— The tip of the electrolyte probe is in ionic and electronic contact with the MIEC membrane; the porous platinum electrode on the other side of the probe sits in the flow gas. Assuming that equilibrium exists between the platinum and the gas and between the probe tip and the MIEC membrane, Eq. 11 and 2 states that the probe voltage, V_{probe} , is a direct measure of the difference in μ_{O_2} between the inlet flow gas and MIEC membrane surface

$$\Delta \mu_{O_2} = \mu_{O_2}^{\text{in}} - \mu_{O_2}^{\text{MIEC}} = 4FV_{\text{surface}} \quad [21]$$

where $\mu_{O_2}^{\text{in}}$ and $\mu_{O_2}^{\text{MIEC}}$ are the oxygen chemical potential, of the inlet flow gas, and just inside the MIEC surface, respectively. V_{surface} is

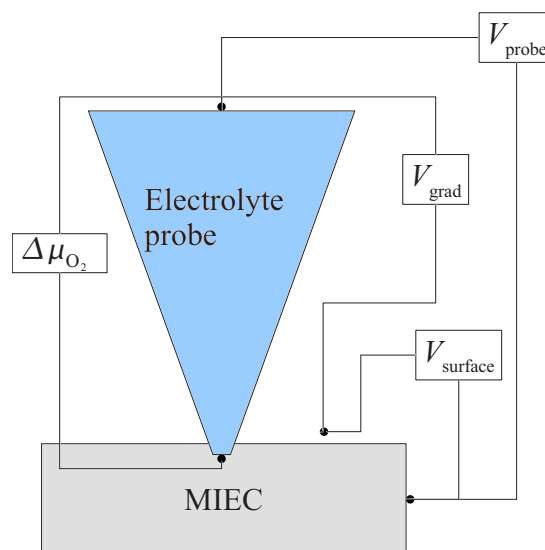


Figure 7. (Color online) A schematic of the origins of the contributions to the probe voltage.

the part of the probe voltage, V_{probe} , attributable to the oxygen exchange over the surface.

Apart from V_{surface} , there are other contributions to V_{probe} when there is a flux of oxygen through the membrane surface (see Fig. 7). They can be written

$$V_{\text{probe}} = V_{\text{surface}} + V_{\text{grad}} + \sum_i V_i \quad [22]$$

where V_{grad} is the contribution from differences in the p_{O_2} of the gas at a point just over the membrane, and the gas at the Pt electrode on the back of the probe. $\sum_i V_i$ is the sum of any other contributions present, such as thermoelectric voltages. $\sum_i V_i$ is assumed to be independent of the ionic current. While $\sum_i V_i$ can be the largest contribution to V_{probe} , it can thus be filtered out by measuring V_{probe} at different currents.

V_{grad} .— V_{grad} is determined by a number of factors including flow geometry, gas diffusion coefficient, T , etc. In the next section it will be shown that V_{probe} is independent of the gas diffusion coefficient. Assuming that the flow geometry is symmetric with respect to a gas flow direction reversal, V_{grad} can be estimated:

The p_{O_2} of the gas close to the probe/MIEC membrane interface, $p_{O_2}^{\text{tip}}$ is different from the p_{O_2} of the inlet gas, due to the oxygen flux through the membrane. Assuming that the flux distribution over the membrane is independent of the flow rate and p_{O_2} of the gas, $p_{O_2}^{\text{tip}}$ can be written

$$p_{O_2}^{\text{tip}} = p_{O_2}^{\text{in}} + \beta(p_{O_2}^{\text{out}} - p_{O_2}^{\text{in}}) \quad [23]$$

where β is a variable depending on the geometry of the setup. If the gas flow direction is reversed such that it enters via the outer pipe and exits via the inner pipe Eq. 23 becomes

$$p_{O_2}^{\text{tip}} = p_{O_2}^{\text{in}} + (1 - \beta)(p_{O_2}^{\text{out}} - p_{O_2}^{\text{in}}) \quad [24]$$

β was extracted using Eq. 23 and 24 by comparing measurements with the regular and inverted gas flow directions. It was found to be 0.09 ± 0.02 ; $p_{O_2}^{\text{tip}}$ is thus close to $p_{O_2}^{\text{in}}$.

V_{grad} can then be calculated as

$$V_{\text{grad}} = \frac{RT}{4F} \ln \left(\frac{p_{O_2}^{\text{in}} + \beta \frac{I_{\text{pump}} RT \Phi_f}{4F \Phi_f}}{p_{O_2}^{\text{in}}} \right) \quad [25]$$

The polarization resistance, r_s , of the surface is finally calculated as

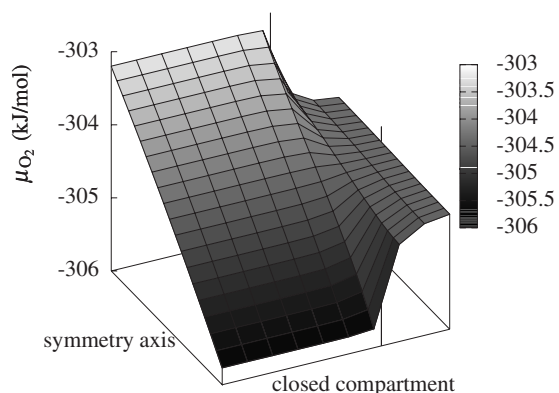


Figure 8. μ_{O_2} in the MIEC membrane in a steady state measurement at $T = 1273$ K in air. $i_{\text{pump}} = -65$ A/m². Two vertical lines mark the edge of the gold seal. The total length along the symmetry axis is 2.253 mm. The total length along the closed compartment edge is 10.15 mm (half the membrane diameter), with the aperture extending 7 mm from the symmetry axis, where the gold seal blocks the flux (see Fig. 6).

$$r_s = \frac{V_{\text{probe}} - V_{\text{mixing}} - \sum_i V_i}{i_{\text{pump}}} \quad [26]$$

In the data analysis we will assume that $\sum_i V_i$ is independent of i_{pump} and equal to V_{probe} when $i_{\text{pump}} = 0$. V_{mixing} is between 0 and 10% of V_{probe} in the measurements of this study. This difference is largest when the p_{O_2} is low and when the temperature is high.

Results and Discussion

In this section the transport properties of the (La_{0.6}Sr_{0.4})_{0.99}Co_{0.2}Fe_{0.8}O_{3- δ} (LSCF) as measured by the described method are presented, including discussions of possible misinterpretations of the measured data.

Steady state measurements.— Steady state experiments are performed by applying various currents to the pump. This creates a gradient in μ_{O_2} . The calculated μ_{O_2} in the membrane in a steady state measurement at $T = 1273$ K in air and with pumping current $i_{\text{pump}} = -65$ A/m² is shown in Fig. 8.

For each value of T and $p_{\text{O}_2}^{\text{in}}$ several different values of i_{pump} were applied; i_{pump} is plotted as function of the reference voltage at $T = 1248$ K in a 4:1 N₂:air gas mixture (i.e., p_{O_2} about 4 kPa), in Fig. 9a). Negative values of i_{pump} correspond to oxygen being pumped from the closed compartment to the reference compartment. Negative values of V_{ref} means that the p_{O_2} in the closed compartment is

lower than the p_{O_2} in the reference compartment. There is clearly a linear relation between i_{pump} and V_{ref} . $G = 10.11 \pm 0.03$ kS/m² is extracted as the slope of the linear fit. Note that the standard deviation is less than 0.1% of the measured value. Most of the G measurements in this study were done with standard deviations less than 1%.

In the same experiment a linear relationship between V_{probe} and i_{pump} is also found and shown in Fig. 9b). Positive values of V_{probe} correspond to μ_{O_2} being lower at the probe tip, than at the platinum electrode at the back of the probe. The deviation of V_{probe} from 0 at $i_{\text{pump}} = 0$ is attributed to thermoelectric voltages in the measuring circuit. $\partial V_{\text{probe}}/\partial i_{\text{pump}}$ was extracted as 6.4 ± 0.2 $\mu\Omega$ m². The standard deviation is about 3%. The standard deviation will be lower at lower T and lower p_{O_2} and higher at higher T and higher p_{O_2} .

The low standard deviations are a consequence of the oxygen pump which allows very precise measurement of the flux and p_{O_2} difference compared to traditional mass-balance experiments.^{11-13,36} Furthermore, the reliance on steady state allows very long sampling times of the measured voltages resulting in high precision on the individual measurement. The nonlinearities caused by the nonconstant σ_{O} of a membrane in a μ_{O_2} gradient are avoided due to the very small μ_{O_2} gradients. This has the added effect that the measured G_c and r_s values are indeed the values at $p_{\text{O}_2} = p_{\text{O}_2}^{\text{in}}$ and not integrated quantities over a large p_{O_2} interval.

p_{O_2} dependence of G_c and r_s .— The $p_{\text{O}_2}^{\text{in}}$ dependence of G_c for different temperatures is plotted in Fig. 10a). An increase in G_c is seen when $p_{\text{O}_2}^{\text{in}}$ is lowered. G_c is proportional to $p_{\text{O}_2}^{-0.114}$ at $T = 1123$ K and $p_{\text{O}_2}^{-0.104}$ at $T = 1273$ K. The $p_{\text{O}_2}^{\text{in}}$ dependence of r_s for different temperatures is plotted in Fig. 10b); r_s is proportional to $p_{\text{O}_2}^{-0.42}$ at $T = 1123$ K and $p_{\text{O}_2}^{-0.65}$ at $T = 1273$ K. The error bars show the precision of the measurements. It is remarkable that the precision of the measurement of r_s is reasonably small, even when measuring at $T = 1273$ K and $p_{\text{O}_2} = 0.21$ atm. The origin of the scatter is discussed in the section on *measurement reproducibility*.

Temperature dependence of G_c and r_s .— G_c is plotted as function of $1/T$ for air ($p_{\text{O}_2}^{\text{in}} = 20.9$ kPa) and an air/N₂ gas mixture ($p_{\text{O}_2}^{\text{in}} = 1.8$ kPa) in Fig. 11a). G_c has an Arrhenius dependence with an activation energy of $E_A = 99$ kJ/mol for both gasses; r_s is plotted as function of $1/T$ for flow gas with $p_{\text{O}_2}^{\text{in}} = 20.9$ and 1.8 kPa in Fig. 11b). The surface reaction $\propto r_s^{-1}$ has an Arrhenius dependence with an activation energy E_A of 172 and 95 kJ/mol for each gas mixture, respectively.

The diffusion coefficient of O₂ in a gas is proportional to $T^{3/2}$.²⁹ Oxygen concentration gradients in the gas phase will thus add terms

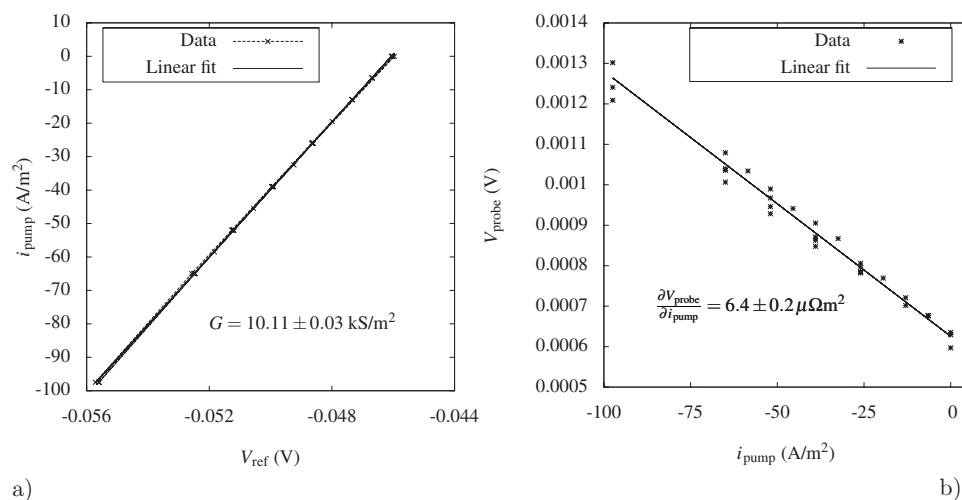


Figure 9. (a) i_{pump} as function of V_{ref} . $T = 1248$ K, gas 20% air + 80% N₂. (b) V_{probe} as function of i_{pump} . $T = 1248$ K, gas 20% air + 80% N₂

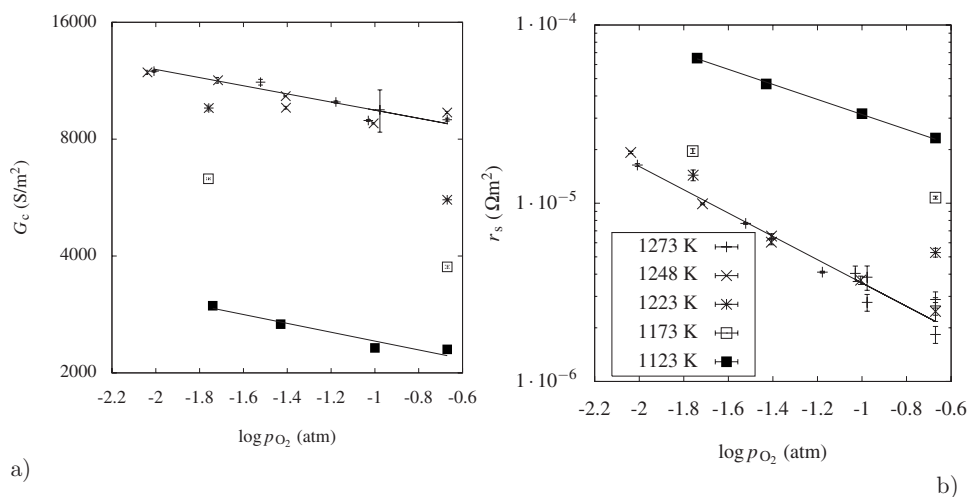


Figure 10. (a) G_c as function of $p_{O_2}^{\text{in}}$ and temperature. (b) r_s as function of $p_{O_2}^{\text{in}}$ and temperature. (a) and (b) share legends.

to G_c and r_s that will cause a deviation from activated behavior. As both G_c and r_s show no deviation from thermally activated behavior, we conclude that oxygen concentration gradients in the gas phase are not influencing our measurements.

Steady state model results.—The steady state model already described was used to extract D_v^0 from the data of G_c and r_s . This has been converted to σ_O and plotted in Fig. 12a. The fraction of $\Delta\mu_{O_2}$ spent over the bulk part of the membrane, f_{bulk} , is plotted in Fig.

12b). As can be seen, the bulk transport limits the transport at high p_{O_2} , while the surface exchange will dominate at low p_{O_2} as r_s is much larger at low p_{O_2} . The large r_s at low p_{O_2} will at $\log p_{O_2}$ (atm) < -2.5 begin to lower G_c (shown later).

The temperature dependence of D_v^0 and k_O for air ($p_{O_2}^{\text{in}} = 20.9$ kPa) and air/N₂ ($p_{O_2}^{\text{in}} = 1.8$ kPa) flow gas are shown in Fig. 13a and b, respectively. The activation energies are 62 and

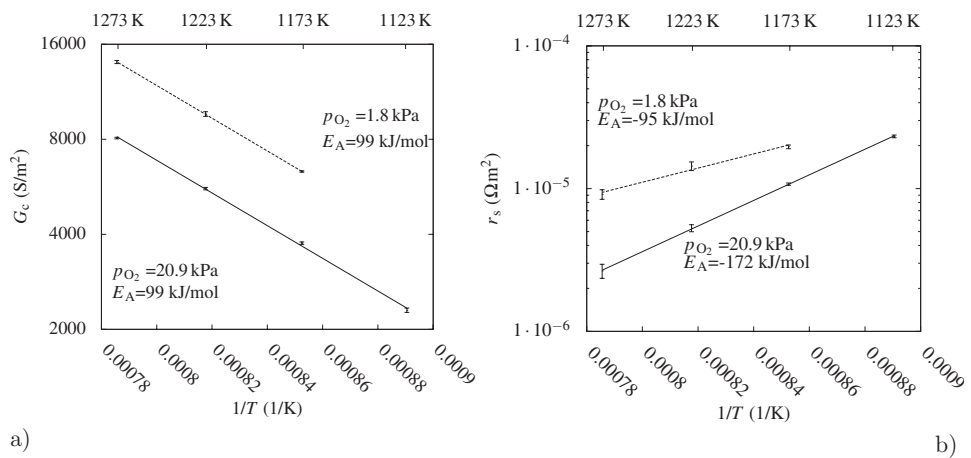


Figure 11. (a) G_c as function of $1/T$ plotted for air ($p_{O_2} = 20.9$ kPa) and air/N₂ ($p_{O_2}^{\text{in}} = 1.8$ kPa) flow gas. The activation energies are 99 kJ/mol for both gas compositions. (b) r_s as function of $1/T$ plotted for flow gas with $p_{O_2}^{\text{in}} = 20.9$ and 1.8 kPa. The activation energies for the surface reaction ($\propto r_s^{-1}$) are 172 and 95 kJ/mol, respectively.

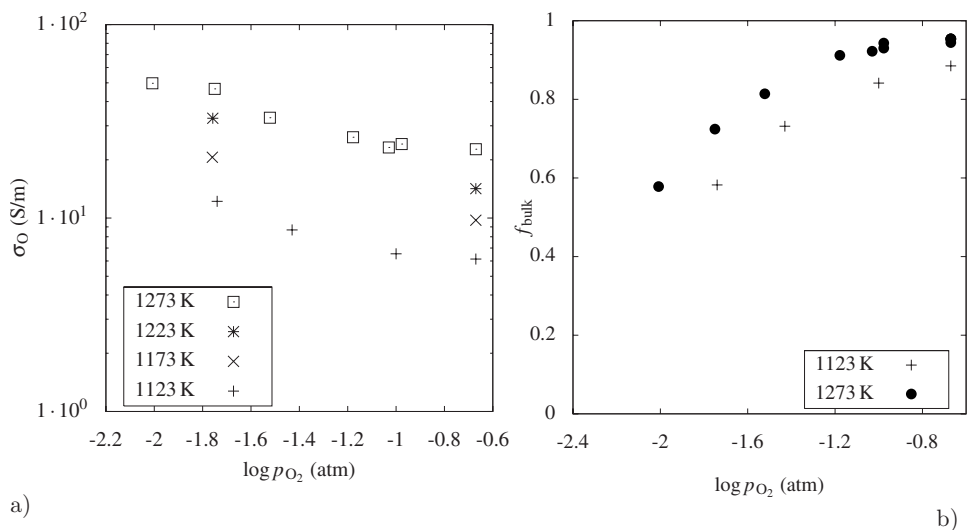


Figure 12. (a) σ_O as function of T and $\log p_{O_2}$. (b) f_{bulk} as function of T and p_{O_2} .

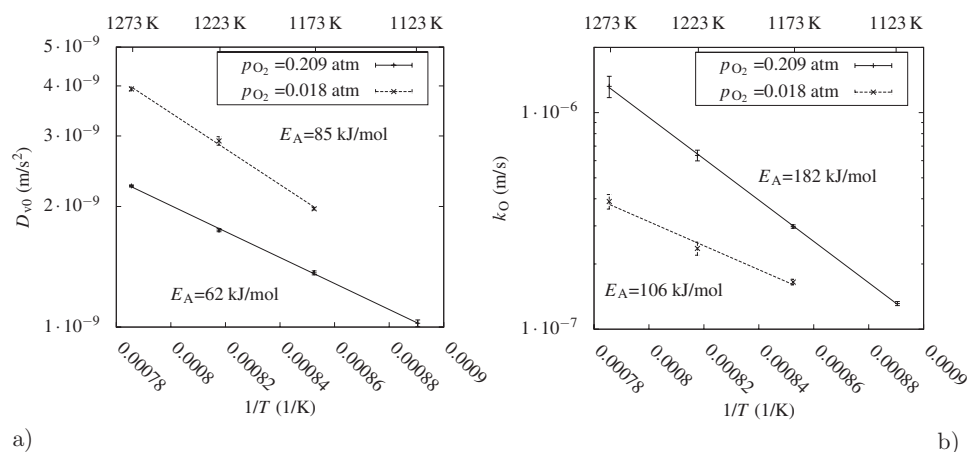


Figure 13. (a) D_v^0 as function of $1/T$ plotted for air ($p_{O_2}^{\text{in}} = 20.9$ kPa) and air/N₂ ($p_{O_2}^{\text{in}} = 1.8$ kPa) flow gas. The activation energies are 62 and 85 kJ/mol, respectively. (b) k_O as function of $1/T$ plotted for flow gases with $p_{O_2}^{\text{in}} = 20.9$ and 1.8 kPa. The activation energies for k_O are 182 and 106 kJ/mol, respectively.

85 kJ/mol for D_v^0 in the two gases. For k_O they are 182 and 106 kJ/mol, respectively; k_O was calculated from r_s using the formulas in Fig. 3.

Transport parameters of LSCF measured using steady state measurements are summarized in Table I. Values of δ and γ used in the computations are also provided.

Relaxation experiments.—Relaxation experiments were performed by making sudden changes of the composition, and thus μ_{O_2} , of the flow gas in the upper compartment, while keeping $i_{\text{pump}} = 0$. An example of μ_{O_2} in the membrane in the relaxation experiment shown in Fig. 15a at time, $t = 120$ s, is shown in Fig. 14. Measured and calculated values of V_{probe} and V_{ref} for a sudden change in the gas composition $p_{O_2}^{\text{in}}: 7.1$ kPa \rightarrow 3.2 kPa at $T = 1273$ K are plotted in Fig. 15a. Fitting parameter values of $D_v^0 = 2.9 \times 10^{-9}$ and $k_O = 5.6 \times 10^{-5}$ m/s were used. At large times the calculated and measured values of V_{ref} disagree. As the model uses many input data of finite precision (δ , membrane dimensions, voltages), this disagreement is well within the expected precision. A contour plot of the summed square deviation³¹ of the measured and calculated values is given in Fig. 15b. SSD is normalized to the best fit (SSD_{min}). The SSD is more sensitive to D_v^0 than k_O , and thus D_v^0 is determined with higher precision than k_O .

Measurement verification.—In the following we present measurements designed to rule out possible misinterpretations of the measured data.

Influence of corrections.—A few corrections were proposed earlier in the paper. These corrections are not necessarily specific to this setup, and will be equally relevant in, e.g., a conductivity relaxation experiment, where the perturbation of $p_{O_2}^{\text{in}}$ can be considerable. To illustrate the impact of these corrections G_c and r_s (i.e., corrected values) are compared to G and $V_{\text{probe}}/i_{\text{pump}}$ (i.e., uncorrected values) in Fig. 16a and b. The represent extrapolation towards lower $p_{O_2}^{\text{in}}$ lines illustrating the steady state, model prediction values have been added. These lines were calculated using a weakly p_{O_2} dependent D_v^0 and a $\partial V_{\text{probe}}/\partial i_{\text{pump}} \propto p_{O_2}^{-0.68}$. The corrections to G and $V_{\text{probe}}/i_{\text{pump}}$ in the measured $p_{O_2}^{\text{in}}$ regime are less than 15%. However, in accor-

dance with Eq. 18 and 25 the extrapolations show that these corrections will increase as the $p_{O_2}^{\text{in}}$ is reduced. Measurements in inert gas with $p_{O_2}^{\text{in}}$ below 1 kPa should therefore be interpreted with care. In regimes with high r_s (i.e., low temperature and low $p_{O_2}^{\text{in}}$), it will be possible to conduct good measurements using low pumping currents; this reduces the corrections.

Influence of gas diffusion.—The diffusion of oxygen molecules in a gas is not infinitely fast; it is therefore important to test whether limitations of gas diffusion will develop gradients in the oxygen concentration, reducing the accuracy of the measurement. i_{pump} is shown as a function of V_{ref} for two experiments in Fig. 17a. One experiment is performed in a gas mixture of 90% He + 10% air, and the other in a gas mixture of 90% N₂ + 10% air. The diffusion coefficient of O₂ in He is about four times larger than the diffusion coefficient of O₂ in N₂. Gradients in the gas above the membrane would thus be larger in the 90% N₂ + 10% air gas mixture than in 90% He + 10% air. Evidently G is independent of the gas diffusion coefficient of oxygen in the upper compartment. V_{probe} is shown in Fig. 17b as function of i_{pump} at $T = 1273$ K in the same two experiments. It can be seen that $\partial V_{\text{probe}}/\partial i_{\text{pump}}$ is also independent of the gas diffusion coefficient.

The measurements shown in Fig. 9a and b are cautious measurements in the sense that oxygen is pumped out of the compartment below the membrane (negative values of i_{pump}) and the flux magnitude is kept low; this protects the membrane and glass seal from large mechanical pressure gradients. In the measurements in Fig. 17a and b the gradients and fluxes are much larger; this allows us to observe a deviation from the linear regime of G in Fig. 17a. This deviation is only seen when pumping oxygen into the closed compartment. We believe that this is caused by leaks in the glass seal being forced open by oxygen overpressure at the membrane perimeter.

The $\partial V_{\text{probe}}/\partial i_{\text{pump}}$ shows a nonlinearity when pumping gas out of the closed compartment in Fig. 17b; this is attributed to a decrease of $p_{O_2}^{\text{in}}$ (i.e., the p_{O_2} in the gas beside the tip of the probe) and the corresponding increase in r_s . Both G and $\partial V_{\text{probe}}/\partial i_{\text{pump}}$ are independent of pumping direction for small fluxes; this shows that the

Table I. Transport and stoichiometry parameters of LSCF in air, k_{chem} is the chemical surface exchange coefficient, analogous to D_{chem} .

T (K)	D_v^0 (m ² /s)	D_{chem} (m ² /s)	σ_O (S/m)	r_s (Ω m ²)	k_{chem} (m/s)	k_O (m/s)	δ	γ
1123	1.02×10^{-9}	2.24×10^{-9}	6.14	2.32×10^{-5}	1.57×10^{-5}	1.31×10^{-7}	5.50×10^{-2}	119
1173	1.37×10^{-9}	3.50×10^{-9}	9.74	1.08×10^{-5}	3.34×10^{-5}	2.98×10^{-7}	6.85×10^{-2}	112
1223	1.75×10^{-9}	5.19×10^{-9}	14.2	5.28×10^{-6}	6.92×10^{-5}	6.34×10^{-7}	8.17×10^{-2}	109
1273	2.25×10^{-9}	7.69×10^{-9}	20.2	2.65×10^{-6}	1.44×10^{-4}	1.32×10^{-6}	9.43×10^{-2}	109

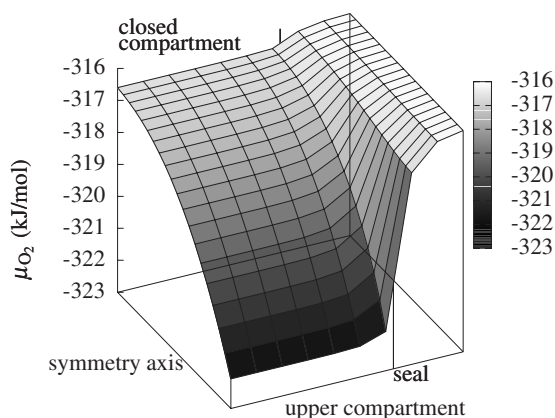


Figure 14. μ_{O_2} in the membrane in the relaxation experiment shown in Fig. 15a at time, $t = 120$ s. The total length along the symmetry axis is 2.253 mm. The total length along the upper compartment edge is 10.15 mm (half the membrane diameter), with the aperture extending 7 mm from the symmetry axis, where the gold seal blocks the flux (see Fig. 6).

rate limiting step of gas entering and leaving the membrane surface is the same. This independence of the current direction is seen in all measurements for small fluxes.

V_{probe} evaluation.—As mentioned earlier, V_{probe} has several contributions apart from V_{surface} . If electronic contact, but no ionic contact between the MIEC membrane and the probe exists, V_{probe} will be the sum of these extra contributions only. In effect V_{probe} will be mea-

suring $p_{\text{O}_2}^{\text{tip}}$, i.e., the p_{O_2} of the gas close to the contact point of the probe and membrane, and possibly some thermoelectric voltages. A measurement of V_{probe} as function of the pumping current for a membrane of the LSCF derived MIEC material ($\text{La}_{0.6}\text{Sr}_{0.2}\text{Ba}_{0.2}\text{Fe}_{0.8}\text{Co}_{0.2}\text{O}_{3-\delta}$ (LSB20CF) with a light load applied to the probe is compared to a measurement on LSCF with a heavy load applied to the probe, in Fig. 18. The $V_{\text{probe}}/i_{\text{pump}}$ was about 20 times less for the lightly loaded LSB20CF than for the heavily loaded LSCF (Fig. 18a). Except for the load on the probe and the sample composition, all other measurement parameters were equivalent. We interpret the low V_{probe} as a sign of poor ionic contact to the membrane. In conclusion, with good ionic contact to the membrane, V_{grad} is small compared to the contribution from V_{surface} . Our recommendation is to apply a pressure of 10–100 MPa to ensure ionic contact.

Measurement reproducibility.—While the precision error bars on the individual measurements in Fig. 10a and b, 11a, and b are smaller than the symbols used to mark the measurements, some scatter still exists. We believe that this is caused by slow relaxation of the membrane surface and microstructure. First, a series with varying $p_{\text{O}_2}^{\text{in}}$ was run at $T = 1273$ K. Neither G_c or r_s were reproducible. Another series was then run at $T = 1248$ K. This series was reasonably reproducible. After a month of measurements a new series was recorded at $T = 1273$ K. The values for G_c or r_s agreed reasonably with the last eight measurements of the first $T = 1273$ K series. We conclude that more than one week of equilibration is needed before the membrane is stable. This is also seen

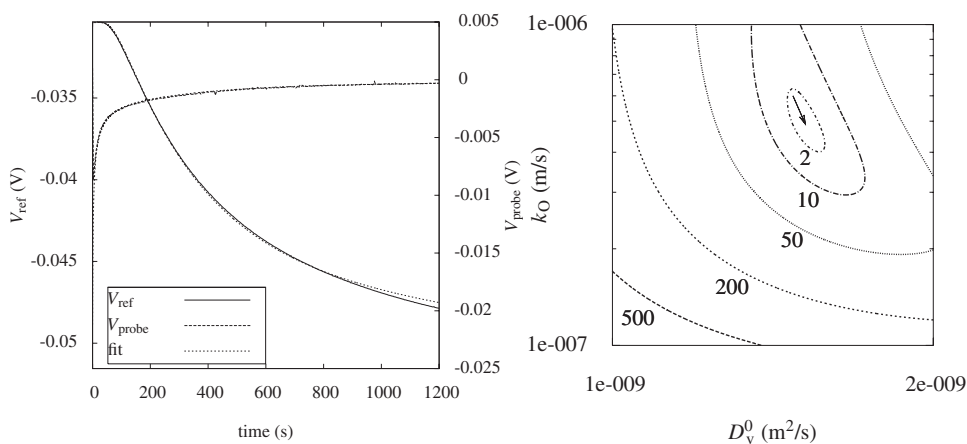


Figure 15. (a) Measured V_{probe} and V_{ref} together with calculated values of V_{probe} and V_{ref} for a sudden gas composition change $p_{\text{O}_2}^{\text{in}}$: 7.1 kPa \rightarrow 3.2 kPa at $T = 1273$ K. (b) Contour plot of $\text{SSD}/\text{SSD}_{\text{min}}$ as function of D_v^0 and k_{O} .

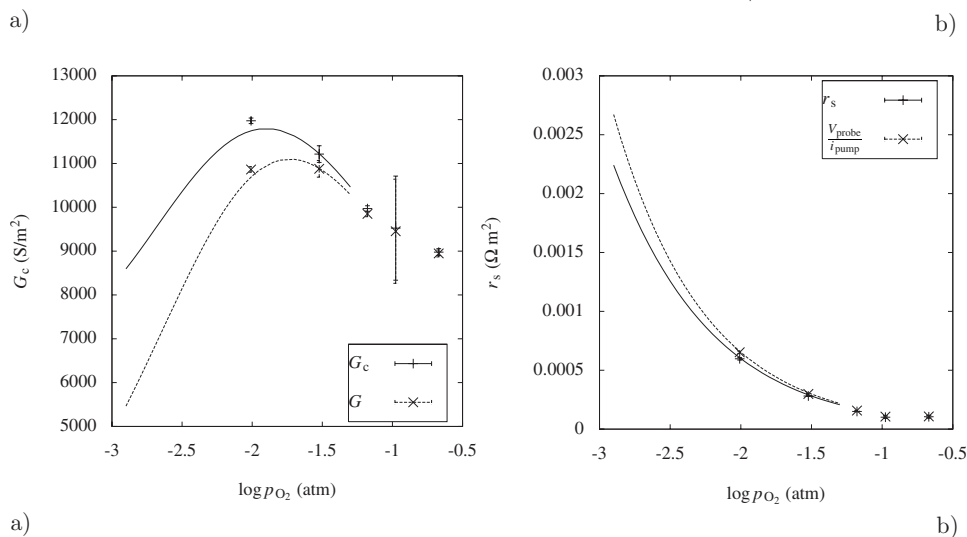


Figure 16. (a) G_c and G at 1273 K as a function of $\log p_{\text{O}_2}^{\text{in}}$. The lines are model predictions. (b) r_s and $\partial V_{\text{probe}}/\partial i_{\text{pump}}$ at 1273 K as a function of $\log p_{\text{O}_2}^{\text{in}}$. The lines are model predictions.

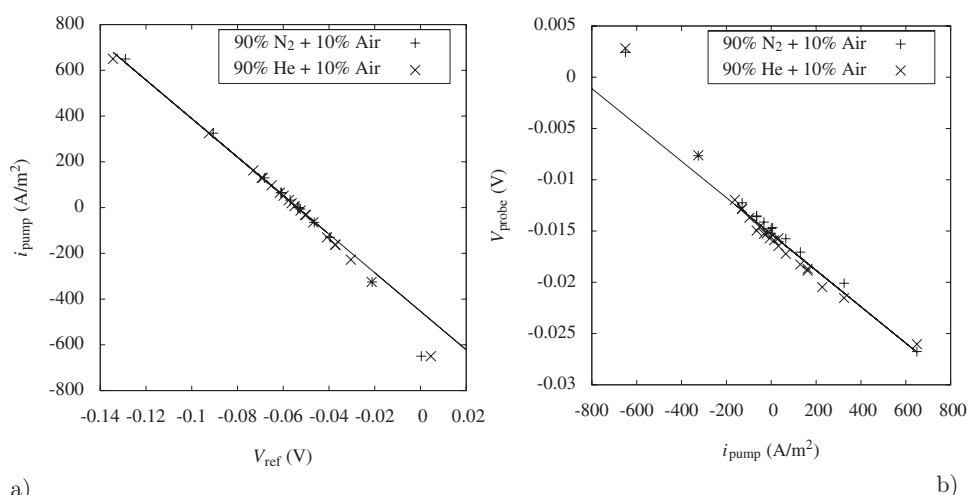


Figure 17. Measurements performed in a gas mixture of 90% He + 10% air (i.e., 2% O₂ + 8% N₂), and in 90% N₂ + 10% air. (a) i_{pump} as function of V_{ref} at 1273 K. (b) V_{probe} as function of i_{pump} at 1273 K.

with other membrane materials in this setup, and we attribute it to relaxation of the microstructure and surface, as well as possible cation diffusion.

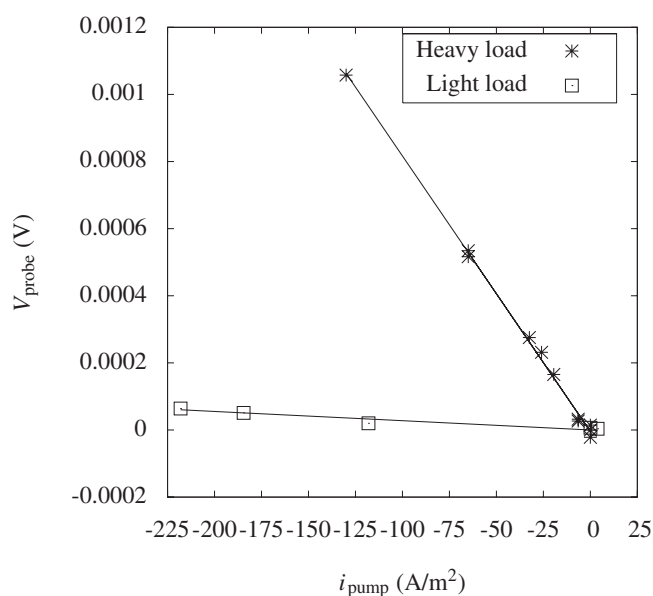


Figure 18. Measurement of V_{probe} at $p_{\text{O}_2} = 4$ kPa, $T = 1273$ K on $(\text{La}_{0.6}\text{Sr}_{0.4})_{0.99}\text{Co}_{0.2}\text{Fe}_{0.8}\text{O}_{3-\delta}$ (heavy load) and on $(\text{La}_{0.6}\text{Sr}_{0.2}\text{Ba}_{0.2})_{0.99}\text{Fe}_{0.8}\text{Co}_{0.2}\text{O}_{3-\delta}$ (light load).

Comparison of values from steady state and relaxation experiments.— An advantage of having both steady state and relaxation measurements performed in the same setup, is that comparison of the results from each measurement type can point out discrepancies in the modeling. D_{v}^0 as deduced from the steady state measurements and the relaxation experiments are plotted in Fig. 19a. As mentioned earlier, D_{v}^0 should not be directly dependent on C_{v} , and should thus be a measure of the mobility of the vacancies. The values obtained using steady state measurements show an increase in D_{v}^0 at low p_{O_2} . The values for D_{v}^0 obtained using relaxation experiments are constant when $p_{\text{O}_2}^{\text{in}}$ is reduced at $T = 1123$ K and decrease slightly when $p_{\text{O}_2}^{\text{in}}$ is reduced at $T = 1248$ K and $T = 1273$ K. While the D_{v}^0 values in air are very close and $p_{\text{O}_2}^{\text{in}}$ dependencies of the steady state and relaxation values are not too different, the agreement is not complete. The conversion from D_{chem} to D_{v}^0 makes use of the thermodynamic factor γ as well as δ . Any error in these parameters will accumulate in D_{v}^0 . While the coulometric titration gives a high precision on $\gamma = \partial \ln p_{\text{O}_2} / 2 \partial \ln c_{\text{O}}$, the absolute value of δ is difficult to obtain as illustrated by the variation in the values reported in the literature (cf. Tai,³² Den Otter,³³ or Stevenson¹²). Furthermore, D_{v}^0 and k_{O} are linearly correlated in a steady state experiment, while the correlation is more complicated in the fit of a relaxation experiment. This could give a systematic difference in the results of steady state measurements and relaxation measurements.

The k_{O} values calculated from the steady state and the relaxation experiments are plotted in Fig. 19b. The steady state and relaxation values of k_{O} are in excellent agreement, if it is assumed that it is the

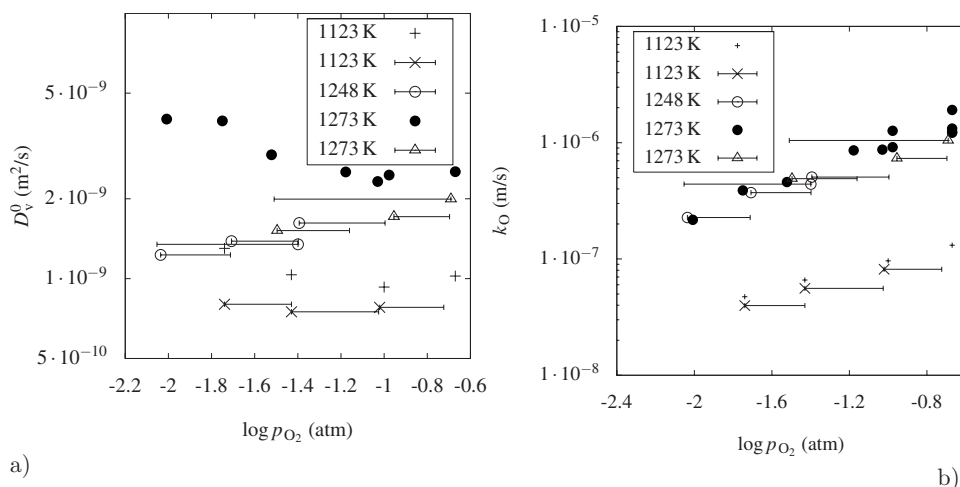


Figure 19. (a) D_{v}^0 calculated from steady state and relaxation experiments. The marks connected to lines are relaxation experiments, where the marks signify the end $p_{\text{O}_2}^{\text{in}}$ value of a relaxation and the bar spanning the p_{O_2} variation of the relaxation. (b) k_{O} calculated from the steady state and the relaxation experiments.

end $p_{O_2}^{\text{in}}$ that determines k_O in the relaxation. This is expected, as the $\mu_{O_2}^{\text{MIEC}}$ just below the surface will adjust its $\mu_{O_2}^{\text{MIEC}}$ to a value close to the $\mu_{O_2}^{\text{MIEC}}$ of the flow gas during the first tenth of the fitted relaxation time; k_O is proportional to $p_{O_2}^{0.41}$ at $T = 1123$ K and $p_{O_2}^{0.52}$ at $T = 1273$ K. It is worth noting that the precision of the measurement of the steady state values of k_O is on the order of the symbol size or less, even when measuring at $T = 1273$ K and $p_{O_2} = 0.21$ atm.

The increase in D_v^0 with lower $p_{O_2}^{\text{in}}$ in the steady state measurements could be caused by misinterpretation of the measured data. At low p_{O_2} surface oxygen exchange processes become more important in determining the oxygen transport (about 45% of the μ_{O_2} difference is found at the surfaces at $p_{O_2}^{\text{in}} = 1$ kPa and $T = 1123$ K), and the increase in D_v^0 could be related to an overestimation of r_s . As the $(V_{\text{probe}}/i_{\text{pump}})$ plots used to determine r_s are close to linear at all measured $p_{O_2}^{\text{in}}$ values, terms causing an overestimation not accounted for must have the same linearity.

One error term with such a linearity could be caused by the probe penetrating deep into the membrane and probing μ_{O_2} some distance into the bulk. Another term could come from the probe blocking the flux and altering the μ_{O_2} profile through the membrane, resulting in an enhanced V_{probe} . To investigate these possibilities, each of the relaxation measurements were fitted again, but this time with the cone depth as a fitting parameter. No common cone probe depth was, however, found for the relaxations.

Thermovoltages have also been dismissed as an error term. While heat from the oxidation reaction will change the temperature at the probe tip, the enthalpy of oxidation should only be weakly temperature dependent and as the thermopower of the CG10 probe increases with temperature,³⁴ the thermovoltage would not experience the decrease of r_s with temperature seen in Fig. 11b. Furthermore, localized ohmic heating in the membrane will scale with $(i_{\text{pump}})^2$ in contrast with the linear relations seen in Fig. 9b. Finally the Seebeck coefficient is less than 20 $\mu\text{V}/\text{K}$.

Conclusion

A setup combining an oxygen pump with an electrolyte probe was demonstrated. From this setup the μ_{O_2} difference across a LSCF membrane and across the LSCF membrane surface could be determined from the voltage across a pair of reference electrodes on the oxygen pump and from the voltage across the electrolyte probe in contact with the membrane surface. During steady state the voltages had a linear dependence on the oxygen flux. The area specific conductance, G_c (a measure of the permeability of the membrane), and the surface resistance, r_s , could thus be extracted. G_c and r_s were found to be independent of the flux direction, showing that the rate limiting step of the oxygen exchange reaction on the surface is the same when oxygen is incorporated into the membrane and when it is released from the membrane. From G_c and r_s a model was used to extract values of the ionic conductivity, σ_O and the concentration independent vacancy diffusion factor D_v^0 . Excellent precision was achieved using small perturbations, illustrating the advantages of (i) measuring at steady state (with very large sampling times), (ii) direct electronic measurement of the oxygen flux, (iii) direct electronic measurement of the μ_{O_2} difference across the membrane surface and (iv) direct electronic measurement of the total μ_{O_2} difference across the membrane. Furthermore, precise determination of r_s was not limited to low temperature and low p_{O_2} , where r_s is normally dominating the transport, but could also be recorded at high temperature and p_{O_2} , where it is usually too small to determine accurately. A model based on Ficks law of diffusion was used to fit the measured voltages after stepwise gas composition changes and thus extract D_v^0 and k_O . The values extracted from steady state and relaxation measurements were reasonably consistent.

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Table A-I. Value of n_i to be used in Eq. A-1 and A-2.

n_1	1.06778
n_2	3.50042
n_3	0.166961×10^{-7}
n_4	1.01258
n_5	0.944365
n_6	2242.45 K
n_7	11,580.4 K

Appendix A

Chemical Potential of Oxygen in the Gas Phase

The following equations and values are all taken from Oxygen, IUPAC International Tables of the Fluid State 9 in Ref. 24.

The entropy, $S^{\text{id}}(T)$, and the enthalpy, $H^{\text{id}}(T)$, of an ideal gas at the temperature, T , and the pressure $P_A = 1$ atm can be found from the equations below

$$S^{\text{id}}(T, P_A) = R \left[-\frac{2n_1}{3} \left(\frac{T}{K}\right)^{-3/2} + n_2 \ln\left(\frac{T}{K}\right) + \frac{n_3}{2} \left(\frac{T}{K}\right)^2 + n_4 \left(\frac{ue^u}{e^u - 1} - \ln(e^u - 1)\right) + n_5 \left(\frac{\frac{2}{3}\varphi e^{-\eta}}{1 + \frac{2}{3}e^{-\eta}} + \ln\left(1 + \frac{2}{3}e^{-\eta}\right) \right) \right]_{298.15 \text{ K}} + 205.146 \text{ J mol}^{-1} \text{ K}^{-1} \quad \text{[A-1]}$$

$$H^{\text{id}}(T, P_A) = R(1(K)) \left[-2n_1 \left(\frac{T}{K}\right)^{-1/2} + n_2 \frac{T}{K} + \frac{n_3}{3} \left(\frac{T}{K}\right)^3 + n_4 \frac{u \left(\frac{T}{K}\right)}{e^u - 1} - n_5 \frac{\eta \left(\frac{T}{K}\right)}{1 + \frac{2}{3}e^{-\eta}} \right]_{298.15 \text{ K}} \quad \text{[A-2]}$$

In the above equations $u = n_6/T$ and $\eta = n_7/T$ and the values of $n_1, n_2, n_3, n_4, n_5, n_6,$ and n_7 can be found in Table A-I.

Using these formulas the entropy and the enthalpy is determined at a temperature T . The term 205.146 J mol⁻¹ K⁻¹ (in Eq. A-1) originates from the entropy difference between the state of the ideal crystal at 0 K and that of the ideal gas at (298.15 K, 0.1 MPa).

The chemical potential of oxygen in the gas phase is then calculated as

$$\mu_{O_2} = \mu_{O_2}^{\ominus} + RT \ln\left(\frac{P}{P_A}\right) \quad \text{[A-3]}$$

where

$$\mu_{O_2}^{\ominus} = H^{\text{id}}(T, P_A) - TS^{\text{id}}(T, P_A) \quad \text{[A-4]}$$

Appendix B

Derivation of the Average Driving Force above the Membrane

When a molar flow of oxygen, J_{O_2} [mol s⁻¹], passes through the membrane it will change the p_{O_2} in the upper compartment

$$p_{O_2} = p_{O_2}^{\text{in}} + \frac{J_{O_2}}{J_{N_2/O_2} + J_{O_2}} \approx p_{O_2}^{\text{in}} + \frac{J_{O_2}RT\Phi_f}{P\Phi_f} \quad \text{[B-1]}$$

where J_{N_2/O_2} is the molar flow of the nitrogen/air mixture, led from the external gas supply, through the tubes, into the upper compartment. P is the absolute pressure, Φ_f is the volume flow, and T_{Φ_f} is the T where volume flow is measured.

The gas flow enters the upper compartment above the center of the membrane and moves towards the periphery. The p_{O_2} above the center of the membrane will thus be close to the p_{O_2} above the center of the membrane will thus be close to the p_{O_2} of the gas entering the chamber, while the p_{O_2} of the gas at the periphery will be close to the gas leaving the chamber. V^{upper} will thus depend on the distance from the center of the membrane, r

$$V^{\text{upper}}(r) = \frac{RT}{4F} \ln\left(\frac{p_{O_2}^{\text{in}} + \frac{J_{O_2}(r)RT\Phi_f}{P\Phi_f}}{p_{O_2}^{\text{ref}}}\right) \quad \text{[B-2]}$$

where $J_{O_2}(r)$ is the amount of oxygen passing through the membrane inside a circle with radius r . $J_{O_2}(r)$ is a fraction of J_{O_2}

$$J_{\text{O}_2}(r) = \int_0^r J_{\text{O}_2} \frac{2\pi r'}{\pi r_0} dr' = \frac{r^2}{r_0^2} J_{\text{O}_2} \quad [\text{B-3}]$$

where r_0 is the total exposed membrane surface. The average V^{upper} on the surface will be

$$V^{\text{upper}} = \frac{1}{\pi r_0^2} \int_0^{r_0} V^{\text{upper}}(r) 2\pi r dr \quad [\text{B-4}]$$

which leads to Eq. 18.

References

1. *The CRC Handbook of Solid State Electrochemistry*, P. Gellings and H. Bouwmeester, Editors, CRC, Boca Raton, FL (1997).
2. S. J. Skinner and J. A. Kilner, *Mater. Today*, **6**, 30 (2003).
3. P. V. Hendriksen, P. H. Larsen, M. Mogensen, F. W. Poulsen, and K. Wiik, *Catal. Today*, **56**, 283 (2000).
4. H. J. M. Bouwmeester, *Catal. Today*, **82**, 141 (2003).
5. I. Yasuda and T. Hikita, *J. Electrochem. Soc.*, **141**, 1268 (1994).
6. J. ten Elshof, M. Lankhorst, and H. Bouwmeester, *Solid State Ionics*, **99**, 15 (1997).
7. J. Lane and J. Kilner, *Solid State Ionics*, **136-137**, 997 (2000).
8. M. Sogaard, P. V. Hendriksen, M. Mogensen, F. W. Poulsen, and E. Skou, *Solid State Ionics*, **177**, 3285 (2006).
9. S. Carter, A. Selcuk, R. J. Chater, J. Kajda, J. A. Kilner, and B. C. H. Steele, *Solid State Ionics*, **53-56**, 597 (1992).
10. S. J. Benson, R. J. Chater, and J. A. Kilner, *Proc. Electrochem. Soc. PV 97-24*, p. 596 (1999).
11. J. ten Elshof, H. Bouwmeester, and H. Verweij, *Solid State Ionics*, **81**, 97 (1995).
12. J. W. Stevenson, T. Armstrong, R. Carnheim, L. Pederson, and W. Weber, *J. Electrochem. Soc.*, **143**, 2722 (1996).
13. J. Lane, S. Benson, D. Waller, and J. Kilner, *Solid State Ionics*, **121**, 201 (1999).
14. M. Lankhorst, Ph.D. Thesis, Universiteit Twente, Enschede (1997).
15. M. Sogaard, P. V. Hendriksen, and F. W. Poulsen, in *Proceedings of the 26th Risø International Symposium on Materials Science*, 1, 355 (2005).
16. Y. L. Yang, A. J. Jacobsen, C. L. Chen, G. P. Luo, K. D. Ross, and C. W. Chu, *Appl. Phys. Lett.*, **79**, 776 (2001).
17. B. Zachau-Christiansen, T. Jacobsen, and S. Skaarup, *Solid State Ionics*, **86-88**, 725 (1996).
18. A. Belzner, T. M. Gür, and R. A. Huggins, *Solid State Ionics*, **40-41**, 535 (1990).
19. S. Diethelm, Ph.D. Thesis, Lausanne EPFL (2001).
20. D.-K. Lee and H.-I. Yoo, *J. Electrochem. Soc.*, **147**, 2835 (2000).
21. J. Fouletier, P. Fabry, and M. Kleitz, *J. Electrochem. Soc.*, **123**, 204 (1976).
22. H.-D. Wiemhöfer, H.-G. Bredes, U. Nigge, and W. Zipprich, *Solid State Ionics*, **150**, 63 (2002).
23. L. A. Chick, G. L. Maupin, L. R. Pederson, D. E. McCready, and J. L. Bates, *Mater. Res. Soc. Symp. Proc.*, **249**, 159 (1992).
24. W. Wagner, K. M. de Reuck, Oxygen, IUPAC International Tables of the Fluid State 9.
25. J. Maier, *Solid State Ionics*, **112**, 197 (1998).
26. R. Tilley, *Principles and Applications of Chemical Defects*, Stanley Thornes Ltd., Cheltenham, U.K. (1998).
27. C. Wagner, *Prog. Solid State Chem.*, **10**, 3 (1975).
28. J. Crank, *The Mathematics of Diffusion*, Oxford University Press, Oxford (1975).
29. R. C. Reid, J. M. Prausnitz, and B. E. Polling, *The Properties of Gases and Liquids*, McGraw-Hill, New York (1987).
30. M. Sogaard, P. V. Hendriksen, and M. Mogensen, *J. Solid State Chem.*, **180**, 1489 (2007).
31. P. Bevington and D. Robinson, *Data Reduction and Error Analysis for the Physical Sciences*, 3rd ed., McGraw-Hill, New York (2003).
32. L.-W. Tai, M. Nasrallah, H. U. Anderson, D. Sparlin, and S. Sehlin, *Solid State Ionics*, **76**, 273 (1995).
33. M. den Otter, Ph.D. Thesis, Universiteit Twente, Enschede (2000).
34. H.-T. Lim and A. V. Virkar, *J. Power Sources*, **161**, 676 (2006).