



Selective flue gas cleaning with ionic liquids

Kegnæs, Søren; Due-Hansen, Johannes; Harris, Pernille; Berg, Rolf W.; Riisager, Anders; Fehrmann, Rasmus

Publication date:
2010

Document Version
Publisher's PDF, also known as Version of record

[Link back to DTU Orbit](#)

Citation (APA):
Kegnæs, S., Due-Hansen, J., Harris, P., Berg, R. W., Riisager, A., & Fehrmann, R. (2010). *Selective flue gas cleaning with ionic liquids*. Poster session presented at Conference on Molten Salts and Ionic Liquids 2010, Bamberg, Germany.

General rights

Copyright and moral rights for the publications made accessible in the public portal are retained by the authors and/or other copyright owners and it is a condition of accessing publications that users recognise and abide by the legal requirements associated with these rights.

- Users may download and print one copy of any publication from the public portal for the purpose of private study or research.
- You may not further distribute the material or use it for any profit-making activity or commercial gain
- You may freely distribute the URL identifying the publication in the public portal

If you believe that this document breaches copyright please contact us providing details, and we will remove access to the work immediately and investigate your claim.

Søren Kegnæs, Johannes Due-Hansen, Pernille Harris, Rolf W. Berg, Anders Riisager and Rasmus Fehrmann.

Centre for Catalysis and Sustainable Chemistry, Department of Chemistry, Technical University of Denmark, DK-2800 Kgs. Lyngby, Denmark. E-mail: skk@kemi.dtu.dk

Introduction

- Combustion of fossil fuels leads to the emission of acidic gases e.g. NO_x and SO₂ which is a major concern in relation to atmospheric pollution.
- These gases have to be effectively removed from flue gases.
- It is difficult to find adsorbents which allow reversible and selective absorption of acidic gases.
- Pressure- and temperature swing adsorption processes, provide excellent resources for production of sulfuric acid and nitric acid.
- Ionic liquids (ILs) have in no (or negligible) vapour pressure and large liquidus.
- The structures of ILs are well-ordered even in liquid state with regular cavities which can host selected solute species depending on the IL ion composition.
- Therefore IL can be tuned to absorb selected gas molecules making them promising materials for selective, reversible absorption of gaseous pollutants in e.g. power plant flue gases and other exhaust or industrial off-gases.^[1,2,3,4]

Structural Characterisation

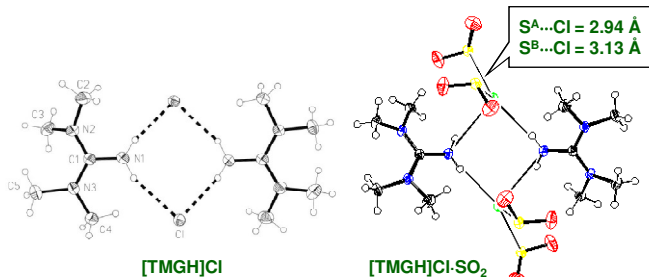


Fig. 1. X-ray crystal structures of [TMGH]Cl^a (left) and [TMGH]Cl·SO₂ (right)^[5]

Table 1. Selected hydrogen bond distances to chloride ions^[5]

Compound	D-H...A (Å)	D-H (Å)	H...A (Å)	D...A (Å)	D-H...A (°)
[TMGH]Cl ^a	N-H ^A ...Cl	0.92(2)	2.42(2)	3.29(2)	158.4(1)
(δ = 1.22 g/cm ³)	N-H ^B ...Cl ^b	0.87(2)	2.37(2)	3.23(2)	168.9(1)
[TMGH]Cl·SO ₂	N-H ^B ...Cl	0.87(2)	2.49(2)	3.26(2)	158.4(1)
(δ = 1.38 g/cm ³)	N-H ^A ...Cl	0.86(2)	2.46(2)	3.22(2)	168.9(1)

^a [TMGH]Cl: 1,1,3,3-Tetramethylguanidinium chloride
^b Fischer and Jones, *Acta Cryst. E*, 58, 2002, o218

SO₂ absorption with SILP absorber

- Supported Ionic Liquid-Phase (SILP) absorber: [TMGH]Cl (0-50 wt%) on high surface area SiO₂
- Fixed-bed reactor: 100 ml/min SO₂ (0-30.000 ppm) in N₂.
- Absorption / desorption temperatures 30-110 °C.

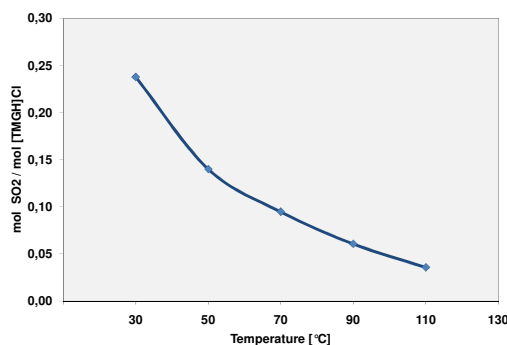


Fig. 2. Absorption of SO₂ gas in 20 wt% [TMGH]Cl on supported SiO₂ at different temperatures.^[6]

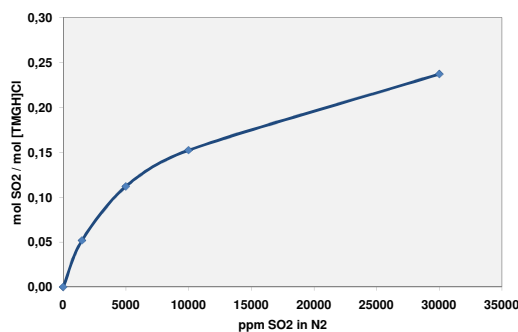


Fig. 3. Absorption of different concentrations of SO₂ gas in 20 wt% [TMGH]Cl on supported SiO₂.^[6]

Conclusions

- The results show that SO₂ can be reversible and selective adsorbed using [TMGH]Cl.
- Supported Ionic Liquid-Phase (SILP) absorbers are promising materials for flue gas cleaning.
- Absorption / desorption can be tuned with temperatures, pressures and gas concentrations.

References

- [1] J. Huang, A. Riisager, P. Wasserscheid, R. Fehrmann, *Chem. Commun.*, 2006, 4027-4029.
- [2] W. Wu, B. Han, H. Gao, Z. Liu, T. Jiang, J. Huang, *Angew. Chem. Int. Ed.*, 2004, 43, 2415-2417.
- [3] Z. Zhang, L. Wu, J. Dong, B.-G. Li, S. Zhu, *Ind. Eng. Chem. Res.*, 2009, 48, 2142-2148.
- [4] J. Huang, A. Riisager, R. W. Berg, R. Fehrmann, *J. Mol. Catal. A*, 2008, 279, 170-176.
- [5] A. Riisager, P. Harris, R. W. Berg, R. Fehrmann, submitted.
- [6] S. Kegnæs, A. Riisager, R. Fehrmann, submitted.