



Positron Lifetimes in Pure and Doped Ice and in Water

Eldrup, Morten Mostgaard; Mogensen, O.; Trumpy, Georg

Published in:
Journal of Chemical Physics

Link to article, DOI:
[10.1063/1.1677990](https://doi.org/10.1063/1.1677990)

Publication date:
1972

Document Version
Publisher's PDF, also known as Version of record

[Link back to DTU Orbit](#)

Citation (APA):
Eldrup, M. M., Mogensen, O., & Trumpy, G. (1972). Positron Lifetimes in Pure and Doped Ice and in Water. *Journal of Chemical Physics*, 57(1), 495-504. <https://doi.org/10.1063/1.1677990>

General rights

Copyright and moral rights for the publications made accessible in the public portal are retained by the authors and/or other copyright owners and it is a condition of accessing publications that users recognise and abide by the legal requirements associated with these rights.

- Users may download and print one copy of any publication from the public portal for the purpose of private study or research.
- You may not further distribute the material or use it for any profit-making activity or commercial gain
- You may freely distribute the URL identifying the publication in the public portal

If you believe that this document breaches copyright please contact us providing details, and we will remove access to the work immediately and investigate your claim.

Positron Lifetimes in Pure and Doped Ice and in Water

M. ELDRUP AND O. MOGENSEN

Chemistry Department, Danish Atomic Energy Commission Research Establishment, Risø, Roskilde, Denmark

AND

G. TRUMPY

Laboratory of Applied Physics II, Technical University of Denmark, Lyngby, Denmark

(Received 8 February 1972)

Positron lifetime spectra were measured in mono- and polycrystalline light ice, polycrystalline heavy ice, doped light ice, as well as in light and heavy water. All spectra were resolved into three components. At temperatures between -196° and -100°C the lifetimes and relative intensities are virtually constant, being 0.12 nsec and 20%, 0.44 nsec and 28%, 0.68 nsec and 52%, and attributed to annihilation of para-Ps, free positrons, and ortho-Ps, respectively. Above -100°C the longest lifetime increases with temperature to 1.1 nsec at 0°C . Its relative intensity is unchanged, while the parameters for the other components show a complex behavior. The spectra for mono- and polycrystalline light ice and for polycrystalline heavy ice are identical. For water long lifetime components attributed to ortho-Ps are 1.86 nsec, 27% for H_2O and 2.01 nsec, 22% for D_2O . Theoretical explanations are suggested. Fast frozen solutions of HF with concentrations larger than approximately 10^{-5} mole fraction have a lifetime component at -160°C of 1.23 nsec, 52% attributed to trapping of ortho-Ps in amorphous regions. Very good agreement is found between the experiments and a trapping model. Irreversible changes of the spectra are found by heating above approximately -120°C . Measurements on a number of fast frozen aqueous solutions of acids, bases, and salts are reported, none of them showing as strong influence on the ortho-Ps lifetime as HF.

I. INTRODUCTION

A positron entering a solid is nearly always thermalized before annihilating with an electron into photons. The photon properties depend on the electron-positron state on annihilation. Hence experimental studies of the annihilation photons^{1,2} may give useful information about the structure of the solid. In particular the interaction of positrons with defects has been a subject of interest for some years past. Recent experiments have indicated that for defect concentration above approximately the parts per million level, the positrons may become trapped in the defects (e.g., vacancies in metals,^{3,4} F centers, and cation vacancies in ionic crystals^{5,6}) before annihilation.

Two experimental methods¹ are mainly used in positron annihilation research: the positron lifetime technique (see next section) and the angular correlation technique. The two photons in the 2γ annihilation are emitted in nearly opposite directions. An angular correlation setup determines the distribution of photon pairs as function of the angle θ , where $\pi - \theta$ is the angle between the wave vectors of the two photons. In a rough first approximation θ is proportional to the center-of-mass momentum of the electron-positron pair.

In ice about 70% of the positrons form a hydrogenlike atom, positronium (Ps: $\frac{1}{4}$ para-Ps and $\frac{3}{4}$ ortho-Ps) before annihilation. The positrons will therefore in ice mainly decay in three different ways: (1) by "free" annihilation with the outer electrons of the molecules (intensity $I \approx 30\%$, lifetime $\tau \approx 0.4$ nsec, gives a broad angular correlation curve); (2) by ortho-Ps pickoff annihilation, i.e., the positron annihilates with an outer electron of the molecules during an ortho-Ps molecule collision ($I \approx 52\%$, $\tau = 0.7$ – 1.3 nsec, gives a broad angular correlation curve); and (3) by para-Ps intrinsic

annihilation ($I \approx 18\%$, $\tau \approx 0.12$ nsec; some para-Ps pickoff annihilation is also found). A narrow central and narrow sidepeaks in the angular correlation curves⁷ for single crystals of ice at low temperatures have been interpreted as due to intrinsic para-Ps annihilation from a center-of-mass Bloch function state, i.e., para-Ps (and hence also ortho-Ps) is delocalized on annihilation. Para-Ps trapped in defects will give a somewhat broader central component of the angular correlation curve due to a high zero-point energy in the trap. The sidepeaks disappear on trapping. Because of reduced pickoff annihilation, trapping of Ps increases the ortho-Ps lifetime.

In this Article we discuss the results of positron lifetime studies of the motion of Ps in ice and doped ice. We have found that Ps is trapped in defects at higher temperatures and for high enough concentrations of impurities (mainly HF doping has been studied). The work is part of an experimental series on positron annihilation in ice and doped ice. In Ref. 8 we showed that a change in ortho-Ps lifetime and yield takes place at the melting point for both light and heavy water. Reference 7 established the fact that Ps is delocalized in ice at low temperatures. Angular correlation curves for frozen aqueous solutions were published in Ref. 9. Work on angular correlation as function of temperature in pure ice,¹⁰ on angular correlation and lifetime studies of HF-doped ice single crystals, and on angular correlation magnetic quenching in ice¹¹ are in progress.

II. EXPERIMENTAL METHODS

A ^{22}Na positron source emits a 1.27 MeV photon nearly simultaneously with the emission of the positron. The positron lifetime is determined as the difference in time between detection of the 1.27 MeV photon and

one of the 0.51 MeV annihilation photons. The measured distribution of lifetimes is known as the lifetime spectrum. Each annihilation process contributes to the spectrum a decaying exponential, its decay constant being the inverse of the mean lifetime.

For the measurements we used a conventional¹ fast-slow delayed-coincidence system mainly consisting of Ortec electronic units and a Nuclear Data 2200 multichannel analyzer with 4096 channels of which 512 were used for recording a spectrum. The width of the prompt curve was 0.425 nsec FWHM. As a source ²²NaCl, of approximately 50 μ Ci sandwiched between two foils of Mylar, was used. The polycrystalline ice samples were prepared by freezing the water in a cylindrical brass container of dimensions 10 mm diameter and 15 mm height. Before freezing, the source was inserted as a plane through the cylinder axis, so that it would be lying between two half-cylinders of ice. The light water used for the polycrystalline samples was distilled twice. Its specific conductivity was approximately 10^{-6} ohm⁻¹.cm⁻¹ and it contained an equilibrium amount of air. In order to investigate the possible influence of dissolved oxygen, samples were prepared containing no dissolved oxygen. However, within the experimental uncertainties no difference was found between these samples and ordinarily prepared ones. The heavy water had a specific conductivity of less than 10^{-5} ohm⁻¹.cm⁻¹ before freezing. The monocrystalline samples were cut from a larger monocrystal to dimensions $10 \times 10 \times 4$ mm³. After being ground and polished two such pieces of crystal were assembled in a sample container with the source between them.

The frozen solutions were prepared by immersing the brass container with the solution in liquid nitrogen. The container reached liquid nitrogen temperature in about 1 min. This way of cooling will be referred to as "fast freezing." All spectra have been corrected for annihilation in the source, which amounted to 7% to 10% of the total number of annihilations.¹²

III. DATA ANALYSIS

The lifetime spectra were analyzed by use of the least-squares fitting computer program POSITRONFIT described in Ref. 12. In principle it fits to an experimental spectrum a curve composed of a sum of decaying exponentials and a constant background. The results of the fitting procedure are the reciprocal decay constant (i.e., lifetime) and the relative intensity of each component (i.e., the area under each exponential in percent of the whole area). The goodness of the fit is expressed by the quantity, the "variance of the fit." It is essentially a normalized sum of the squares of the deviations between the fitted curve and the measured spectrum. For a fit to be satisfactory, the variance of the fit should be close to unity, since perfect fits will give variances which have a normal distribution, with a mean value of unity and with a standard deviation which for the present measurements is 0.072. A large

deviation from unity suggests that the resulting parameters are not a good representation of the measured spectrum.

One of the experiences gained from the analysis of our measurements was that all spectra could be fitted completely by three exponential terms, i.e., the variance of a fit including four terms was nearly the same as that for three terms. On the other hand a two-term fit was always appreciably poorer than a three-term fit except for a few special cases.

These facts do not necessarily mean that all spectra contained no more than three lifetime components, but only that the number of fitting parameters were sufficient to give a complete representation of a measured spectrum. If a spectrum is composed of more than three components the question then arises as to the physical significance of the parameters obtained from an analysis. Strictly speaking they are only mathematical numbers describing a curve which is a good approximation to the spectrum. If in an analysis weight is attached to one part of the spectrum, and in another analysis of the same spectrum to another part, the parameters resulting from the analyses are likely to be different in the two cases. The weight can be changed for example by changing the numerical weight ascribed to each point in a least-squares fitting procedure or by imposing some sort of constraints on the parameters to be fitted. An often used method of analysis is to determine the component with the longest lifetime and subtract it from the spectrum. The difference is then considered as an unresolvable mixture of shorter lifetimes. This method attaches great weight to the long-lived component and more or less ignores the detailed shapes of the short-lived ones.

If the longest lifetime is appreciably longer (i.e., two to three times or more) than the shorter ones and has an intensity of at least a few percent, it is so well defined that all reasonable methods of analysis should give the same result for this component, while the results for the shorter-living components may differ (especially when they are of the order of magnitude of the equipment time resolution). However, if the various lifetimes in a spectrum are rather close in magnitude, it may be difficult to extract directly well-defined parameters. During the analysis it may then be necessary to use knowledge about some of the parameters gained from other measurements. The results presented below for water and pure ice are examples of those two different situations.

In general it cannot be said what particular method of analysis should be used for a spectrum which contain more than a resolvable number of components. The only requirement is that the resulting parameters must represent a curve that fits the measured spectrum reasonably well. The method to be chosen for a certain problem should be the one which most efficiently extract the wanted physical information from the spectrum.

TABLE I. Lifetimes and relative intensities for H₂O and D₂O at 20°C. The uncertainties are standard deviations as estimated by the computer.

	$\tau_1 \pm \Delta\tau_1$ (nsec)	$\tau_2 \pm \Delta\tau_2$ (nsec)	$\tau_3 \pm \Delta\tau_3$ (nsec)	$I_1 \pm \Delta I_1$ (%)	$I_2 \pm \Delta I_2$ (%)	$I_3 \pm \Delta I_3$ (%)
H ₂ O	0.22 ± 0.02	0.46 ± 0.02	1.86 ± 0.02	25.0 ± 5.1	48.2 ± 4.7	26.9 ± 0.5
D ₂ O	0.31 ± 0.03	0.53 ± 0.06	2.01 ± 0.03	44.6 ± 12.6	33.2 ± 12.1	22.2 ± 0.6

The problems involved in the analysis of positron lifetime spectra have only been discussed to a rather small extent in the literature. The above considerations and the following discussion of the analysis of our data show, however, that these problems deserve more attention than has usually been paid to them.

IV. EXPERIMENTAL RESULTS

A. Pure Water and Ice

The results published in Ref. 8 demonstrated that a discontinuity exists in the lifetime and relative intensity of ortho-Ps at the ice-water phase transition. In order to be able to use the more refined method of analysis provided by POSITRONFIT,¹² the spectra were remeasured and the temperature region extended down to liquid nitrogen temperature. Within the experimental uncertainties the results of the analysis were in agreement with Ref. 8. The results for light and heavy water at 20°C are given in Table I.

Here and in the following the lifetimes and relative intensities are numbered from shortest to longest lifetime. This is done since no convention is commonly used for results analyzed in terms of more than two lifetimes.

The results for H₂O and D₂O ice as analyzed by a three-term fit are shown in Fig. 1. The standard deviations as estimated by the computer are given as vertical bars on some of the experimental points. For the sake of clarity the points for I_1 and I_2 have been omitted, and only typical uncertainties have been indicated. Furthermore, a few points with extremely large uncertainties have been omitted. Note that large uncertainties in this case indicate that a good fit to the experimental data can be obtained within a wide range of the parameters. Thus, the points shown in Fig. 1 reproduce the data extremely well.

It is seen that particularly the short lifetime $\tau_1 \approx 0.12$ nsec and its intensity I_1 are well-defined quantities by this procedure. As discussed earlier, one can obtain a better determination of parameters and a possibility for physically more meaningful results by introducing constraints, reducing the number of free parameters from essentially five (three lifetimes and two relative intensities) to essentially three or four (for details about the parameters used in the fitting procedure, see Ref. 12).

In Fig. 2 is shown a two-term fit (three parameters) for the same set of data as above. It is seen that the data for H₂O and D₂O agree closely. The curves in Fig. 2 naturally contain essential information about the experimental data, but they do not provide a complete reproduction of the experimental data as is shown by the difference $\Delta_{2,3}$ between the variances of the two-

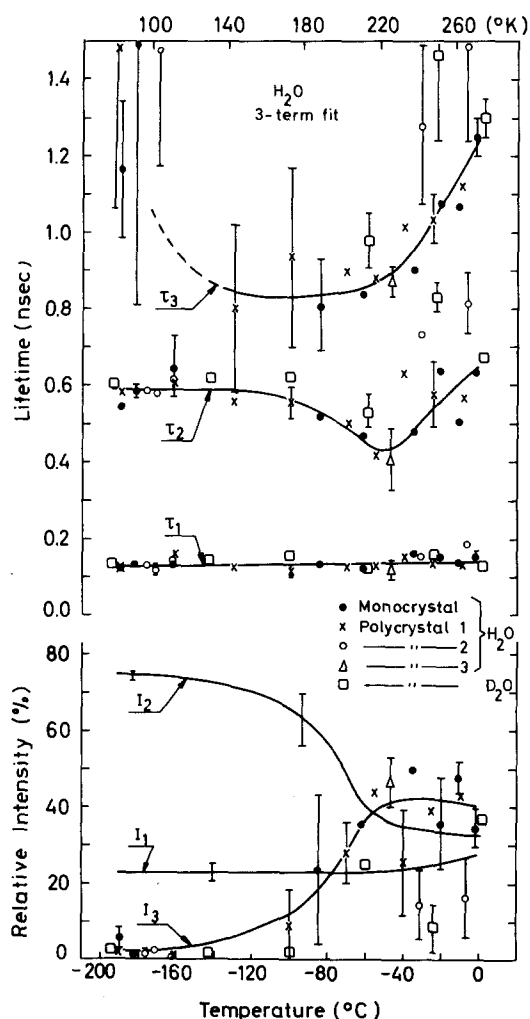


FIG. 1. Lifetimes and relative intensities for a three-term analysis of lifetime spectra for light and heavy ice as function of temperature. The experimental points for I_1 and I_2 have been omitted for the sake of clarity. Typical uncertainties indicated on some of the experimental points are standard deviations as estimated by the computer.

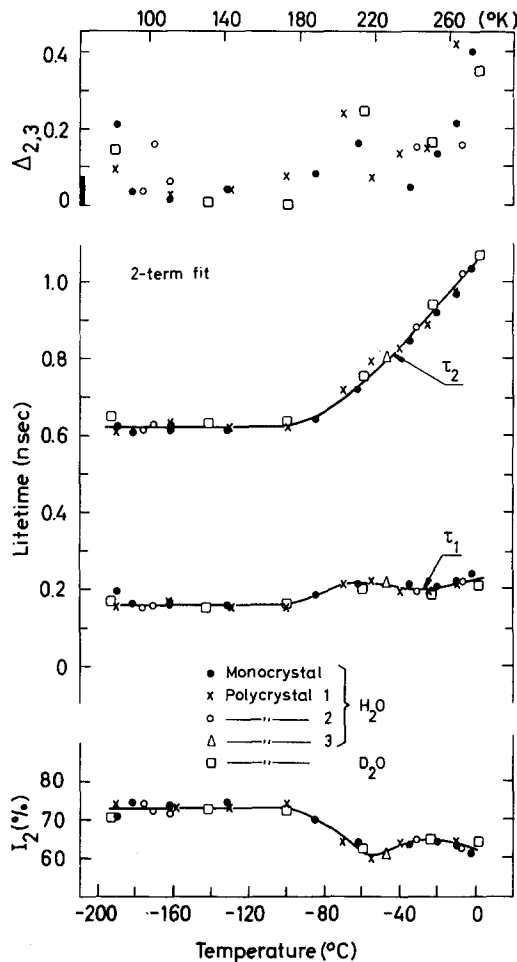


FIG. 2. Lifetimes and relative intensities for a two-term analysis of lifetime spectra for light and heavy ice as function of temperature. $\Delta_{2,3}$ indicates the difference between the variance of a two-term and the corresponding three-term fit. The statistically determined standard deviation for the variance is indicated by a black bar.

term and the corresponding three-term fits. We mentioned earlier that the variance of a three-term fit is the smallest that can be obtained for a particular spectrum. The value of the variance for various spectra is subject to statistical fluctuations. Since these fluctuations are almost eliminated in $\Delta_{2,3}$, this quantity more directly shows the goodness of a two-term fit than does the variance of the fit itself. As the values of $\Delta_{2,3}$ in Fig. 2, except for a few points, are of the order of or larger than the standard deviation of the variance (shown by the black bar), the two-term fits must be considered rather poor.

In order to obtain a physically more useful result we now try a four-parameter fit under the assumption that I_3 is the ortho-Ps annihilation intensity and that it is independent of temperature. This assumption is based upon the following considerations:

1. The Ps-formation probability is expected to

depend only slightly on temperature from 0°C and down, since the energies involved in temperature changes are of the order of 1% or 2% of an electron volt, while energies involved in determination of the Ore gap are of the order of 1 to 10 eV.¹³

2. From angular correlation measurements one finds that the total relative area of narrow components associated with para-Ps is equal to $15.2 \pm 0.5\%$ for temperatures between -182 and 0°C .^{7,10} Taking the pickoff process into account a total para-Ps intensity of $18 \pm 0.5\%$ is found. This entails an ortho-Ps intensity of $54 \pm 1.5\%$, if the reasonable assumption is made that no ortho-para conversion takes place.

3. In the measurements on doped ice described below we find for the higher HF concentrations a long-lived component of intensity $52 \pm 2\%$ at -160°C . It may reasonably be ascribed to ortho-Ps. When these measurements are performed as a function of temperature, the intensity of the long-lived component stays roughly constant, varying between 42% and 54% in the interval from -196 to -20°C .

4. The 3γ -annihilation rate is, to a good approximation, proportional to the product of the ortho-Ps lifetime and its intensity¹:

$$R_{3\gamma} \propto I_{\text{ortho}} \cdot \tau_{\text{ortho}}$$

In Fig. 3 the measurements of Wagner and Hereford¹⁴ of the 3γ rate are compared with the products of the longest lifetimes and their corresponding intensity as

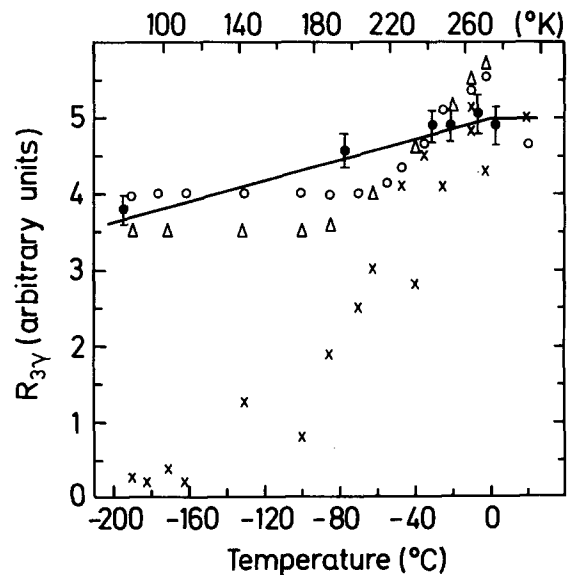


FIG. 3. Rate of 3γ annihilations in ice, as obtained by Wagner and Hereford,¹⁴ compared with the product of the longest lifetime and its relative intensity from the present measurements. Results are shown for different fitting procedures. The ordinate has been normalized to give the best visual fit to the results of Wagner and Hereford. ● (with vertical error bars): Wagner and Hereford; ○: two-term fit; ×: three-term fit; △: three-term fit, $I_3 = 52\%$.

determined by two- and three-term fits of the present measurements. It is evident that the variation of I_{ortho} with temperature must be much smaller than of I_3 as obtained by the first three-term fit.

These points, taken together, lead us to the conclusion that the total amount of ortho-Ps formed in ice must be $52 \pm 2\%$, varying only little with temperature. Accordingly, the lifetime spectra have finally been fitted by use of three lifetime terms, but subject to the constraint that $I_3 = I_{\text{ortho}} = 52\%$ (four-parameter fit). This implies that τ_3 is now assumed to be an average ortho-Ps lifetime. The computed results for H_2O are shown in Fig. 4. In this case the lifetime spectra are quite well represented by the obtained parameter set, as shown by the small differences $\Delta_{3,3}$ between the variance of the four-parameter three-term fit and the five-parameter

three-term fit. These differences are generally considerably smaller than the standard deviation of the variance.

Figure 4 shows that the ortho-Ps lifetime is nearly constant (around 0.68 nsec) for temperatures below -100°C . Above -100°C it increases with increasing temperature. Furthermore, we note that the results for mono- and polycrystalline H_2O agree closely. The obtained product $I_3\tau_3$ has been plotted in Fig. 3 for comparison with $R_{3\gamma}$. The temperature dependence are similar but not quite identical.

One may naturally ask why the results in Fig. 1 are so different from those of Fig. 4. The main reason probably is that the analyzed spectra contain more than three exponential terms, even when corrected for source annihilation. As seen in Fig. 1 a long-lived component of a relative intensity of 1% or 2% is resolved at low temperatures. The influence of this component on the fitted parameters is greatly reduced in the two-term fit and in the three-term fit with the intensity I_3 fixed at 52%. However, it strongly influences the results of the three-term fit without constraints. The origin of this component may be of various natures. It may be an inherent part of a true ice spectrum. It may be caused by small uncontrolled amounts of impurities (the deviating results of polycrystal run 2 in Fig. 1 are probably caused by this), or it may be caused by insufficient source correction. The uncertainties involved in the correction procedures are discussed in Ref. 15.

B. Frozen Solutions of HF

In many areas of ice physics, ice is doped with hydrogen fluoride and ammonia in order to create defects (so-called L and D defects).¹⁶⁻¹⁸ With the purpose of investigating the influence of HF doping on the lifetime of positrons, aqueous solutions of different HF concentrations were fast frozen, and measurements were made at -160°C .

We found that a long-lived component appeared in the spectrum even for rather small HF concentrations. For concentrations above 5×10^{-5} mole fraction its lifetime and relative intensity as determined from several spectra were 1.23 ± 0.02 nsec and $52 \pm 2\%$. As is known from angular correlation results (see previous paragraph) there is good reason to assume that ortho-Ps is formed with a relative yield close to 52% in pure ice. Accordingly we interpret the increased lifetime as due to ortho-Ps trapped in HF-created defects in which the lifetime against pickoff is 1.23 nsec.

A simple model has been used to describe the trapping of positrons.^{3,5} According to the model, positrons in the bulk will either annihilate (with rate λ_b) or become trapped in defects (rate κ , proportional to defect concentration). A positron in a trap will annihilate with the rate λ_t ($=1/\tau_t$). The model shows that a lifetime spectrum for such positrons will have two components,

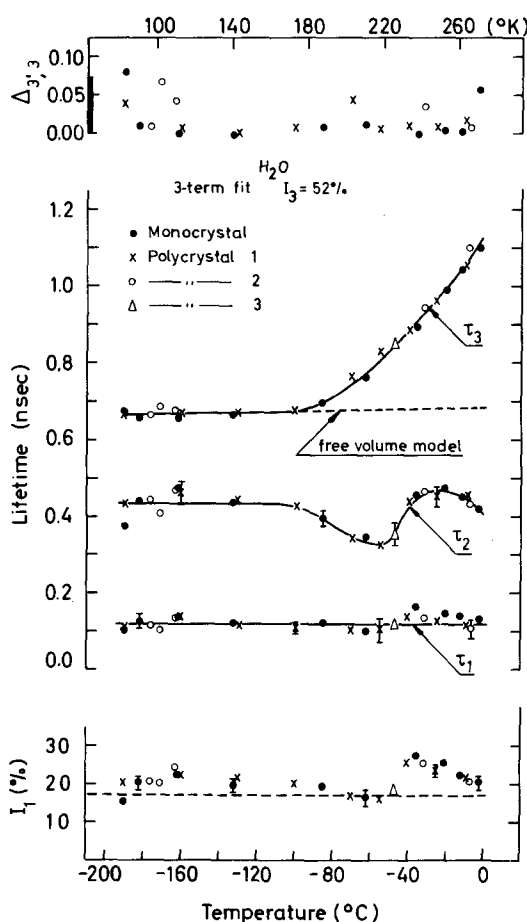


FIG. 4. Lifetimes and relative intensities for a three-term analysis of lifetime spectra for light ice subject to the constraint $I_3 = 52\%$. The indicated typical uncertainties are the standard deviations as estimated by the computer. $\Delta_{3,3}$ is the difference between the variance of this analysis and the corresponding three-term fit without constraints. The statistically determined standard deviation for the variance is indicated by the black bar. The calculated curve for the free volume model is fitted to experiments at -190°C .

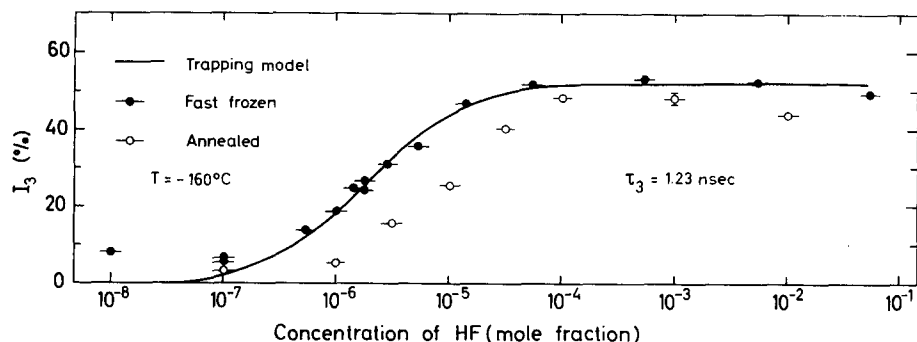


FIG. 5. Relative intensity of the component with lifetime 1.23 nsec as function of HF concentration for fast frozen samples, and for fast frozen samples that were annealed for 20 h at -18°C and subsequently cooled at the rate of about $1^{\circ}\text{C}/\text{min}$. The results are compared with the trapping model, Eq. (1).

one with the decay rate $\lambda_b + \kappa$ and one with the decay rate λ_t . The relative intensity of the latter component is

$$I_t = I_0 \kappa / (\lambda_b - \lambda_t + \kappa), \quad (1)$$

I_0 being the total intensity. By applying this model to trapping of ortho-Ps we analyzed the measured spectra by three-term fits, keeping one lifetime fixed at 1.23 nsec. (This is another case, where results obtained from other measurements are used to extract physical information from lifetime spectra which contain more than three lifetime components). The intensity of the component with lifetime 1.23 nsec as a function of HF concentration is compared with the trapping model curve in Fig. 5 (κ assumed proportional to HF concentration, $I_0 = 52\%$). Furthermore the figure shows the results for samples that after fast freezing were annealed for 20 h at -18°C and then cooled to -160°C with the rate of $1^{\circ}\text{C}/\text{min}$.

In order to further investigate the transition from the fast frozen to the annealed state, samples were heated to different temperatures. Then they were again cooled to -160°C and their lifetime spectra recorded. We found that the transition takes place below -100°C since no

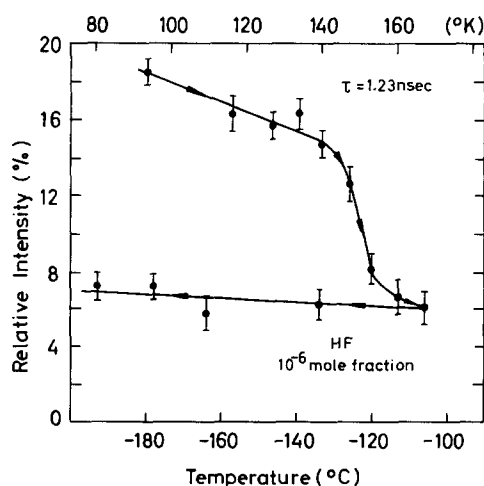


FIG. 6. Transition from fast frozen state to annealed state as indicated by the relative intensity of the component with lifetime 1.23 nsec. The interval between subsequent measurements were 1 day, the cooling rate approximately $1^{\circ}\text{C}/\text{min}$.

further change of the spectra occurred by heating the samples to higher temperatures. The temperature region up to -100°C was subsequently investigated in more detail for a HF concentration of 10^{-6} mole fraction. The intensity of the 1.23 nsec component as function of temperature is shown in Fig. 6. An irreversible transition is clearly seen to take place between approximately -130 and -120°C .

C. Other Frozen Solutions

The main difference between the spectra for frozen solutions and that for pure ice (all at -160°C) was the appearance of a long-lived component of an intensity varying from a few percent to around 50% depending on solution and concentration. To get an approximate quantitative expression for the intensity of this long-lived component all spectra were fitted with three terms, one lifetime being fixed at 1.23 nsec. The intensity of this component gives the wanted information.

Frozen NH_3 -solution measurements were made also at other temperatures. Only small differences from the results of pure ice were seen. The most important results are given in Table II.

A number of other fast frozen solutions of salts and acids in ice were investigated at one or two different concentrations in order to compare with the HF and NH_3 results. None of them influenced the lifetime as strongly as HF. The results are listed in Table III.

TABLE II. Relative intensities of component with lifetime 1.23 nsec at -160°C and lifetime and intensity of long-lived component at -9°C for frozen aqueous solutions of NH_3 . The last three rows were obtained in chronological order from above to below. The uncertainties are standard deviations as estimated by the computer.

Concentration (mole fraction)	Temperature ($^{\circ}\text{C}$)	$\tau_3 \pm \Delta\tau_3$ (nsec)	$I_3 \pm \Delta I_3$ (%)	Comment
4×10^{-2}	-160	1.23	9.1 ± 0.9	fast frozen
10^{-3}	-160	1.23	5.0 ± 0.1	fast frozen
10^{-3}	-9	1.22 ± 0.07	34.5 ± 7.6	
10^{-3}	-160	1.23	5.7 ± 0.5	annealed

V. DISCUSSION

A. Water and the Ice-Water Transition

The light water values for the long lifetime and its relative intensity agree very well with most of the published data on these quantities.¹ This component is attributed to annihilation of ortho-Ps. The relative intensity also agrees with the value of $36 \pm 1\%$ for the total Ps formation as determined by angular correlation magnetic quenching experiments.¹¹

When going from the solid to the liquid phase there is a considerable increase in the long lifetime. Also for a few other molecular materials values of the long lifetime in both solid and liquid phase have been published. Generally the long lifetime is longer in the liquid than in the solid.^{13,19,20} For some substances the change has been correlated to a change in "free volume" (i.e., the volume not occupied by the molecules) in agreement with the "free volume model",^{13,21} but in a few other cases this is not possible²² and the lifetime in the liquid is longer than predicted by the volume change. Water is an example of this.^{8,21}

Ps is able by exchange forces to create a bubble around itself in a liquid. This fact has been used to explain the very long lifetimes (tens of nanoseconds) observed in some liquified gases,² and has also been applied more or less directly to ordinary liquids.^{2,23,24} Because of the reduced electron density in the bubble the ortho-Ps lifetime against pickoff annihilation will be longer than could be expected from the average electron density of the liquid. It is likely that bubble formation is responsible for the general increase in ortho-Ps lifetime in liquids compared to solids as mentioned above.

Another noteworthy fact is the difference between liquid H₂O and D₂O when one considers the long life-

times and relative intensities, as given in Table I. No such difference is found between the two solid compounds.

The viscosity of D₂O is 20% larger than for H₂O.¹⁸ If the bubble model applies for Ps in water, this may be related to a smaller amplitude of oscillation of the bubble volume, which in turn will produce a smaller pickoff annihilation rate. In ice the rigid structure will prevent formation of bubbles, and since no difference exists between the lattice parameters of H₂O and D₂O, no difference in lifetime should be expected. The localization of Ps in a bubble in water may explain the broadening of the narrow component in angular correlation curves²⁵ compared to low temperature ice, in which Ps is delocalized.⁷

An explanation of the differences in Ps-formation probabilities in ice and water and between liquid H₂O and D₂O will probably require a detailed knowledge of the slowing-down processes of positrons and Ps since a simple Ore gap model does not apply to the problem. No attempt to explain these processes will be carried out here.

It seems, however, to be a general feature of a liquid whose molecules contain hydrogen, that deuteration will decrease the Ps yield.^{19,26,27}

One would expect that the components τ_1, I_1 and τ_2, I_2 in a spectrum for water could be associated with annihilation of para-Ps and of free positrons, respectively. This is not directly possible (Table I) since it would imply that $\tau_1 \approx 0.12$ nsec and $I_1 \approx I_3/3$. The lifetime of free positrons in water is best found in an aqueous solution of a Ps oxidizer, such as KMnO₄. The spectrum of this solution contains one single lifetime $\tau_{\text{free}} = 0.395 \pm 0.005$ nsec in both H₂O and D₂O. As discussed earlier, a reason for the fact that the results for pure water (Table I) do not agree with expectations may be that the spectrum contains more than three lifetime components. That will be the case if more annihilation processes take place in water than considered above. Also an incomplete correction for annihilation in the source will have that effect.¹⁵ Actually an analysis with τ_2 fixed at 0.40 nsec gives a reasonably good fit (within one standard deviation of the variance) and the parameters thus obtained are quite close to the expected ones.

TABLE III. Relative intensity of component with lifetime 1.23 nsec for fast frozen solutions measured at -160°C . The uncertainties are standard deviations as estimated by the computer. For comparison values for pure ice and HF are included. The uncertainties on the concentrations are estimated to be $\pm 20\%$.

	Concentration (mole fraction)	$I_3 \pm \Delta I_3$ (%)
Pure ice	0	1.5 ± 0.5
HF	10^{-4}	52 ± 2
LiF	10^{-4}	11.9 ± 0.8
NH ₄ F	1.3×10^{-4}	6.5 ± 0.7
CsF	10^{-4}	8.2 ± 0.7
HCl	10^{-4}	31.2 ± 0.6
HCl	10^{-2}	50.3 ± 0.7
HBr	10^{-4}	37.2 ± 0.8
HI	10^{-4}	23.3 ± 1.4
H ₂ SO ₄	1.1×10^{-4}	3.2 ± 0.6
NaCl	10^{-6}	5.8 ± 0.8

B. Pure Ice

For pure ice at temperatures below -100°C we may ascribe the three lifetime components in Fig. 4 to annihilation of ortho-Ps (τ_3, I_3 , the basis of the analysis), free positrons (τ_2, I_2), and para-Ps (τ_1, I_1), respectively. The value $\tau_2 = 0.44$ nsec seems reasonable compared to the free positron lifetime (0.40 nsec) in water, which is approximately 9% more dense than ice. The value $\tau_1 = 0.12$ nsec is in good agreement with the lifetime of free para-Ps (which becomes 0.105 nsec when pickoff is taken into account). Furthermore, I_1 is fairly close to $\frac{1}{3}I_3$ (the dashed line).

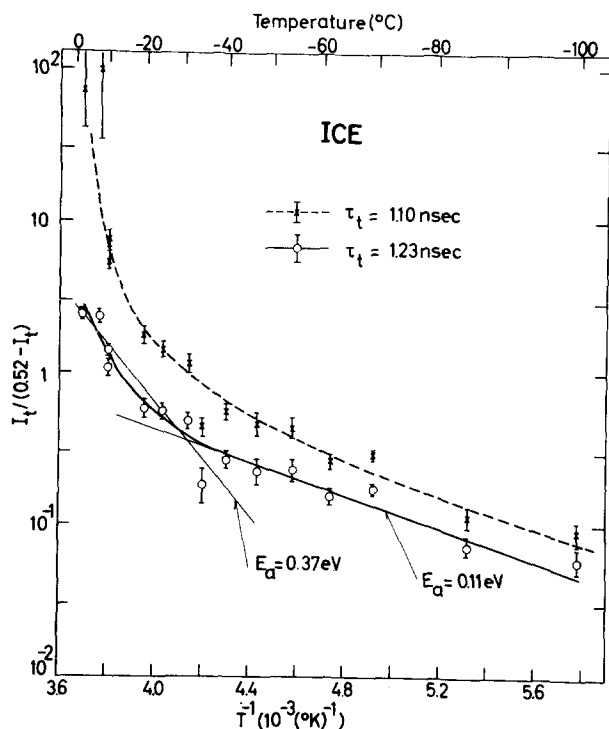


FIG. 7. The quantity $[I_t/(0.52-I_t)]$, which from a simple trapping model is proportional to the number of traps, plotted against $1/T$ (Arrhenius plot) for two different τ_t , the assumed ortho-Ps lifetime in the traps. I_t is the relative intensity of the component with lifetime τ_t . The slopes of the two straight lines estimate activation energies for possible temperature-created defects which trap ortho-Ps.

The rather steep rise of τ_3 with temperature above -100°C exceeds the variation expected from the free volume model.²¹ Ice has a free volume of around 66%, and the cubical expansion of ice is about 1.6% when the temperature is raised from -200° to 0°C .¹⁶ The effect of this temperature change, when calculated from a free volume model,^{21,24} will be an increase in the ortho-Ps lifetime of approximately 2%. The expected variation is shown in Fig. 4. The dashed line was fitted to measurements at -190°C .

A more reasonable explanation of the large increase of the long lifetime is that trapping of ortho-Ps takes place in defects, the number of which increases with temperature. This explanation is supported by the temperature dependence of angular correlation curves for ice single crystals.¹⁰ With increasing temperature the area of the side peaks and the narrow central peak decreases and a component with a width of about 3.2 mrad emerges. This strongly suggests that with increasing temperature an increasing number of para-Ps atoms becomes localized within a few angstroms (i.e., becomes trapped). Hence, also ortho-Ps must be trapped, with a probability larger than for para-Ps, since the ortho-Ps lifetime is longer than that of para-Ps. This involves two characteristic decay rates of ortho-Ps. When defects

are present, also single positrons, not bound in Ps, may be trapped in the way one has seen for metals.³ Hence it is likely that the decay of the positrons not forming Ps will also be governed by two different rates, and the lifetime spectra for ice at higher temperatures may contain at least five components. Therefore, τ_3 in Fig. 4 is only certain to be an average of the longest lifetimes in a spectrum (mainly due to ortho-Ps), and the peculiar behaviors of τ_2 and I_1 may be caused by the mixing of the shorter lifetime components.

If the traps for ortho-Ps are a kind of temperature-created defects, the number of traps N should depend upon temperature T as

$$N = N_0 \exp(-E_a/kT),$$

where N_0 is a constant, E_a is the activation energy for creation of a defect, and k is Boltzmann's constant. Experimentally the relative variation of the number of traps can be determined if the ortho-Ps lifetime in the traps, τ_t , is known. Keeping one lifetime fixed at τ_t in the analysis of the spectra gives the relative intensity I_t of this lifetime component. It is easy to show from Eq. (1) that the number of traps then will be proportional to $N' = [I_t/(0.52-I_t)]$, where the total ortho-Ps yield is 52%. Hence, a plot of $\log N'$ against $1/T$ (Arrhenius plot) will show a straight line if the above mentioned exponential dependence exists. From the slope of the line, E_a can be determined.

This analysis (three terms) has been carried out for different fixed values of τ_t , ranging from 1.1 to 1.7 nsec. The plot is shown in Fig. 7 for the values 1.1 and 1.23 nsec. (The value 1.23 nsec was taken from the results for frozen solutions of HF. However, since the nature of traps is probably different in the two cases, the lifetimes in the traps need not be the same. Actually, Fig. 1 suggests that $\tau_t > 1.23$ nsec in pure ice.) The important point is that it is not possible to fit the points to a single straight line in a reasonable way whatever value is used for τ_t . To determine the order of magnitude of the activation energies involved, the points for $\tau_t = 1.23$ nsec was fitted with two straight lines. The associated energies are 0.11 and 0.37 eV. The same or slightly smaller activation energies are found for larger τ_t values.

The nonlinearity of the Arrhenius plot indicates that the increase of lifetime with temperature is not simply caused by trapping of ortho-Ps in one kind of thermally created defects in which it has one well-defined lifetime.

It is difficult to find a correlation between the different kinds of defects that are generally believed to exist in ice and the trapping of ortho-Ps. An estimated vacancy concentration at -10°C is approximately 10^{-10} mole fraction, while the concentration of interstitials and ionic defects are about 10^{-11} mole fraction.¹⁶ Since it seems to be quite general that defect concentrations have to be of the order of parts per million to be detected by positrons or Ps (previous section and Refs. 3, 4, 5, 6, and 28), these defect concentrations in ice are too small to act as efficient trapping centers. Also the

effect of dislocations on Ps must be too small to be observed. A typical density is 10^6 cm^{-2} which roughly speaking is equivalent to at most 10^{-9} mole fraction of point defects. It is likely that the number of dislocations is different for the various samples, in particular between the mono- and polycrystalline samples. The fact that the results are the same for the different ice samples shows that dislocations have no influence on the positron lifetimes. Neither has any annealing effect been observed, although of two subsequent measurements, the first was normally carried out at a high temperature, and the second at a lower one. The only defects of sufficient concentration are L and D defects. Ice consists of oxygen atoms connected tetrahedrally by hydrogen bonds. An L defect exists if the proton is missing in a bond, and a D defect exists if two protons are in the same bond. At -10°C the concentration of these defects is about 10^{-6} mole fraction, and the energy required to form a pair of these defects is 0.68 eV.^{16,18} The apparent activation energy for the formation of an L or a D defect will then be $\frac{1}{2} 0.68 \text{ eV} = 0.34 \text{ eV}$,¹⁶ close to the 0.37 eV estimated from Fig. 7.

It should be noted, however, that the picture of L defects capturing Ps is not quite consistent with our results for impurities in ice. The concentration of L defects is generally believed to increase with HF doping and decrease with NH_3 doping (and vice versa for D defects). However, both of these dopings tend to increase the long positron lifetime (see Fig. 5 and Table III), thus suggesting that neither L nor D defects are responsible for the trapping of Ps.

C. Impurities in Ice

The good agreement (except for very low HF concentrations) between the trapping model and the measurements on fast frozen HF solutions strongly supports the assumption that ortho-Ps gets trapped in defects created by HF in the ice matrix during fast freezing. By differential thermal analysis Vuillard has shown that a number of fast frozen solutions of acids, bases, and salts are in a glassy state at low temperatures and exhibit a glass transition at a well-defined temperature in the interval between -145 and -110°C .²⁹ In the light of this, our results may be interpreted as follows. By the fast freezing, an amorphous region of H_2O molecules is formed about each impurity, the detailed constitution of which is not known (could be F^- , HF, HF_2^- , or more complex molecules). The delocalized Ps atom can become trapped in the larger holes in the amorphous region and thus, due to more space, it obtains a longer lifetime. The reaction constant for the trapping of ortho-Ps, as calculated from the results of Fig. 5 is 6×10^{12} liters/mole·sec, which is about a hundred times more than the largest quenching constants known for aqueous solutions. The corresponding capture cross section is 22 \AA^2 , when the Ps atom is assumed to have thermal velocity. This large cross

section might be interpreted in terms of delocalization of Ps.

When the amorphous regions are heated above approximately -120°C , an irreversible transition to a crystalline state is apparently taking place (Fig. 6). Due to the presence of the impurities, however, the lattice is distorted and will still give rise to trapping of Ps, but with smaller probability than for the amorphous regions.

We see from Fig. 5 that a characteristic lower defect concentration to be detected by positrons is 10^{-6} to 10^{-5} mole fraction in ice. As mentioned in the previous paragraph this seems to be quite generally the order of magnitude for the trapping center concentration necessary to get a measurable influence on positrons or positronium.

The influence on positron lifetimes by other acids and salts in fast frozen aqueous solutions, as investigated here (Table III), is smaller than that of HF although all of them most probably form glassy phases by fast freezing.²⁹ The hydrogen halides show a quite strong influence, whereas a large molecule like H_2SO_4 has almost no effect. The results for the fluoride salts indicate that the presence of the F^- ion in the solution before freezing does not necessarily lead to a large increase in lifetime. Furthermore, the results in Ref. 30 show that a glassy structure need not entail a lengthening of the ortho-Ps lifetime compared to crystalline ice. An additional fact is that the influence of HCl impurities on the angular correlation curve of ice is stronger than that of HF impurities.⁹

The present results and those discussed in Ref. 9 provide some knowledge about the influence of different impurities on the behavior of positrons and Ps in ice. However, more extensive investigations in the field of frozen aqueous solutions are necessary to get a detailed picture of the different annihilation processes.

VI. CONCLUSIONS

The present work has been concentrated on an experimental investigation of positron lifetimes in the solid and liquid forms of light and heavy water and in frozen aqueous solutions. The temperature region which was investigated ranged from -190 to $+20^\circ\text{C}$. All spectra were resolved into three lifetime components. The following major results were obtained.

Previously⁸ it was established that an abrupt change in ortho-Ps yield and lifetime takes place at the solid-liquid phase transition for both light and heavy water. In the liquid the lifetime is longer and the yield smaller in heavy ($2.01 \pm 0.06 \text{ nsec}$, $22.2 \pm 1.2\%$) than in light water ($1.86 \pm 0.04 \text{ nsec}$, $26.9 \pm 1.0\%$) while in ice no difference is found. Also mono- and polycrystalline samples give identical results. In the temperature range from -190 to -100°C essentially no change of the lifetime spectrum takes place as a function of temperature, in agreement with a free volume model. From

–100°C to the melting point the longest lifetime increased from 0.68 ± 0.02 to 1.12 ± 0.02 nsec, its intensity in both intervals being 52%. The increase in lifetime was ascribed to trapping of ortho-Ps in temperature-created defects although no clear-cut correlation with already known defects could be established. The large Ps yield of $70 \pm 3\%$ found in ice is probably the biggest one observed in any substance.

In frozen aqueous solutions of HF trapping of Ps was found to occur even for concentrations less than parts per million.

The measurements could be interpreted in terms of a very simple trapping theory, thus serving as a first clear demonstration of Ps trapping in defects. At –160°C the ortho-Ps lifetime in the traps was 1.23 nsec. The very high reaction constant associated with the trapping process in fast frozen solutions (about hundred times larger than the highest values known from aqueous solutions) was ascribed to the delocalization of Ps in ice. An irreversible transition was observed about –125°C by heating of a fast frozen HF solution of a concentration of one part per million. It was ascribed to a phase transition from an amorphous to a crystalline state of a region around each impurity. Several other frozen aqueous solutions were investigated, but none of them showed as strong an influence on the lifetime spectrum as HF did.

From ice physics point of view the most interesting problem raised by our work is: What is the nature of the defects which trap Ps (in particular in pure, but also in doped ice).

The present work, together with corresponding measurements by the angular correlation technique, will probably make ice the substance which has been investigated in most detail by positron annihilation. The results obtained for this solid may therefore (although ice in some respects is a rather special substance) serve as a starting point in the interpretation of measurements on other materials. To serve this aim better, however, more measurements on ice would be desirable. Apart from the measurements in progress (see the Introduction) they might comprise an extension of the temperature interval down to liquid helium temperature, lifetime magnetic quenching measurements, and angular correlation magnetic quenching experiments on

monocrystals. Also a detailed determination of the temperature dependence of the 3γ -annihilation rate and the effect of irradiation could be desired.

ACKNOWLEDGMENT

We are grateful to Dr. G. Kvajić for supplying the ice single crystals.

- ¹ V. I. Goldanskii, *At. Energy Rev.* **6**, 3 (1968).
- ² *Positron Annihilation*, edited by A. T. Stewart and L. O. Roellig (Academic, New York, 1967).
- ³ B. Bergersen and M. J. Stott, *Solid State Commun.* **7**, 1203 (1969).
- ⁴ B. T. A. McKee, W. Triftshäuser, and A. T. Stewart, *Phys. Rev. Letters* **28**, 358 (1972).
- ⁵ W. Brandt, H. F. Waung, and P. W. Levy, *Phys. Rev. Letters* **26**, 496 (1971).
- ⁶ W. C. Mallard and F. H. Hsu, *Phys. Letters* **38A**, 164 (1972).
- ⁷ O. Mogensen, G. Kvajić, M. Eldrup, and M. Milošević-Kvajić, *Phys. Rev. B* **4**, 71 (1971).
- ⁸ K. Petersen, M. Eldrup, and G. Trumpy, *Phys. Letters* **31A**, 109 (1970).
- ⁹ M. Milošević-Kvajić, O. Mogensen, G. Kvajić, and M. Eldrup, *J. Chem. Phys.* **56**, 2567 (1972).
- ¹⁰ O. Mogensen *et al.* (unpublished).
- ¹¹ L. Smedskjaer and G. Trumpy (unpublished).
- ¹² P. Kirkegaard and M. Eldrup, *Computer Phys. Commun.* (to be published).
- ¹³ W. Brandt and I. Spirn, *Phys. Rev.* **142**, 231 (1966).
- ¹⁴ R. T. Wagner and F. L. Hereford, *Phys. Rev.* **99**, 593 (1955).
- ¹⁵ M. Eldrup, Ph.D. thesis, Risø Report No. 254, 1971.
- ¹⁶ N. H. Fletcher, *The Chemical Physics of Ice* (Cambridge U. P., London, 1970).
- ¹⁷ *Physics of Ice*, edited by N. Riehl, B. Bullemer, and H. Engelhardt (Plenum, New York, 1969).
- ¹⁸ D. Eisenberg and W. Kauzmann, *The Structure and Properties of Water* (Oxford U. P., London, 1969).
- ¹⁹ S. Y. Chuang and S. J. Tao, *Phys. Letters* **33A**, 56 (1970).
- ²⁰ H. S. Landes, S. Berko, and A. J. Zuchelli, *Phys. Rev.* **103**, 828 (1956).
- ²¹ W. Brandt, S. Berko, and W. W. Walker, *Phys. Rev.* **120**, 1289 (1960).
- ²² W. H. Holt, S. Y. Chuang, A. M. Cooper, and B. G. Hogg, *J. Chem. Phys.* **49**, 5147 (1968).
- ²³ J. Lee and G. J. Celians, *J. Chem. Phys.* **44**, 2506 (1966).
- ²⁴ R. K. Wilson, P. O. Johnson, and R. Stump, *Phys. Rev.* **129**, 2091 (1963).
- ²⁵ P. Colombino, B. Fiscella, and L. Trossi, *Nuovo Cimento* **38**, 707 (1965).
- ²⁶ P. R. Gray, C. F. Cook, and G. P. Sturm, *J. Chem. Phys.* **48**, 1145 (1968).
- ²⁷ K. P. Singh, R. M. Singru, and C. N. R. Rao, *J. Phys. B* **4**, 261 (1971).
- ²⁸ G. Coussot, Ph.D. thesis, Rapport C.E.A., R4138, 1971.
- ²⁹ G. Vuillard, *Ann. Chim. (Paris)* **2**, 231 (1957).
- ³⁰ M. Eldrup, O. Mogensen, and L. Kevan, *Chem. Phys. Letters* **10**, 379 (1971).